



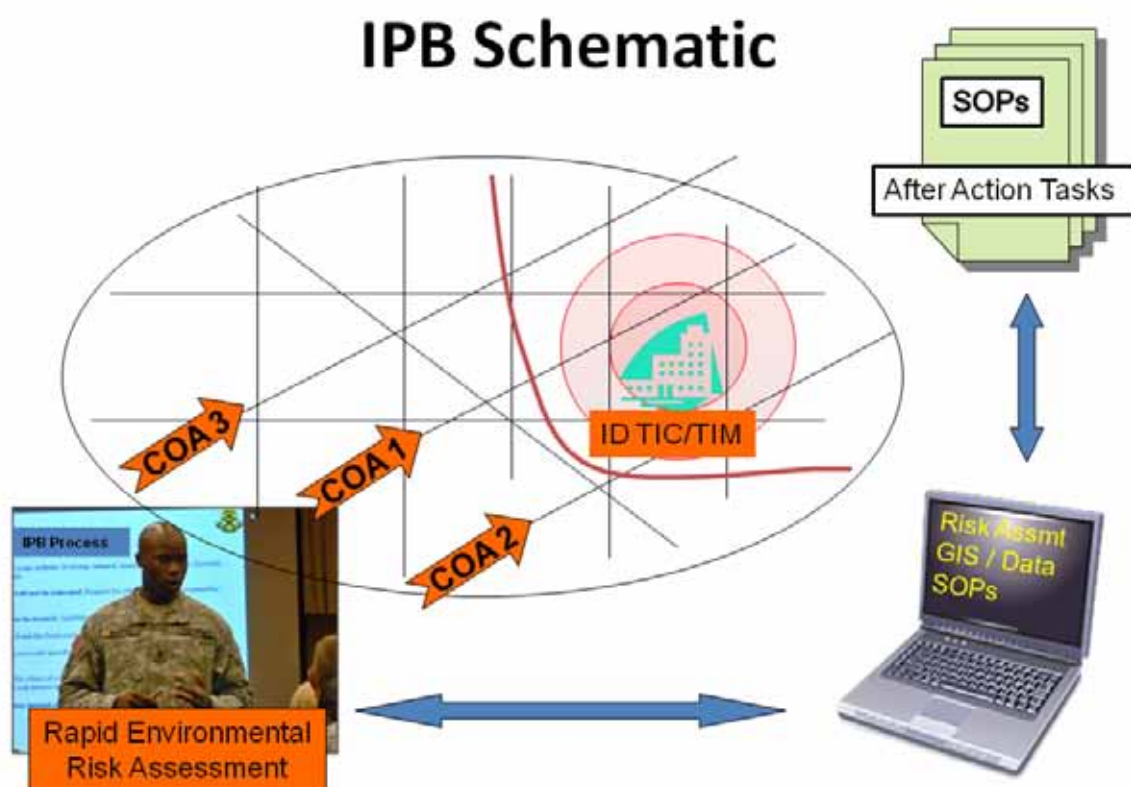
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*Army Environmental Quality/Installations Program*

## Environmental Assessment During Intelligence Preparation of the Battlefield (IPB) – Evaluation of the TIC-Master Computer Program

Victor F. Medina, Dennis L. Brandon, K. James Hay,  
Dhiren S. Khona, Heather K. Smith, and Agnes B. Morrow

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Victor F. Medina, Dennis L. Brandon, Heather K. Smith, and Agnes B. Morrow

*U.S. Army Engineer Research and Development Center  
Environmental Laboratory  
3909 Halls Ferry Road  
Vicksburg, MS 39180-6199*

K. James Hay

*Construction Engineering Research Laboratory  
2902 Newmark Ave.  
Champaign, IL 61826-9005*

Dhiren S. Khona

*Army Geospatial Center  
7701 Telegraph Rd.  
Alexandria, VA 22315-3864*

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## Abstract

The Army needs to manage threats from toxic industrial chemicals (TICs) during the Intelligence Preparation of the Battlefield (IPB) Process. This work unit involved developing a database of TICs that could pose environmental/operational risk, creating an industry/chemical crosswalk, linking the chemical database to a GIS component, applying and evaluating the system in five hypothetical scenarios. The resultant program, which was christened "TIC-Master," has the capability to enhance urban theater of operations analysis to identify TIC/toxic industrial material (TIM) affiliated activities within the battlefield operational footprint and evaluate consequence management alternatives for decision support during the IPB process. TIC-Master was tested in a series of demonstrations. Facilities in the Jackson/Vicksburg MS area were used to test the program. Actual chemical data were obtained for several facilities by securing Tier 2 EPA (Environmental Protection Agency) reports for the facilities of interest. A test team was assembled to develop scenarios based on the industries represented by the facilities. TIC-Master was then used to address the scenarios. Although some deficiencies were identified, the program was successful at encompassing the chemical risk presented by the TICs found at the facilities. Plume modeling was used to assess release scenarios. Subsequent briefings of the program were conducted with potential users at the Army Maneuver Support Center of Excellence (MSCoE) and the Army Intelligency Center (AIC). Constructive suggestions were received and are presented in this document. Overall, the potential users agree that the TIC-Master could be a valuable tool for the Army and other organizations. Ultimately, populating tactical decision-making tools with critical information leads to a greater knowledge of the battlefield and increases survivability.

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## Preface

This report evaluates the TIC-Master Computer Program. This work was part of a project funded by the Army Environmental Quality/Installations Program titled “Environmental Assessment during the Intelligence Preparation of the Battlefield.” This report was prepared by Dr. Victor F. Medina, Environmental Engineering Branch (EEB), Dr. Dennis L. Brandon, Environmental Risk Assessment Branch (ERAB), Environmental Processes and Engineering Division (EPED), Environmental Laboratory (EL), U.S. Army Engineer Research and Development Center (ERDC), Vicksburg, MS; Dr. K. James Hay, ERDC Construction Engineering Research Laboratory (CERL), Champaign, IL; and Dhiren S. Khona, Army Geospatial Center, Alexandria, VA. Technical reviews were provided by Debbie Felt and Dr. Molly Reif, both of ERDC-EL.

This work was conducted under the general supervision of Andy Martin, Chief, EEB, and Warren Lorentz, Chief, EPED. Dr. Elizabeth Ferguson was the Technical Director.

At the time of publication of this report, Dr. Beth Fleming was Director of EL. COL Kevin J. Wilson was Commander of ERDC, and Dr. Jeffery P. Holland was the Director of ERDC.

## Acronyms

%Prob – Percent probability. Used in describing the probability of a chemical being present for a specific industrial process based on analysis conducted by the TIC-Master Database.

AERTA – Army Environmental Requirements and Technology Assessments

AIC – United States Army Intelligence Center (Fort Huachuca)

ALOHA – Areal Locations of Hazardous Atmospheres, a computer program that estimates atmospheric release of chemicals

CAMEO – Computer Aided Management of Emergency Operations

CANUKUS – Canada, the United Kingdom, and the United States. The acronym refers to an international task force consisting of these countries.

CAS – Chemical Abstract Service

CCIR – Commander's Critical Information Requirements

CBRN – Chemical Biological Radiological and Nuclear

CCIR – Commander's Critical Information Requirements

C/JMTK – Commercial Joint Mapping Toolkit

COA: – Course of Action

CoBRA® – Chemical Biological Response Aid chemical database

CONUS: – Within the Continental United States

COP – Common Operational Picture

CDC – Centers for Disease Control

DOA – Department of the Army

DoD – Department of Defense

DPW – Department of Public Works

EPCRA – Emergency Planning and Community Right to Know Act

ERDC – U.S. Army Engineer Research and Development Center

ESRI – Environmental Systems Research Institute. The key vendor of GIS products to US Federal Agencies.

FHP – Force Health Protection

FSO – Full Spectrum Operations

GAO – General Accounting Office

G – Gas

GIS – Geographic Information Systems

GWOT – Global War on Terror

HPAC – Hazard Prediction and Assessment Capability Model

HPC – Hazard Probability Code. A numeric hazard probability score based on physical state, product, and weapon/accident history

HSC – Hazard Severity Category. A numerical ranking combining maximum hazard from toxicity, flammability, and instability)

HSDB – Hazardous Substances Data Bank

HIFLD-WG – Homeland Infrastructure Foundation Level Database Working Group

ICCW – Industry/Chemical Crosswalk

ICSC – International Chemical Safety Cards

IPB – Intelligence Preparation of the Battlefield

ITF-25, 40 – International Task Forces 25 and 40, respectively

JEM – Joint Effects Model

L – Liquid

Lat – Latitude

Long – Longitude

LTC – Lieutenant Colonel

MDMP – Military Decision Making Process

MEMA – Mississippi Emergency Management Agency

MSCoE – United States Army Maneuver Support Center of Excellence  
(Fort Leonard Wood)

OCONUS – Outside the Continental United States

NAICS – North American Industrial Classification System

NFPA – National Fire Protection Association. F - Flammability (ranking),  
H - Health (ranking)

NIOSH – National Institute for Occupational Safety and Health

NOAA – National Oceanic and Atmospheric Administration

ODUSD(I&E) – Office of the Deputy Under Secretary of Defense  
(Installations and the Environment)

RA – Risk Assessment

RM – Risk Management

RR – Risk Ranking. A ranking combining HSC and HPC parameters.

S – Solid

SOP(s) – Standard Operating Procedure(s)

TIC – Toxic Industrial Chemical

TIM – Toxic Industrial Material

TRI – Toxic Release Inventory

UK – United Kingdom

US – United States

USACE – United States Army Corps of Engineers

USACHPPM or CHPPM – U.S. Army Center for Health Promotion and Preventive Medicine. Now known as the Army Public Health Command.

USBCANZPO – United States, Britain, Canada, Australia, and New Zealand Program Office

USEPA or EPA – U.S. Environmental Protection Agency

USDOA – United States Department of the Army

VFD – City of Vicksburg Fire Department (Vicksburg, MS)

WISER – Wireless Information System for Emergency Responders

WMD – Weapons of Mass Destruction

## Definitions

**Environmental considerations:** The spectrum of environmental media, resources, or programs that may impact, or are affected by, the planning and execution of military operations; factors may include, but are not limited to, environmental compliance, pollution prevention, conservation, protection of historical and cultural sites, and protection of flora and fauna (USDOD 2001).

**Environmental hazard:** All activities that may pollute, create negative-noise-related effects, degrade archeological/cultural resources, or negatively affect threatened or endangered species habitats: also included are environmental health-related hazards.

**Environmental stewardship:** The integration and application of environmental values into the military mission in order to sustain readiness, improve quality of life, strengthen civil relations, and preserve valuable natural resources (U.S. Department of Defense (USDOD) 2001).

**Full Spectrum Operations (FSO):** Full spectrum operations describe the combination of offensive, defensive, and stability or civil support operations simultaneously as part of an interdependent joint force to seize, retain, and exploit the initiative, accepting prudent risk to create opportunities to achieve decisive results. They employ synchronized action-lethal and nonlethal-proportional to the mission and informed by a thorough understanding of all variables of the operational environment. Mission command that conveys intent and an appreciation of all aspects of the situation guides the adaptive use of Army forces.

**Hazard:** A condition with the potential to cause injury, or death of personnel; damage to or loss of equipment or property; or mission degradation (USDOD 2001).

**Intelligence Preparation of the Battlefield (IPB):** IPB is an analytical methodology that is conducted at all levels of command to produce graphical intelligence assessments, estimates, and other intelligence products in support of the commander's decision-making process (United

States, British, Canadian, Australian and New Zealand Armies' Program Office 2009).

**Natural environment:** The human ecosystem, including both the physical and biological systems, that provides the resources (i.e. clean air, clean water, healthy surroundings, and sufficient food) necessary to sustain productive human life. Included in the natural resource environment are manmade structures such as water and waste water treatment facilities and natural/cultural resources (U.S. Department of the Army (USDOA) 2004).

## Unit Conversion Factors

Multiply	By	To Obtain
feet	0.3048	meters
gallons (U.S. liquid)	3.785412 E-03	cubic meters
pounds (force)	4.448222	newtons
pounds (force) per square inch	6.894757	kilopascals
pounds (mass)	0.45359237	kilograms

# 1 Introduction

## Intelligence Preparation of the Battlefield

The Intelligence Preparation of the Battlefield (IPB) is a planning process undertaken in preparation for combat operations. It involves gaining key intelligence that will assist commanders in determining the best way to maneuver their forces and apply firepower to gain maximum advantage over their adversaries. The steps involved with the IPB process are:

1. Define the battlefield environment.
2. Describe the battlefield's effects.
3. Evaluate the threat.
4. Determine threat course of action (COA) (U.S. Department of the Army (USDOA) 1994).

Defining the battlefield environment usually focuses on two key pieces of information. The first is the size, disposition, and condition of the enemy forces. The second focuses on the terrain. Although chemical weapons are a critical part of the IPB process (Center for Counterproliferation Research 2000), industrial chemicals have not traditionally been considered as a substantial threat. However, because of changes in the battle space environment, areas of concentrated industrial chemicals, such as those found at industrial facilities, could cause casualties and impact movement if released or if used intentionally (Hauschild and Lee 2004). An extreme example of the latter was the use of released oil and oil fires by defeated Iraqi forces during the first Gulf War. There has been a report of a chlorine tanker being used in an improvised explosive device attack in Afghanistan; however, in that case, the explosion consumed the chlorine gas, limiting its effect.<sup>1</sup>

## Implication of environmental considerations

The U.S. Army has developed policy that values the interdependence among military missions, the global community, and the environment, which is known as full spectrum operations (FSO). Environmental stewardship represents the reflection of leader and individual awareness of and

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<sup>1</sup> Personal Communication. 2011. LTC Robert Tucker, Fort Leonard Wood, MO.

commitment to protecting the environment. It is a proactive, values-based concept that helps to ensure the sustainability and conservation of resources. Environmental considerations include a myriad of specifics dealing with protection of the natural and cultural environment, as well as environmental considerations in Force Health Protection (FHP). Commanders must integrate environmental considerations into the mission analysis phase of the Military Decision Making Process (MDMP) (Jederberg et al. 2002; USDOA 2010). During the mission analysis, a risk assessment (RA), which includes identifying all environmental-related hazards, is ideally conducted (Jederberg et al. 2002). The RA would identify hazards, which the military planners must anticipate, plan for, and mitigate. The risk management (RM) process is one of controlling risk arising from operational factors and balancing that risk with mission accomplishment (Langenberg 2007). Environmental risks are those risks both to and from the environment that must be included in the RM process.

Many of the factors that affect the health of service members may be directly related to environmental considerations within the theater and to the effects on the environment created by military actions. In many cases, environmental considerations need to be integrated into the IPB process. In certain circumstances, environmental considerations may be significant enough to be included in the commander's critical information requirements (CCIR). These environmental considerations will impact all levels of command and should be integrated, as appropriate, into the planning process (USDOA 2010).

## **Urban battlefield**

Discussions conducted with experts from both the Army Maneuver Support Center of Excellence (MSCoE) and the Army Intelligence Center (AIC) indicated that before the Global War on Terror (GWOT), including Operations Enduring Freedom and Iraqi Freedom, American military doctrine was focused on the destruction of foreign standing armies using rapid movement in open spaces. This process was brilliantly displayed during the First Gulf War (Operation Desert Storm), when coalition forces used maneuver and concentrated firepower to route Iraqi forces that had invaded Kuwait. Urban spaces were considered unsuitable for this type of combat model. The use of maneuver and firepower to defeat hostile forces was again successfully demonstrated in the early phases of the liberation of Afghanistan (Operation Enduring Freedom) and the invasion of Iraq (Operation Iraqi Freedom or the Second Gulf War). However, as operations

in both Iraq and Afghanistan continued, American and coalition forces have had to engage in operations in urban environments to deal with insurgents. The U.S. Army has had to develop new strategies, weapons, and equipment for this challenging environment. Urban environments frequently have industrial centers where industrial chemicals are used. Urban operations increase the possibility of troop exposure to industrial chemicals, as well as their release during operations.

## **Toxic industrial chemicals**

Chemicals are a part of modern industry. Production of virtually all goods requires at least some chemical processes. The tenets of toxicology indicate that any chemical can be toxic, as it is just a matter of dosage. However, some chemicals are clearly more problematic than others. One category of chemicals that is of particular concern is the toxic industrial chemicals (TICs) category.

TICs are a special group of chemicals that are deemed of great threat to human health and have the possibility of being weaponized. TICs have been defined by two international task forces (ITFs). The first was ITF-25, which was started in 1994 and consisted of three countries, the United States, the United Kingdom (UK), and Canada (CANUKUS 1994). ITF-25 created a prioritized list of industrial chemicals for military use to assist in identifying, assessing, and controlling exposure risks from industrial chemicals in the battlefield (Steumpfle et al. 1996). This database identifies whether chemicals are listed on the ITF-25 priority list and to what degree (low, medium, or high). The ITF-25 assessed chemicals primarily based on two criteria, toxicity and producibility, which favor chemicals that are acute inhalation hazards produced in large quantities. The ITF-25 priority list consists of 98 chemicals that met certain hazard index criteria. Resta (2001) discussed the activities of international task forces (ITFs, consisting of the United States, the United Kingdom, and Canada) to define and evaluate toxic industrial chemicals. He emphasized the need to evaluate more chemicals than the number included in the first (ITF-25) agreement, which ultimately resulted in ITF-40.

The ITF-40 (which was expanded to include Australia and New Zealand) reviewed, updated, and enhanced the methodology of the ITF-25 work to produce a new list of prioritized industrial chemicals (U.S. Army Center for Health Promotion and Preventive Medicine (USACHPPM) 2003; U.S. British, Canadian, Australian, and New Zealand Armies' Program Office

(USBCANZPO) 2009). The task force determined that the primary deficiency of the ITF-25 list is that it did not consider other important hazardous chemical properties, such as reactivity, ignitability, or corrosiveness. In particular, the task force recognized that a large number of industrial disasters have involved flammable chemicals. ITF-40 attempted to correct this deficiency as well as making several other improvements in ranking hazards and increasing the number of evaluated chemicals. The ITF-40 database includes 1756 chemicals, 1042 of which are ranked, and 63 of which are ranked high enough to be listed as military priority industrial chemical hazards.

Several sources have emphasized the threat of TICs during military operations. In a Master's thesis for the Army Command and General Staff College, Small (2002) concluded "that in the hands of terrorists, aggressive states or nonstate actors, TICs will serve as WMD (weapons of mass destruction), undermine regional stability, and threaten US interests." Small (2002) concluded that joint doctrine for combat and noncombat operations inadequately takes account of current TIC realities. Based on the documents cited in this thesis, "the US Armed Forces have not collectively identified overall force structure requirements for dealing with this threat." In 1998, the Government Accounting Office (GAO) concluded that DoD had not developed doctrine that addressed low-level exposures to chemical agents found on the battlefield. DoD has initiated research to identify the cause (environmental exposures) and effects (medical outcome) relationships associated with various chemicals (Hauschild 2004).

The USBCANZPO (2009) stated that the presence and threat of TICs continues to increase. In environments where normal security regulation has broken down, the failure to secure TIC, CBRN weapons, or precursors, may have significant security or health implications for military personnel. Identification and security of such materials will need to be addressed. This will be especially relevant to operations conducted in complex urban and suburban terrain. As a recent example, senators have asked the Pentagon's inspector general to investigate the precise circumstances surrounding the apparent exposure of Army National Guard soldiers to the chemical sodium dichromate at an industrial facility in Iraq. The issues debated include 1) the Army's role in warning soldiers about the contamination, 2) the personal protective gear provided, and 3) the subsequent testing of soldiers to determine the health impact of the exposure (USEPA 2009).

The USBCANZPO (2009) also provides a preliminary planning guide to enhance preparation, resourcing, and the conduct of operations to mitigate the effects of CBRN/TIM hazards in order to conduct effective operations across the spectrum of conflict. The following CBRN/TIM factors need to be incorporated into the planning process:

- a. Meteorology.
- b. National Infrastructure.
  - 1) Population and industrial centers.
  - 2) Communications (air, road, rail, waterway transportation).
  - 3) Energy resource demands (oil, electricity, gas, nuclear and water).
  - 4) Emergency response capability (mitigation and containment).

Hill (2009) set forth the Army Environmental Requirements and Technology Assessments (AERTA), which represent the critical environmental technology needs for accomplishing the Army's mission while minimizing impact to the environment. These requirements are Army-level requirements and include installation needs only when the need is critical to the execution of the Army's mission. TICs in FSO are applicable to two AERTAs: ER-6-07-01, "Environmental Restoration during Full Spectrum Operations (FSO)," and ER-10-07-01 "Avoidance of Environmental Risk during Contingency Operations" (Hill 2009).

The combined military services have identified TICs as an increasing hazard to military operations. Cutclasure and McClellan (2009) prepared a briefing discussing a Joint Program to improve assessment of TIC hazards in order to improve consequence management approaches.

In addition, the Chief of the U.S. Army Corps of Engineers (USACE), General R.L Van Antwerp, directed the COE to revise policies to evaluate hazardous and toxic agents that may be encountered with overseas contingency operations (Van Antwerp 2010). This further indicates the recognition that these chemicals could present significant hazards to U.S. forces and surrounding native populations.

## **Emergency Planning and Community Right to Know Act**

The Emergency Planning and Community Right to Know Act (EPCRA) was passed in 1986 and has been amended and reauthorized several times since then.

### **Community Right to Know Reporting (EPA Tier II Reporting)**

The EPCRA requires industrial facilities to annually report their hazardous and toxic chemicals to the USEPA, the State agency equivalent, and to local emergency responders, particularly fire departments (Cornell University Law School (CULS), undated; USEPA 2010). Industries that maintain either 10,000 lb of total hazardous chemicals or 500 lb of extremely hazardous materials must conduct inventory reporting.

The reports consist of two parts. Tier I reports must be submitted by 1 March of each year and include total chemicals for specific hazardous chemical categories. Tier II reporting is much more detailed, involving specific hazardous chemicals and their quantities. The definition of hazardous chemicals is quite broad involving lists from other regulations as well as a criteria definition. Thresholds exist for the chemicals, which vary from 1000 kg to as little as 1 kg for the most hazardous. If the chemical is below the threshold, the company is not required to report it.

Tier II reports are an excellent, in fact definitive, source of chemical information for a specific facility, as a wide range of chemicals are reported. One limitation is that if the chemical is not above the threshold, it does not have to be reported. The good news is that most facilities will take a conservative approach and report a below-threshold chemical if (1) there is any chance they might reach or get near the threshold in order to avoid a penalty, and (2) as a courtesy to local emergency responders. A second limitation is that companies can withhold some specific chemicals if they are trade secrets. They do, however, have to report the quantity and properties of any trade secret chemicals.

As mentioned, Tier II reports are valuable records in that they define the chemicals used by a specific facility. Unfortunately, as of this date, Tier II reports are not digital. This means there is no means of easily searching them. This limits their effectiveness in terms of determining chemicals generated from an industry.

### **Toxic Release Inventory (TRI) reporting**

Another component of EPCRA requires industries to file annual reports of chemical releases and wastes. There are approximately 650 chemicals and chemical categories that are subject to reporting by these industries. The USEPA maintains this collected data in the Toxic Release Inventory (TRI),

where the public may view and analyze the data in many different ways. For example, the data may be sorted by facility, geographical location, chemical, and industry. The ability to search these data makes this a valuable resource for predicting chemicals used at industrial facilities. That said, the TRI database does have some key limitations. The chemicals list is relatively small (about 650 chemicals). Second, the chemicals identified are limited only to those released over the timeframe searched.

### **The TIC-Master Project**

TIC exposures may result during military urban operations involving industrial or manufacturing facilities. Even if soldiers are adequately protected, TIC releases can affect neutral or friendly populations in the environs of an urban operation. Better understanding and mapping of the types of TICs that may be present in a battlefield from industrial sources are needed as an important part of an assessment tool that could be incorporated into the IPB Process. A database/GIS system known as the TIC-Master program was built containing over 2,000 TICs, primarily from the ITF-40 and EPA TRI. The database included physical properties, hazard assessment information, and environmental effects and remediation information. A data model was also constructed to relate these chemicals and industrial types using the 6-digit codes from the North American Industrial Classification System (NAICS). Data to relate the chemicals to the industries was queried from the TRI database. Analysis of the data allows for a crude estimation of the probability to encounter specific TICs at a particular industrial facility. A custom GIS standards-based platform has been designed to specifically map industrial facilities within a designated region. The application allows relevant information about the chemical(s) to be queried from the TIC database for a specific site and the mapping of areas of highest hazard probability. The industrial chemical database has been integrated with a GIS mapping method. Testing compared the prediction of primary industrial chemicals at specific industrial locations within the geographical region to the actual chemicals and locations.

## 2 TIC-Master Program

### Approach

In order to evaluate the capability to estimate TICs associated with the industrial process during the IPB process, a computer-based program was created, which will be referred to as the TIC-Master program. This program consists of a database of chemical information based on lists of TICs identified by the ITF-40 and the EPA TRI. This database was then cross-linked to a database that related the chemicals to industries and industrial processes. Finally, this database was placed in a geographic information system (GIS) that allowed for mapping of the facility (or facilities) in an environment that allowed users to identify features in spatial proximity to the site(s), and model releases.

Figure 1 shows how the TIC-Master would be used to improve the IPB process. A course of action (COA) would be proposed. In this case, COA 1 involves interaction with an industrial facility suspected of using TICs. The TIC-Master program will be used to identify TICs associated with the industry, and will also be used to determine some release scenarios. In this case, the pink circles indicate effects of potential chemical releases. It turns out that COA 1 does involve possible exposure to the chemical hazard (be it toxicity, explosion, or fire). Two other possibilities are also evaluating COA 2 and COA 3. COA 2 does involve a possible chemical hazard, but not as significant as COA 1. COA 3 avoids all. However, other considerations may make COA 1 the COA chosen. If so, the TIC-Master can provide chemical information to assist in preparing personnel. Also, environmental standard operating procedures can be used to address environmental issues that may arise from the operation.

### TIC Database

The purpose of the TIC database was to provide hazard information for TICs that may be found in the battlefield and the necessary industrial information for the TIC-Master to provide a reasonable prediction of potential TICs that may be found at a particular type of industrial site. The TIC database was electronically created as a Microsoft Office 97-2003 Excel worksheet file consisting of several spreadsheets, each with a different purpose (Appendix A). For integration into the TIC Master, two spreadsheets were

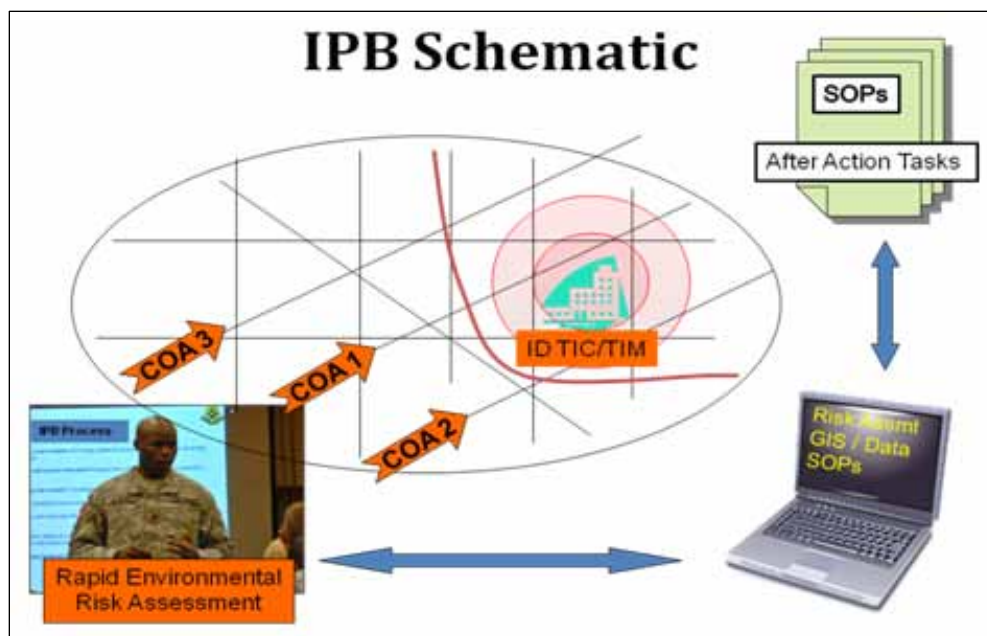


Figure 1. Schematic depicting the use of the TIC-Master program.

utilized. One contains the physical properties and hazard assessment information for each chemical and the other includes a list of industrial types and potential chemicals associated with those industries (industry/chemical crosswalk or ICCW).

## Chemical information

The chemical spreadsheet was created based on a list of 1756 chemicals identified by the ITF-40, (USACHPPM 2003) and is supplemented by 578 listed chemicals from the USEPA Toxic Release Inventory (TRI) – giving a total size of 2092 chemicals (242 chemicals are common between the two lists). For comparison, GeoRAMS, a system designed to model chemical spills that was developed for the U.S. Army, had a library of 44 chemicals (U.S. Army Engineer Research and Development Center (ERDC) – Environmental Laboratory (EL) 2005). Chemicals are presented in ascending order based on their Chemical Abstract Service (CAS) registry number. The spreadsheet includes 26 columns/parameters for each chemical. These include four identifying columns: Chemical ID, CAS, CAS Text, and Chemical Name. Eight parameters are designated for physical properties that include physical state, color, odor, storage state, melting point, boiling point, vapor density, and water solubility. The remaining 14 columns provide hazard assessment data. The hazard assessment columns include National Fire Protection Association (NFPA 2006) diamond ratings; data provided directly from the ITF-40 database;

indicators as to whether the chemical is regarded as a high priority chemical by ITF-40 and ITF-25 (Steumpfle et al. 1996), which was the predecessor task group to ITF-40; and a column for any special considerations to be noted. Each parameter is described in Table 1.

**Table 1. Parameters included in chemical spreadsheet.**

Parameter	Data Type	Description
Chemical_ID	Integer	Primary key for chemical in ascending order
CAS	Integer	Chemical Abstract Service registry number in integer form (minus hyphens)
CAS_Text	Text	Chemical Abstract Service registry number in text form as commonly presented with hyphens
Chemical_Name	Text	Chemical name, usually a commonly accepted systematic name
State	Text	Physical state near 25 degrees Centigrade, listed as G (gas), L (liquid), S (solid), or any combination of the three
Color	Text	Color in the visible spectrum, listed as “no” or simple color description
Odor	Text	Odor description, listed as “no”, “yes”, or simple description
Storage_State	Text	Likely state that chemical will be stored: S (solid), L (liquid in closed container or cylinder), G (compressed gas in vessel), SL (solid or liquid in closed container, LG (liquefied compressed gas in vessel), GLG (liquefied or non-liquefied gas in vessel), GIG (compressed gas diluted with inert gases), GAC (compressed gas dissolved in acetone), A (aqueous liquid solution), Ssol (Solid or liquid solution other than aqueous), SA (solid or aqueous liquid solution)
Melting_Point	Numeric	Melting point in degrees Centigrade
Boiling_Point	Numeric	Boiling point in degrees Centigrade
Vapor_Density	Numeric	Relative vapor density (air = 1.0)
Water_Solubility	Text	Simple description of water solubility (e.g., insoluble, poor, very good, miscible, reaction)
Exposure_Route	Text	Primary exposure route risk as determined by ITF-40
NBCFilter_Protection	Text	NBC filter protection as provided by FM 3-11.4, Table E-1: E (effective), M (marginal), and P (poor)
NFPA_H	Integer	NFPA 704 health hazard rating (0 to 4); or in cases where NFPA 704 health rating were not inputted, ITF-40 health/toxicity score based on highest toxicity rating for three potential exposure routes (1 to 4)
NFPA_F	Integer	NFPA 704 flammability hazard rating (0 to 4)

Parameter	Data Type	Description
NFPA_I	Integer	NFPA 704 instability hazard rating (0 to 4)
NFPA_Special	Text	NFPA 704 special hazard, limited to W (reacts violently or explosively with water) and OX (possesses oxidizing properties)
HPV	Text (Boolean)	High production volume as imported from ITF-40 data: yes or no
ITF25_Hazard	Text	ITF -25 hazard rating: low, medium, or high
ITF40_Priority	Text (Boolean)	Indication of inclusion in ITF-40 military priority industrial chemical hazards, table II-13 (ref): yes or no
ITF40_HSC	Text	ITF-40 Hazard Severity Category: maximum hazard rank from ITF-40 health/toxicity, NFPA704 flammability, and NFPA 704 instability (negligible, marginal, critical, catastrophic)
ITF40_HPC	Text	ITF-40 Hazard Probability Category: summed probability rank based on physical state, production, and history (unlikely, seldom, occasional, likely, or frequent)
ITF40_RR	Text	ITF-40 Risk Ranking: matrix application of ITF-40 HSC and ITF-40 HPC to provide overall risk (low, moderate, high, or extreme)
ITF40_History	Text	ITF-40 history converted to text (accident, listed, or weapon)
Special Considerations	Text	Comments that may provide pertinent additional information

Physical property data and NFPA hazard ratings were collected through the ChemIDplus Lite data searching tool provided by the United States National Library of Medicine (<http://chem.sis.nlm.nih.gov/chemidplus/chemidlite.jsp>). This tool allows a search for chemical information from numerous databases. Of these databases, there were four that provided pertinent physical property data included in this spreadsheet: National Institute for Occupational Safety and Health (NIOSH) International Chemical Safety Cards (NIOSH, undated), NIOSH Pocket Guide to Chemical Hazards (NIOSH 2005), Computer Aided Management of Emergency Operations (CAMEO), and the National Library of Medicine Hazardous Substances Data Bank (HSDB). It should be noted that sometimes the data within these sources are conflicting and best judgment was used to decide on the values entered into the chemical spreadsheet.

## Industry/chemical crosswalk

The industries provided in the industry/chemical crosswalk (ICCW) are listed by the six-digit North American Industrial Classification System (NAICS) codes, which were obtained via <http://www.census.gov/eos/www/naics/>. There are 1,175 six-digit codes, of which 126 were populated with data.

The ICCW provides data that can be used to predict chemicals that may be present at a particular industrial site if the six-digit NAICS code is known or can be determined. Chemical/industry relationships were developed by reviewing the USEPA TRI data for given industries. Custom queries were performed for each six-digit code to list all reported releases of chemicals by a given industry type. These data allow for the generation of a list of chemicals reported, the number of facilities reporting, the frequency a particular chemical is reported, and the probability that a particular chemical will be reported as released by a facility that can be identified by that particular six-digit NAICS code. If the chemicals and probabilities generated by this procedure are assumed to be characteristic of this industry type, then these data can be used to predict chemicals and their likelihood of being found at this type facility. It should be noted that custom queries used to populate the data in the ICCW were performed during May and June 2009 utilizing the TRI database that was current at that time. The exact procedure used to gather and manipulate these data is provided in Appendix B.

Several limitations are associated with the method for populating the ICCW. For each six-digit NAICS code, there may be many subtypes of industries. This means that chemical lists generated will contain possible chemicals for all the subtypes under a particular industry type, which in many cases will be excessive. The chemicals list generated will only contain TRI listed chemicals, of which there are 578 (only includes chemicals with an assigned CAS and does not include chemical mixtures). So, many chemicals that are in the chemical spreadsheet will not be included in a prediction of chemicals. Also, the TRI data only include data from industrial sites located within the United States, which may limit applicability to other countries.

The industry/chemical crosswalk spreadsheet contains six data columns: ICCW ID, NAICS Code, NAICS Title, CAS, TRICHEM Name, and TRI% Prob. These parameters are defined in Table 2. Table 3 lists the industries with populated data in the industry/chemical crosswalk with the number of

Table 2. Parameters included in ICCW.

Parameter	Value	Description
ICCW_ID	Integer	Primary key for industry/chemical relationship in crosswalk, in ascending order
NAICS_Code	Integer	6-digit NAICS code identifying industry
NAICS_Title	Text	Text description of NAICS industry listing
CAS	Integer	Chemical Abstract Service registry number in integer form (minus hyphens)
TRICHEM_NAME	Text	Chemical name as provided by USEPA TRI database
TRI%_Prob	Numeric	Probability that chemical was reported in the USEPA TRI database by a facility associated with the 6-digit NAICS code

Table 3. Industry types with populated data in ICCW with number of predicted chemicals.

NAICS Code	NAICS Title	# Chemicals
212319	Other Crushed and Broken Stone Mining and Quarrying	1
212324	Kaolin and Ball Clay Mining	11
221112	Fossil Fuel Electric Power Generation	76
221310	Water Supply and Irrigation Systems	6
221320	Sewage Treatment Facilities	6
311119	Other Animal Food Manufacturing	37
311222	Soybean Processing	26
311225	Fats and Oils Refining and Blending	25
311421	Fruit and Vegetable Canning	28
311511	Fluid Milk Manufacturing	18
311615	Poultry Processing	41
311919	Other Snack Food Manufacturing	25
311991	Perishable Prepared Food Manufacturing	9
312111	Soft Drink Manufacturing	21
321113	Sawmills	36
321114	Wood Preservation	43
321211	Hardwood Veneer and Plywood Manufacturing	26
321918	Other Millwork (including Flooring)	26
321991	Manufactured Home (Mobile Home) Manufacturing	30
322121	Paper (except Newsprint) Mills	90
322130	Paperboard Mills	81
322211	Corrugated and Solid Fiber Box Manufacturing	33

NAICS Code	NAICS Title	# Chemicals
322212	Folding Paperboard Box Manufacturing	34
322214	Fiber Can, Tube, Drum, and Similar Products Manufacturing	24
324110	Petroleum Refineries	158
324121	Asphalt Paving Mixture and Block Manufacturing	42
324191	Petroleum Lubricating Oil and Grease Manufacturing	105
324199	All Other Petroleum and Coal Products Manufacturing	75
325110	Petrochemical Manufacturing	182
325120	Industrial Gas Manufacturing	137
325131	Inorganic Dye and Pigment Manufacturing	71
325132	Synthetic Organic Dye and Pigment Manufacturing	105
325181	Alkalies and Chlorine Manufacturing	173
325182	Carbon Black Manufacturing	39
325188	All Other Basic Inorganic Chemical Manufacturing	285
325191	Gum and Wood Chemical Manufacturing	88
325192	Cyclic Crude and Intermediate Manufacturing	140
325193	Ethyl Alcohol Manufacturing	52
325199	All Other Basic Organic Chemical Manufacturing	286
325211	Plastics Material and Resin Manufacturing	277
325212	Synthetic Rubber Manufacturing	136
325221	Cellulosic Organic Fiber Manufacturing	89
325222	Noncellulosic Organic Fiber Manufacturing	112
325311	Nitrogenous Fertilizer Manufacturing	76
325312	Phosphatic Fertilizer Manufacturing	32
325314	Fertilizer (Mixing Only) Manufacturing	77
325320	Pesticide and Other Agricultural Chemical Manufacturing	330
325411	Medicinal and Botanical Manufacturing	156
325412	Pharmaceutical Preparation Manufacturing	171
325413	In-Vitro Diagnostic Substance Manufacturing	22
325414	Biological Product (except Diagnostic) Manufacturing	32
325510	Paint and Coating Manufacturing	156
325520	Adhesive Manufacturing	166
325611	Soap and Other Detergent Manufacturing	120
325612	Polish and Other Sanitation Good Manufacturing	113
325613	Surface Active Agent Manufacturing	121

NAICS Code	NAICS Title	# Chemicals
325620	Toilet Preparation Manufacturing	60
325910	Printing Ink Manufacturing	64
325920	Explosives Manufacturing	57
325991	Custom Compounding of Purchased Resins	87
325992	Photographic Film, Paper, Plate, and Chemical Manufacturing	105
325998	All Other Miscellaneous Chemical Product and Preparation Manufacturing	268
326140	Polystyrene Foam Product Manufacturing	46
326150	Urethane and Other Foam Product (except Polystyrene) Manufacturing	71
326160	Plastics Bottle Manufacturing	25
326191	Plastics Plumbing Fixture Manufacturing	32
326192	Resilient Floor Covering Manufacturing	38
326199	All Other Plastics Product Manufacturing	116
326299	All Other Rubber Product Manufacturing	95
327112	Vitreous China, Fine Earthenware, and Other Pottery Product Manufacturing	15
327215	Glass Product Manufacturing Made of Purchased Glass	51
327332	Concrete Pipe Manufacturing	12
327390	Other Concrete Product Manufacturing	35
327910	Abrasive Product Manufacturing	48
331111	Iron and Steel Mills	100
331221	Rolled Steel Shape Manufacturing	51
331316	Aluminum Extruded Product Manufacturing	55
331319	Other Aluminum Rolling and Drawing	42
331422	Copper Wire (except Mechanical) Drawing	54
331491	Nonferrous Metal (except Copper and Aluminum) Rolling, Drawing, and Extruding	63
332311	Prefabricated Metal Building and Component Manufacturing	50
332312	Fabricated Structural Metal Manufacturing	68
332321	Metal Window and Door Manufacturing	62
332439	Other Metal Container Manufacturing	69
332812	Metal Coating, Engraving (except Jewelry and Silverware), and Allied Services to Manufacturers	85
332813	Electroplating, Plating, Polishing, Anodizing, and Coloring	102

NAICS Code	NAICS Title	# Chemicals
332911	Industrial Valve Manufacturing	34
332912	Fluid Power Valve and Hose Fitting Manufacturing	28
332919	Other Metal Valve and Pipe Fitting Manufacturing	42
333319	Other Commercial and Service Industry Machinery Manufacturing	41
333518	Other Metalworking Machinery Manufacturing	21
333911	Pump and Pumping Equipment Manufacturing	46
333924	Industrial Truck, Tractor, Trailer, and Stacker Machinery Manufacturing	53
333991	Power-Driven Handtool Manufacturing	30
334111	Electronic Computer Manufacturing	42
334513	Instruments and Related Products Manufacturing for Measuring, Displaying, and Controlling Industrial Process Variables	36
335122	Commercial, Industrial, and Institutional Electric Lighting Fixture Manufacturing	33
335211	Electric Housewares and Household Fan Manufacturing	29
335311	Power, Distribution, and Specialty Transformer Manufacturing	50
335313	Switchgear and Switchboard Apparatus Manufacturing	46
335911	Storage Battery Manufacturing	36
335929	Other Communication and Energy Wire Manufacturing	47
335931	Current-Carrying Wiring Device Manufacturing	51
335999	All Other Miscellaneous Electrical Equipment and Component Manufacturing	58
336111	Automobile Manufacturing	91
336370	Motor Vehicle Metal Stamping	61
336399	All Other Motor Vehicle Parts Manufacturing	101
336419	Other Guided Missile and Space Vehicle Parts and Auxiliary Equipment Manufacturing	23
336612	Boat Building	52
336999	All Other Transportation Equipment Manufacturing	40
337122	Nonupholstered Wood Household Furniture Manufacturing	57
337124	Metal Household Furniture Manufacturing	30
337127	Institutional Furniture Manufacturing	61
337211	Wood Office Furniture Manufacturing	50
337214	Office Furniture (except Wood) Manufacturing	49

NAICS Code	NAICS Title	# Chemicals
337920	Blind and Shade Manufacturing	32
339999	All Other Miscellaneous Manufacturing	108
424410	General Line Grocery Merchant Wholesalers	3
424690	Other Chemical and Allied Products Merchant Wholesalers	164
424710	Petroleum Bulk Stations and Terminals	89
424910	Farm Supplies Merchant Wholesalers	6
493120	Refrigerated Warehousing and Storage	1
561499	All Other Business Support Services	1
562112	Hazardous Waste Collection	213
562211	Hazardous Waste Treatment and Disposal	372
722310	Food Service Contractors	1

chemicals predicted for each industry. Although the 126 industries covered were just a little over 10% of the industrial possibilities, the range of industries and industrial processes was impressive, ranging from chemical manufacturing and processing, to food service industries, to furniture manufacturing. As an example, 34 industrial types were identified in the Vicksburg/Jackson area, and the TIC-Master industry list thoroughly encompassed these industries.

## Integrated Geographic Information System

### Introduction

Both the intelligence and emergency management communities rely on having accurate, timely, and authoritative sources of geospatial information for their respective missions (Haefner 2004, Channell et al. 2009). Geographic Information Systems (GIS) play a critical role for visualization, and spatial analysis of population, critical infrastructure, and potential impact area to help manage disasters, and share a common operational picture (COP) amongst the stakeholders.

The primary purpose of the TIC-GIS effort was to research methods to design a toxic industrial chemicals database linked to a GIS, develop a framework to share the TIC database for peer review, and receive feedback from the emergency management community.

### Objectives:

1. Develop a TIC data model and a data dictionary.
2. Develop a capability to query the TIC database on a map.
3. Demonstrate the value of a geographic information system through relevant spatial analysis and automated reporting to support emergency management.

### Methodology

A relational TIC data model (Figure 2) was designed based on the tables containing the toxic industrial chemicals and relevant properties created for this project.

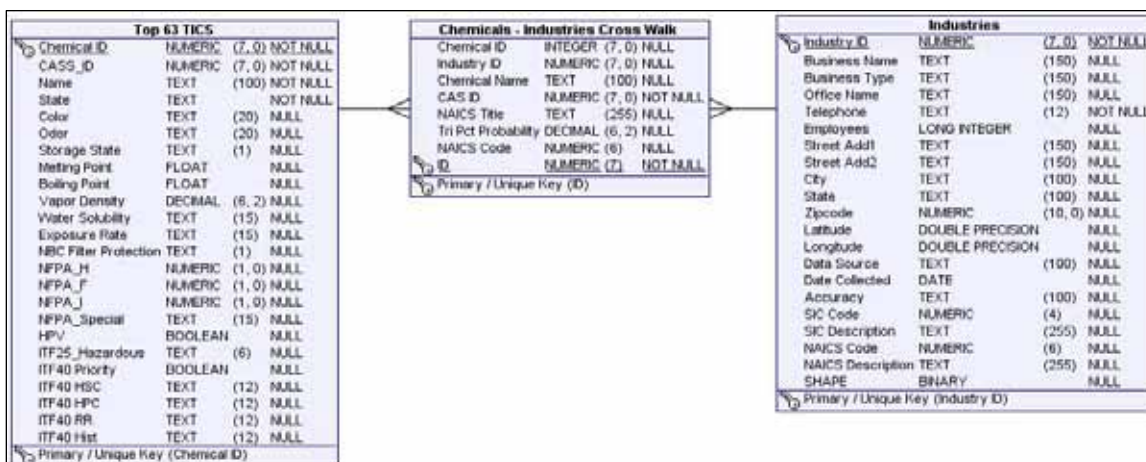


Figure 2. Relational TIC data model.

This model was developed using best industry practices for designing a normalized relational model to include a crosswalk table for handling the multiple relationships between the chemicals and industries tables, respectively.

A TIC database was created with the data model in an Environmental Systems Research Institute (ESRI) file geodatabase format. The ESRI file geodatabase format was selected for a rapid prototyping of a TIC database linked to a geographic information system. The data model enables a user to click a facility on a map and retrieve the list of chemicals likely to be found at that facility based on the facility type (Figure 3).

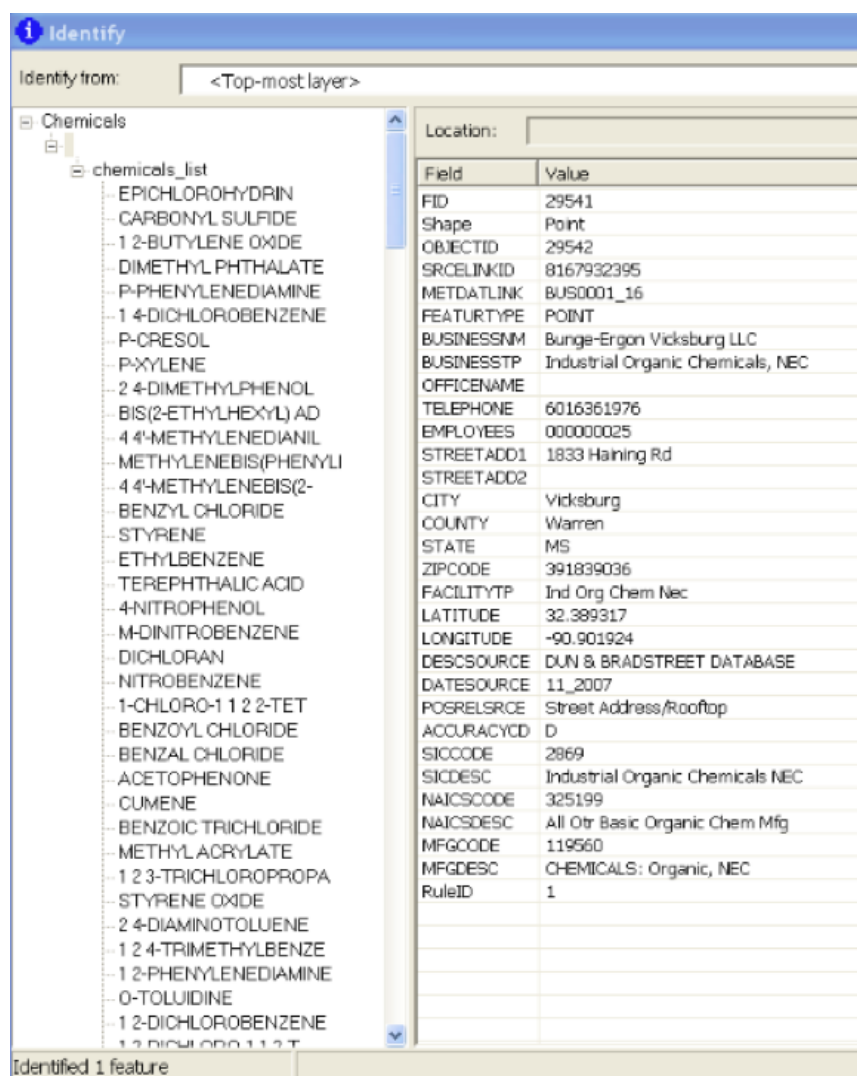


Figure 3. Pull-down screen showing facility type.

Users can click on a chemical to get its properties, as illustrated in Figure 4 for the chemical “phosgene.”

ESRI software is part of the Commercial Joint Mapping Toolkit (C/JMTK) and a viable and completely compatible mechanism for transitioning capabilities to the Warfighter (Figure 5). A data dictionary was developed, containing tables, attributes, definitions of attributes, and attribute type.

A custom graphical user interface was designed in ARCGIS to demonstrate the value of linking a TIC database to an emergency management system developed around a geographic information system. The custom code was written in Visual Basic with the ESRI ARCOBJECTS library. The customized interface allows a user to navigate based on latitude/longitude, city,

**i Identify**

Identify from:

Field	Value
OID	38
ICCW_ID	39
NAICS_TITL	All Other Basic Or
CAS	75445
TRICHEM_NA	PHOSGENE
TRI_PROB	2.12
NAICS	325199
OID_1	101
CHEMICAL_I	102
CAS_1	75445
CAS_TEXT	75-44-5
CHEMICAL_N	phosgene
STATE	G
COLOR	no
ODOR	yes
STORAGE_ST	LG
MELTING_PO	-118
BOILING_PO	8
VAPOR_DENS	3
WATER_SOLU	reaction
EXPOSURE_R	Inh
NBCFILTER_	E
NFPA_H	4
NFPA_F	0
NFPA_I	1
NFPA_SPECI	
HPV	yes
ITF25_HAZA	high
ITF40_PRIO	yes
ITF40_HSC	Critical
ITF40_HPC	Likely
ITF40_RR	High
ITF40_HIST	Weapon
SPECIAL_CO	

Identified 1 feature

Figure 4. Phosgene example.

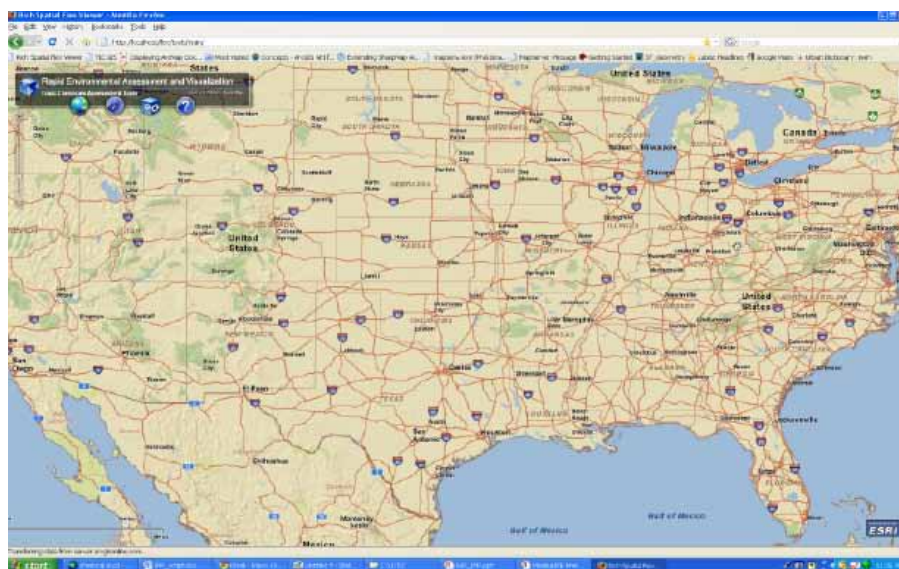


Figure 5. ESRI GIS interface.

state, and ZIP code in the continental United States (Figure 5). The GIS system could be applied to areas all over the world. However, some of the features were only available within the United States. Databases are available to apply most of these capabilities worldwide, but were not investigated as part of the scope for this project since it was developed as a scoping project to see how the concept would function using primarily free source data.

The system allowed for searches of key entities, such as fire and police departments and schools. The system also had the capability to search for facilities that use industrial chemicals within the United States (Figure 6 shows an example of facilities identified within the Warfighter III training area). For the system tested, this feature was only available within the United States (by using the “Homeland Security Infrastructure Protection Database” provided by the National Geospatial-Intelligence Agency); however, international databases exist that allow for similar capabilities around the world (such as the “Directory of World Chemical Producers”). In unstable areas where armed conflict occurs, industrial activity information may be inaccurate. However, most industries have international ties in terms of chemical and machinery sources and in terms of export agreements. Therefore, these records are probably fairly useful for most situations. A polygon feature allows for estimation of populations within a defined area based on census data (Figure 7). Once again, this feature is available with international database add-ons.

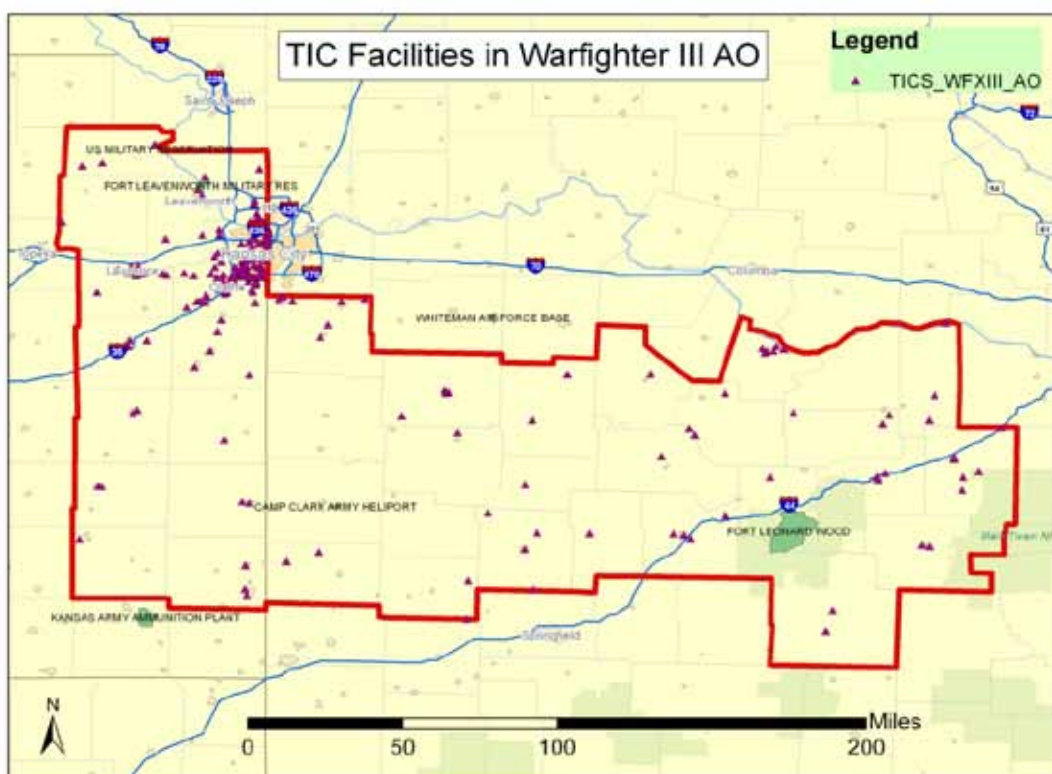


Figure 6. Location of toxic industrial chemical facilities in the 2009 Warfighter III test area.

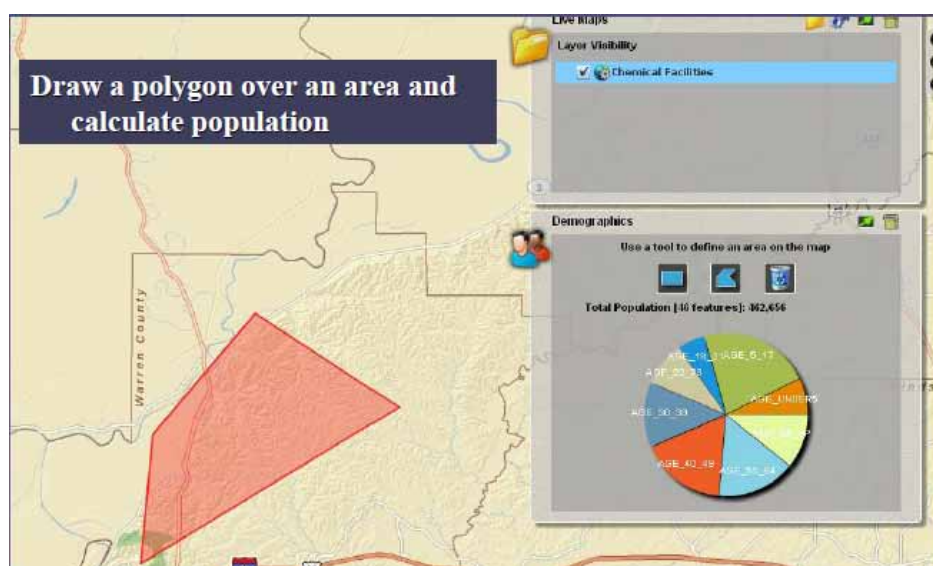


Figure 7. Demonstration of the polygon tool to obtain census data.

## Outreach

An outreach activity was used to expose the TIC-Master Data Model to the Emergency Management Community through the Geospatial Acquisition Support Directorate of the Army Geospatial Center, and the Homeland

Infrastructure Foundation Level Database Working Group (HIFLD-WG). Within the HIFLD-WG, outreach involved collaboration with Craig P. Adams from the Office of the Deputy Under Secretary of Defense (Installations and the Environment) (ODUSD(I&E)). Mr. Adams is the Business Enterprise Integration Program Manager responsible for the central DoD chemicals database. This project provided value-added attributes, including the probability of a toxic industrial chemical to be found at a particular location based on the facility type. The Army Data Model Program Manager was contacted and utilized to transition the TIC data model into the Army Data Model.

### **Release modeling using ALOHA (Areal Locations of Hazardous Atmospheres)**

ALOHA (Areal Locations of Hazardous Atmospheres) was selected as the means of release modeling for this project. On 27 May 2009, Dr. Victor Medina and Agnes Morrow interviewed representatives of the Vicksburg Fire Department (VFD) (Vicksburg, MS) on emergency response issues to guide the development of the TIC-Master. The VFD indicated that they use a number of release models, but that ALOHA is the one they rely on in most cases and that ALOHA is commonly used by other first responder organizations. ALOHA was jointly developed by the National Oceanic and Atmospheric Administration (NOAA) and the USEPA and is designed especially for use by emergency responders, as well as for emergency planning and training. ALOHA models key hazards (toxicity, flammability, thermal radiation (heat), and overpressure (explosion blast force)) related to chemical releases that result in toxic gas dispersions, fires, and/or explosions. ALOHA runs quickly on small, easily transportable computers (Microsoft® Windows® or Macintosh®). It is designed to be easy to use so that it can be successfully operated during high-stress situations. Its chemical library contains information about the physical properties of approximately 1,000 common hazardous chemicals. Its computations represent a compromise between accuracy and speed; ALOHA has been designed to produce good results quickly enough to be of use to responders. ALOHA is also designed to minimize operator error as it checks entered information and warns the user when a mistake seems evident. ALOHA's on-screen help offers quick access to explanations of its features and computations, as well as background information to help in interpreting its output. Although ALOHA was the only program used in this study, other release modeling programs could easily be added to the GIS interface.

## Standard operating procedures for environmental releases of selected chemicals

A goal of the project was to provide detailed environmental response information for TICs – Standard Operating Procedures (SOPs). This process required a detailed literature review for each chemical. Based on the resources for the project, it was determined that it would not be possible to create an SOP for the >2000 chemicals in the database. So, the focus was narrowed to the 63 chemicals identified as most problematic in ITF-40.

Of these 63 chemicals, 35 existed only as a gas. Gaseous releases would likely have a minimal long-term environmental impact, so SOPs were not prepared for those materials. Therefore, SOPs were prepared for 28 chemicals. Figure 8 is an example of the SOP for acrolein. Appendix C contains the SOPs prepared for the project.

<b>Acrolein</b>	<b>107-02-8</b>
<b>ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION</b>	
<b>Air</b>	
Acrolein emitted to air reacts primarily with photochemically generated hydroxyl radicals in the troposphere (Ghilarducci & Tjeerdema, 1995). Minor processes include direct photolysis, reaction with nitrate radicals, and reaction with ozone (Atkinson et al., 1987; Haag et al., 1988a; Howard, 1989; BUA, 1994). Acrolein has been detected in rainwater, indicating that it may be removed by wet deposition (Grosjean & Wright, 1983). The calculated atmospheric half-life of acrolein, based on rate constants for hydroxyl radical reaction, is between 3.4 and 33.7 h (Atkinson, 1985; Edney et al., 1986b; Haag et al., 1988a; Howard, 1989; Howard et al., 1991; BUA, 1994). The overall reactivity-based half-life of acrolein in air, as estimated by Mackay et al. (1995), is less than 10 h. Based on these short estimated half-lives, acrolein is not a candidate for long-range atmospheric transport.	
<b>Water</b>	
Acrolein is removed from surface water primarily by reversible hydration, biodegradation by acclimatized microorganisms, and volatilization (Irwin, 1987; Haag et al., 1988b; Howard, 1989; ATSDR, 1990; Springborn Laboratories, 1993). In groundwater, acrolein is removed by anaerobic biodegradation and hydrolysis (Chou & Spanggard, 1990a). The overall reactivity-based half-life of acrolein in surface water is estimated to be between 30 and 100 h (Mackay et al., 1995). In groundwater, half-lives of 11 days and 336–1344 h (14–56 days) are estimated based on aerobic and anaerobic degradation, respectively (Howard et al., 1991). Observed dissipation half-lives of acrolein applied as a herbicide in irrigation Acrolein 9 canals range from 7.3 to 10.2 h (Jacobson & Gresham, 1991a, b, c; Nordone et al., 1996a). The relatively short observed half-lives of acrolein in surface waters make long-range aquatic transport unlikely. However, groundwater contamination could result contaminant exposures.	

Figure 8. Example of part of an SOP for acrolein.

### 3 Testing of TIC-Master

#### Overview

A testing program was developed to evaluate the TIC-Master program. Key aspects of the test addressed the following questions:

- How accurate was the TIC-Master at predicting specific chemicals from a given industrial process and at determining an overall chemical threat?
- Could TIC-Master prove useful in operational planning (particularly using release modeling)?
- How easy was the program to use?
- What could be improved?

A test team consisting of Dr. Dennis Brandon, Dr. Heather Smith, and Agnes Morrow, developed test scenarios and evaluated the results. This group was not involved in any way with the development of the TIC-Master program. Nor were these team members involved with the TIC-Master operation during testing. The TIC-Master operation was conducted by Dhiren Khona, Dr. James Hay, and Dr. Victor Medina. The Vicksburg/Jackson MS area was chosen as the area to test, primarily because it was convenient for the ERDC Vicksburg staff to obtain industrial information for this area.

#### Scenario development

Five scenarios were developed and submitted to the test team separately. Each was submitted using a standard form (Figure 9), which provided background information for the test team. The industrial process, TIC release mechanism (i.e., area secured after a fierce aerial assault, explosion due to unknown cause, suicide bombers drive vehicles into facility) and atmospheric conditions changed with each scenario. Four scenarios provided the industrial process or NAICS. The other scenarios provided a partial or complete list of chemicals. Several scenarios included multiple sites, other hazards, and additional site characteristics. The primary route of exposure alternated between air, water, and surface soil. The sites varied in their proximity to civilian populations, water bodies, and intra-structure (i.e., schools, hospitals, and historical landmarks).

Test __ Scenario		
Date: _____ Time Submitted: _____ Time Team Response Received: _____		
Team Initial Start Time: _____ Time Team Response Submitted: _____		
<u>Industrial Processes or Process Name:</u>	<u>TIC Release Mechanism:</u>	<u>Atmospheric Conditions:</u> Temperature: Wind Speed: Wind Direction: Precipitation:
<u>Suspected TICs:</u>	<u>Other Industrial Chemicals:</u>	<u>Other Potential Hazards:</u>
<u>Site Designation:</u> Single Site: __ Multiple Sites: __ General Location*: __	Site A*: Longitude: Latitude: Elevation:	Site B: Longitude: Latitude: Elevation:
<u>Other Characteristics:</u>		
* See the <u>Other Characteristics</u> section. For example, the U.S. Army may wish to occupy a 2-square-mile area centered at the location provided by Site A*.		

Figure 9. Test scenario submission form.

The selected facilities were in the vicinity of Jackson/Vicksburg, MS (Table 4). Copies of 2008 EPA TIER 2 reports were obtained from the Mississippi Emergency Management Agency (MEMA). This request was based on the Freedom of Information Act. The evaluation team used these reports to validate the site chemical lists provided by the test team.

Table 4. List of facilities used for TIC-Master testing.

Name	NAICS	Location
AKZO Nobel Coatings Inc.	325510	Clinton, MS
Bunge-Ergon Vicksburg, LLC	311222	Vicksburg, MS
Clorox Products Manufacturing Co.	325612	Pearl, MS
Double G Coatings Co., L.P.	332813	Jackson, MS
Ergon Refining, Inc.	324110	Vicksburg, MS
Georgia Gulf Corp. and Vinyls	325991	Hazelhurst, MS
International Paper	322121	Redwood, MS
Jackson Plating Co.	332813	Jackson, MS
Nissan North America	336111/336112	Canton, MS

## 4 Test Results

### Test 1

#### Scenario

The scenario assumes an explosion has occurred and remnants of the subsequent fire remain. The NAICS code was 325510. Figure 10 is the Test 1 scenario description form.

Test <u>1</u> Scenario		
Date: <u>10/09/09</u> Time Submitted: <u>0830 CDT</u> Time Team Response Received: <u>09:00</u>		
Team Initial Start Time: <u>09:45</u> Time Team Response Submitted: <u>14:45</u>		
We worked on the test from 09:45 to 10:45, then from 13:15 to 14:45. total time 3 hours		
<u>Industrial Processes or Process Name:</u>  NAICS: 325510	<u>TIC Release Mechanism:</u>  An explosion has occurred. Remnants of the subsequent fire remain.	<u>Atmospheric Conditions:</u> Temperature: 79° F Wind Speed: 20 mph Wind Direction: NE Precipitation: None
<u>Suspected TICs:</u>	<u>Other Industrial Chemicals:</u>	<u>Other Potential Hazards:</u>  Flammable liquids are on site.
<u>Site Designation:</u> Single Site: <u>X</u> Multiple Sites: <u>   </u> General Location*: <u>   </u>	Site A*: Longitude: 90° 21' 47.16" W Latitude: 32° 21' 02.57" N Elevation:	Site B: Longitude: Latitude: Elevation:
<u>Other Characteristics:</u>		
* See the <u>Other Characteristics</u> section. For example, the U.S. Army may wish to occupy a 2-square-mile area centered at the location provided by Site A*.		

Figure 10. Scenario description for Test 1.

## Results

### *Response time*

The total response time was 3.5 hr, above the 2.0-hr goal. The challenge came from sorting out chemical hazards, which was done manually for this test.

### *Facility location and identification*

Figure 11 is a location map of the facility based on the GIS coordinates. The facility was correctly identified as a painting and coatings manufacturing facility. The actual facility was AKZO Nobel Coatings.

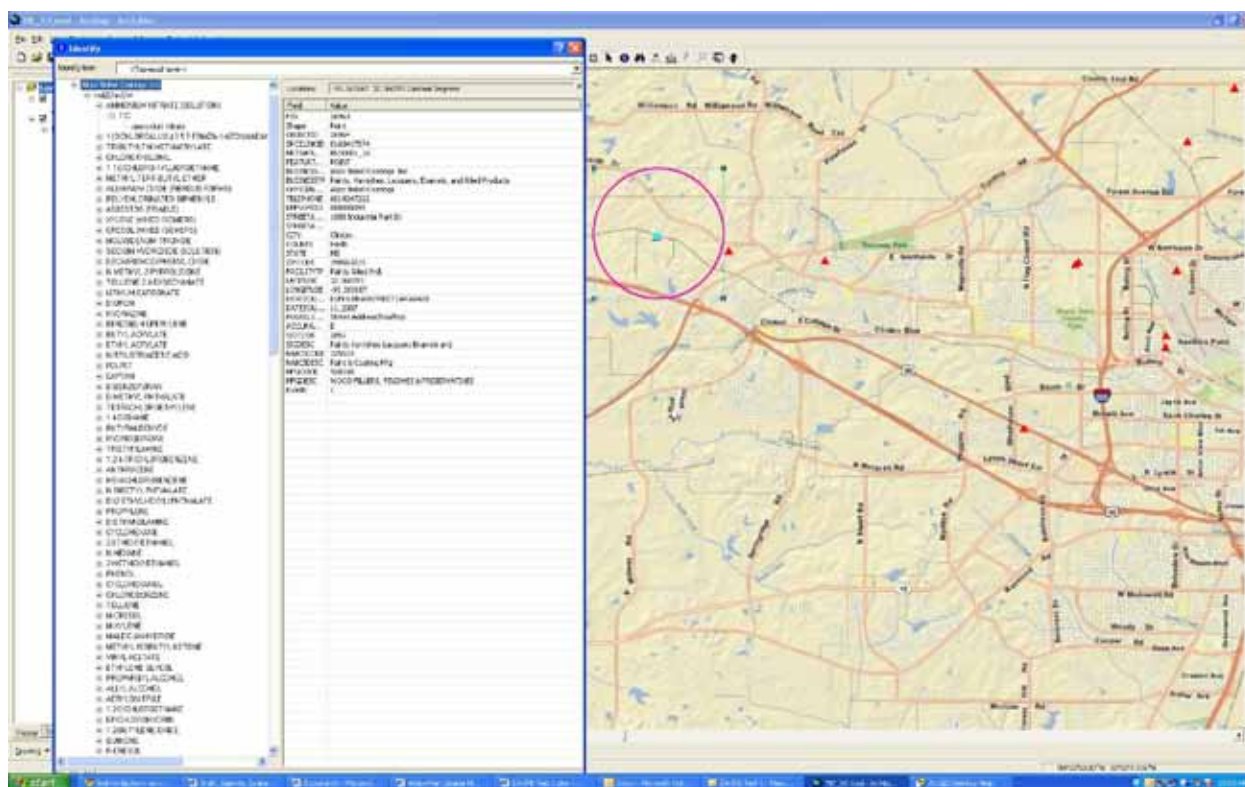


Figure 11. Location map for facility identified in Test 1.

### *Chemical identification*

The 2009 Tier II report (reporting from 1 January to 31 December 2008) submitted by AKZO Nobel Coatings Inc. was used to verify the chemicals identified by the test team. The test team identified 156 TICs (Table 5).

Table 5. Chemical identification list for Test 1 scenario. CAS is the chemical abstracts services number for the chemical.

CAS	Chemical Name	CAS	Chemical Name
50000	FORMALDEHYDE	80626	METHYL METHACRYLATE
56235	CARBON TETRACHLORIDE	84662	DIETHYL PHTHALATE
56359	BIS(TRIBUTYL TIN) OXIDE	84742	DIBUTYL PHTHALATE
57749	CHLORDANE	85018	PHENANTHRENE
64186	FORMIC ACID	85449	PHTHALIC ANHYDRIDE
64675	DIETHYL SULFATE	85687	BUTYL BENZYL PHTHALATE
67561	METHANOL	91087	TOLUENE-2 6-DIISOCYANATE
67630	ISOPROPYL ALCOHOL	91203	NAPHTHALENE
67641	ACETONE	92524	BIPHENYL
68122	N N-DIMETHYLFORMAMIDE	94360	BENZOYL PEROXIDE
71363	N-BUTYL ALCOHOL	95476	O-XYLENE
71432	BENZENE	95487	O-CRESOL
71556	1 1 1-TRICHLOROETHANE	95501	1 2-DICHLOROBENZENE
74873	CHLOROMETHANE	95636	1 2 4-TRIMETHYLBENZENE
75058	ACETONITRILE	96333	METHYL ACRYLATE
75092	DICHLOROMETHANE	98828	CUMENE
75354	VINYLDENE CHLORIDE	98884	BENZOYL CHLORIDE
75456	CHLORODIFLUOROMETHANE	98953	NITROBENZENE
75558	PROPYLENEIMINE	100210	TEREPHTHALIC ACID
75569	PROPYLENE OXIDE	100414	ETHYLBENZENE
75650	TERT-BUTYL ALCOHOL	100425	STYRENE
75694	TRICHLOROFLUOROMETHANE	100447	BENZYL CHLORIDE
75718	DICHLORODIFLUOROMETHANE	101144	4 4'-METHYLENEBIS(2-CHLOROANILINE)
76131	FREON 113	101688	METHYLENEBIS(PHENYLISOCYANATE)
77736	DICYCLOPENTADIENE	101779	4 4'-METHYLENEDIANILINE
78922	SEC-BUTYL ALCOHOL	101906	DIGLYCIDYL RESORCINOL ETHER
78933	METHYL ETHYL KETONE	103231	BIS(2-ETHYLHEXYL) ADIPATE
79005	1 1 2-TRICHLOROETHANE	105679	2 4-DIMETHYLPHENOL
79016	TRICHLOROETHYLENE	106423	P-XYLENE
79061	ACRYLAMIDE	106445	P-CRESOL
79107	ACRYLIC ACID	106514	QUINONE
79469	2-NITROPROPANE	106887	1 2-BUTYLENE OXIDE
79947	TETRABROMOBISPHENOL A	106898	EPICHLOROHYDRIN
80057	4 4'-ISOPROPYLIDENEDIPHENOL	107062	1 2-DICHLOROETHANE
80159	CUMENE HYDROPEROXIDE	107131	ACRYLONITRILE

CAS	Chemical Name	CAS	Chemical Name
107186	ALLYL ALCOHOL	302012	HYDRAZINE
107197	PROPARGYL ALCOHOL	330541	DIURON
107211	ETHYLENE GLYCOL	554132	LITHIUM CARBONATE
108054	VINYL ACETATE	584849	TOLUENE-2 4-DIISOCYANATE
108101	METHYL ISOBUTYL KETONE	872504	N-METHYL-2-PYRROLIDONE
108316	MALEIC ANHYDRIDE	1163195	DECABROMODIPHENYL OXIDE
108383	M-XYLENE	1310732	SODIUM HYDROXIDE (SOLUTION)
108394	M-CRESOL	1313275	MOLYBDENUM TRIOXIDE
108883	TOLUENE	1319773	CRESOL (MIXED ISOMERS)
108907	CHLOROBENZENE	1330207	XYLENE (MIXED ISOMERS)
108930	CYCLOHEXANOL	1332214	ASBESTOS (FRIABLE)
108952	PHENOL	1336363	POLYCHLORINATED BIPHENYLS
109864	2-METHOXYETHANOL	1344281	ALUMINUM OXIDE (FIBROUS FORMS)
110543	N-HEXANE	1634044	METHYL TERT-BUTYL ETHER
110805	2-ETHOXYETHANOL	1717006	1 1-DICHLORO-1-FLUOROETHANE
110827	CYCLOHEXANE	1897456	CHLOROTHALONIL
111422	DIETHANOLAMINE	2155706	TRIBUTYL TIN METHACRYLATE
115071	PROPYLENE	4080313	1-(3-CHLOROALLYL)-3 5 7-TRIAZA-1-AZONIAADAMANTANE CHLORIDE
117817	DI(2-ETHYLHEXYL) PHTHALATE	6484522	AMMONIUM NITRATE (SOLUTION)
117840	N-DIOCTYL PHTHALATE	7429905	ALUMINUM (FUME OR DUST)
118741	HEXACHLOROBENZENE	7439921	LEAD
120127	ANTHRACENE	7439965	MANGANESE
120821	1 2 4-TRICHLOROBENZENE	7439976	MERCURY
121448	TRIETHYLAMINE	7440020	NICKEL
123319	HYDROQUINONE	7440224	SILVER
123728	BUTYRALDEHYDE	7440360	ANTIMONY
123911	1 4-DIOXANE	7440393	BARIUM
127184	TETRACHLOROETHYLENE	7440439	CADMIUM
131113	DIMETHYL PHTHALATE	7440473	CHROMIUM
132649	DIBENZOFURAN	7440484	COBALT
133062	CAPTAN	7440508	COPPER
133073	FOLPET	7440666	ZINC (FUME OR DUST)
139139	NITRILOTRIACETIC ACID	7632000	SODIUM NITRITE
140885	ETHYL ACRYLATE	7647010	HYDROCHLORIC ACID (1995 AND AFTER "ACID AEROSOLS" ONLY)
141322	BUTYL ACRYLATE	7664382	PHOSPHORIC ACID
191242	BENZO(G H I)PERYLENE	7664393	HYDROGEN FLUORIDE

CAS	Chemical Name	CAS	Chemical Name
7664417	AMMONIA	12122677	ZINEB
7664939	SULFURIC ACID (1994 AND AFTER "ACID AEROSOLS" ONLY)	20325400	3,3'-DIMETHOXYBENZIDINE DIHYDROCHLORIDE
7697372	NITRIC ACID	25321226	DICHLOROBENZENE (MIXED ISOMERS)
7726956	BROMINE	26471625	TOLUENE DIISOCYANATE (MIXED ISOMERS)
7782492	SELENIUM	35691657	1-BROMO-1-(BROMOMETHYL)-1,3-PROPANEDICARBONITRILE
7782505	CHLORINE	55406536	3-iodo-2-propynyl butylcarbamate
10294345	BORON TRICHLORIDE	60207901	PROPICONAZOLE

The Tier II report listed 58 chemicals, but of these, only 35 were part of the database used for the project. The test team identified 15 of the 35 chemicals in the database (Table 6). The 20 remaining chemicals are listed in Table 7. Furthermore, the test team identified chemicals similar to those found at the site, including at least one chemical in the following classes: alkane, aldehyde, alcohol, ketone, aromatic, amide, and inorganic metals.

Table 6. Chemicals properly identified in Test 1.

CAS #	Chemical	Reported Max Daily Quantity (lbs)
50000	FORMALDEHYDE	1,000-9,999
67561	METHANOL	10,000-99,999
67630	ISOPROPYL ALCOHOL	100,000-999,999
67641	ACETONE	100,000-999,999
71363	N-BUTYL ALCOHOL	10,000-99,999
78933	METHYL ETHYL KETONE	10,000-99,999
95636	1,2,4-TRIMETHYLBENZENE	10,000-99,999
100414	ETHYLBENZENE	10,000-99,999
108101	METHYL ISOBUTYL KETONE	10,000-99,999
108883	TOLUENE	10,000-99,999
108952	PHENOL	0-99
121448	TRIETHYLAMINE	1,000-9,999
1330207	XYLENE (MIXED ISOMERS)	10,000-99,999
7429905	ALUMINUM (FUME OR DUST)	1,000-9,999
7664417	AMMONIA	1,000-9,999

Table 7. Chemicals not precisely identified by the TIC-Master process for Test 1.

CAS No.	Chemical	NFPA	HSC	Reported max daily quantity (lbs)
64-17-5	Ethyl alcohol	H2, F3	Critical	10,000-99,999
78-83-1	Isobutanol	H2, F3	Critical	10,000-99,999
97-85-8	isobutyl isobutyrate	H1, F3	Critical	1,000-9,999
98-56-6	Parachlorobenzofluoride	H1	Negligible	10,000-99,999
104-15-4	p-toluene sulfonic acid	H2, F1	Marginal	10,000-99,999
108-21-4	Isopropyl acetate	H1, F3	Critical	1,000-9,999
108-65-6	1-methoxy-2-acetolpropane	H1, F2	Marginal	10,000-99,999
110-19-0	Isobutyl acetate	H2, F3	Critical	10,000-99,999
111-76-2	Butoxyethanol	H3, F2	Critical	10,000-99,999
110-43-0	Heptanone	H4, F2	Catastrophic	10,000-99,999
123-86-4	Butyl acetate	H2, F3	Critical	10,000-99,999
763-69-9	Ethyl-3-ethoxypropionate	H1	Negligible	10,000-99,999
1314-13-2	Zinc oxide	H2	Marginal	10,000-99,999
1314-98-3	Zinc sulfide			10,000-99,999
1333-86-4	Carbon black	H2	Marginal	10,000-99,999
8052-41-3	Petroleum hydrocarbon	H2, F2	Marginal	10,000-99,999
21645-51-2	Aluminum trihydroxide		Indeterminant	10,000-99,999
34590-94-8	Methoxymethyl ethoxypropanol	H2, F2	Marginal	1,000-9,999
64742-89-8	Aliphatic hydrocarbon		Indeterminate	10,000-99,999
42978-66-5	Tripropylene glycol diacrylated	H1	Marginal	1,000-9,999

### *Most probable chemicals*

The system allowed for ranking the probability of chemicals that would be found. The top ten chemicals are given in Table 8. Eight of these top ten chemicals were actually found in the inventory, including the first six chemicals on this list.

### *Most significant chemical hazard*

Chlormethane, benzene, and styrene were the most significant hazards identified by the test team. This recommendation was based on the assumption that flammability was the greatest hazard. None of these chemicals were reported in the AKZO Tier II report. However, as these chemicals had flammability hazard indexes of 3 and 4, their effect would be well within flammable hazards from other chemicals at the site.

**Table 8. The 10 most probable chemicals predicted by the TIC-Master database for Test 1. Chemicals found in the actual inventory are starred.**

%Probability	CAS	NAME	Rank
65.41	1330207	XYLENE (MIXED ISOMERS)*	1
56.16	108883	TOLUENE*	2
47.51	78933	METHYL ETHYL KETONE*	3
35.81	108101	METHYL ISOBUTYL KETONE*	4
32.66	71363	N-BUTYL ALCOHOL*	5
31.88	100414	ETHYLBENZENE*	6
29.17	107211	ETHYLENE GLYCOL	7
26.46	67641	ACETONE*	8
25.15	95636	1 2 4-TRIMETHYLBENZENE	9
23.32	67561	METHANOL*	10

There are, however, other ways to rank the most significant chemical hazards. HSC is a hazard severity ranking based on the ITF 40's method. Table 9 summarizes the chemicals identified by the TIC-Master program and found at the site. It is clear that there are several chemicals that could be severe chemical hazards.

**Table 9. Evaluation of TICs in Test 1 for health hazards.**

CAS No.	Chemical	NFPA	HSC	Reported max daily quantity (lbs)
110-43-0	2 heptanone	H-4, F-2	Catastrophic	10,000-99,999
108-95-2	Phenol	H-4, F-2	Catastrophic	0-99
111-76-2	2 butoxyethanol	H-3, F-3	Critical	10,000-99,999
67-56-1	Methyl alcohol	H-3, F-3	Critical	10,000-99,999
108-88-3	Toluene	H-3, F-3	Critical	10,000-99,999
1330-20-7	Xylene	H-3	Critical	10,000-99,999
121-44-8	Triethylamine	H-3, F-3	Critical	1,000-9,999

NFPA: National Fire Protection Association.

H: Health Hazard Rating (0 to 4 Scale). F: Flammability Hazard Rating (0 to 4 Scale).

HSC: ITF-40 Hazard Severity Classification.

### *Release modeling*

Figure 12 is a map identifying the effects of a flammable release of xylene using ALOHA. The estimation was based on a reportable limit of 500 kg. The TIC-Master program does not have the ability to specify storage vessels.

Presumably, it would be possible for military intelligence to obtain storage vessel information. However, that was not provided in this case. ALOHA does, however, list likely storage vessels depending on the contaminant. This feature, as well as information on the density of xylene, was used to estimate that xylene was stored in a 200-gallon spherical tank. The explosive release included the weather conditions given in the start conditions.

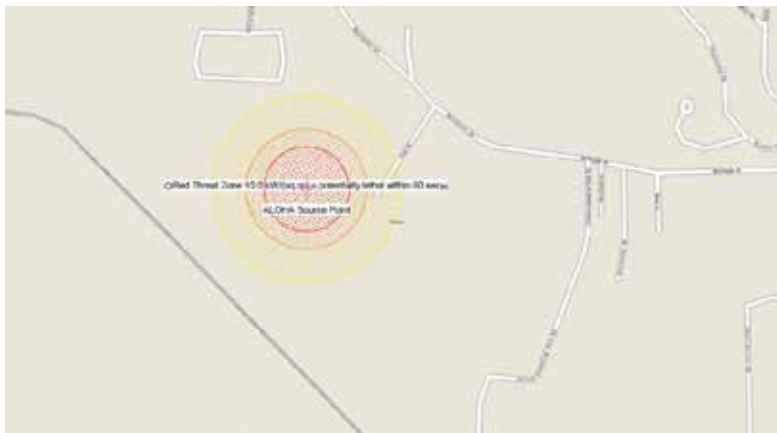


Figure 12. Release modeling of an explosion of a 200-gal tank of xylene. The innermost red circle indicates lethality within 60 seconds, and was about 200 m in diameter. The circular pattern is common for explosions.

As seen, the explosive release from a circular tank creates a circular area of influence that is relatively uninfluenced by the wind. The zone of impact was about 500 m in diameter, with a 200-m zone indicating lethality within 60 seconds. It should be possible to model secondary contaminant release following the explosion, if that is needed; however, it would be reasonable to assume that the contaminants would be largely consumed in the explosion.

#### *Evaluation of chemical risks from ITF-40 63 most toxic chemicals*

Three chemicals were on the SOP list: Acrylonitrile, formaldehyde, and propylene oxide.

#### *Acrylonitrile*

Fate – Biodegrades in soil.

Mitigation – Limit movement in soil and surface water and promote soil biodegradation.

### *Formaldehyde*

Risks – Because it readily biodegrades in soil, the primary issue is air. May also contaminate water (MCL = 0.7 ug/L).

Mitigation – Promote biodegradation in soil and groundwater. Water-based sprays for air and absorbents for soil spills.

### *Propylene oxide*

Risks – Photochemically reactive in atmosphere. Doesn't readily sorb to soil, so may migrate, but also undergoes hydrolysis reaction (reaction with water).

Mitigation – Limit spread in soils and water. Use water mists to remove from air.

### *Issues*

Several issues were identified in this first test. First, it was hoped that all test team members would have access to the program, as the Army has a general license to ArcGIS. However, the license covers only the Reader version of the program, not the Editor version. The Reader version alone would not run the program, which limited the GIS portion of the work to Dhiren Khona.

Initially there were problems with automating the process of ranking the most probable contaminants. However, once this process was straightened out, the ranking system was found to be very accurate.

Ranking of the most hazardous materials was challenging. Initially, there was no search mechanism, which required manually searching the database for hazard information. Fortunately, this deficiency was rectified. Choosing criteria for hazard ranking would prove to be more challenging. Flammability and reactivity could be critical for certain applications. However, health and toxicity are also crucial. Issues regarding factors such as the likelihood of the presence or potential quantity of hazardous materials would continue to be dealt with throughout subsequent tests.

## Test 2 Summary

### Scenario

A fire has occurred at a soybean processing facility. Six inches of rain has fallen within the last 8 hr. There are other flammable liquids on site. The scenario form is given as Figure 13.

Test <u>2</u> Scenario		
Date: <u>10/16/09</u> Time Submitted: <u>0915 CDT</u> Time Team Response Received: <u>09:30</u>		
Team Initial Start Time: <u>09:45</u> Time Team Response Submitted: <u>14:15</u>		
Completed: <u>11:20</u>		
<u>Industrial Processes or Process Name:</u>  Soybean Processing	<u>TIC Release Mechanism:</u>  The fire which occurred is probably due to a lightning strike during a thunderstorm.	<u>Atmospheric conditions:</u> Temperature: 72° F Wind Speed: 5 mph Wind Direction: SE Precipitation: 6 in. within the last 8 hr
<u>Suspected TICs:</u>	<u>Other Industrial Chemicals:</u>	<u>Other Potential Hazards:</u>  Flammable liquids are on site. Some above-ground tanks have been destroyed.
<u>Site Designation:</u> Single Site: <u>X</u> Multiple Sites: <u>   </u> General Location*: <u>   </u>	<u>Site A*:</u> Longitude: 90° 53' 29.99" W Latitude: 32° 23' 30.43" N Elevation:	<u>Site B:</u> Longitude: Latitude: Elevation:
<u>Other Characteristics:</u>		
* See the <u>Other Characteristics</u> section. For example, the U.S. Army may wish to occupy a 2-square-mile area centered at the location provided by Site A*.		

Figure 13. Scenario description for Test 2.

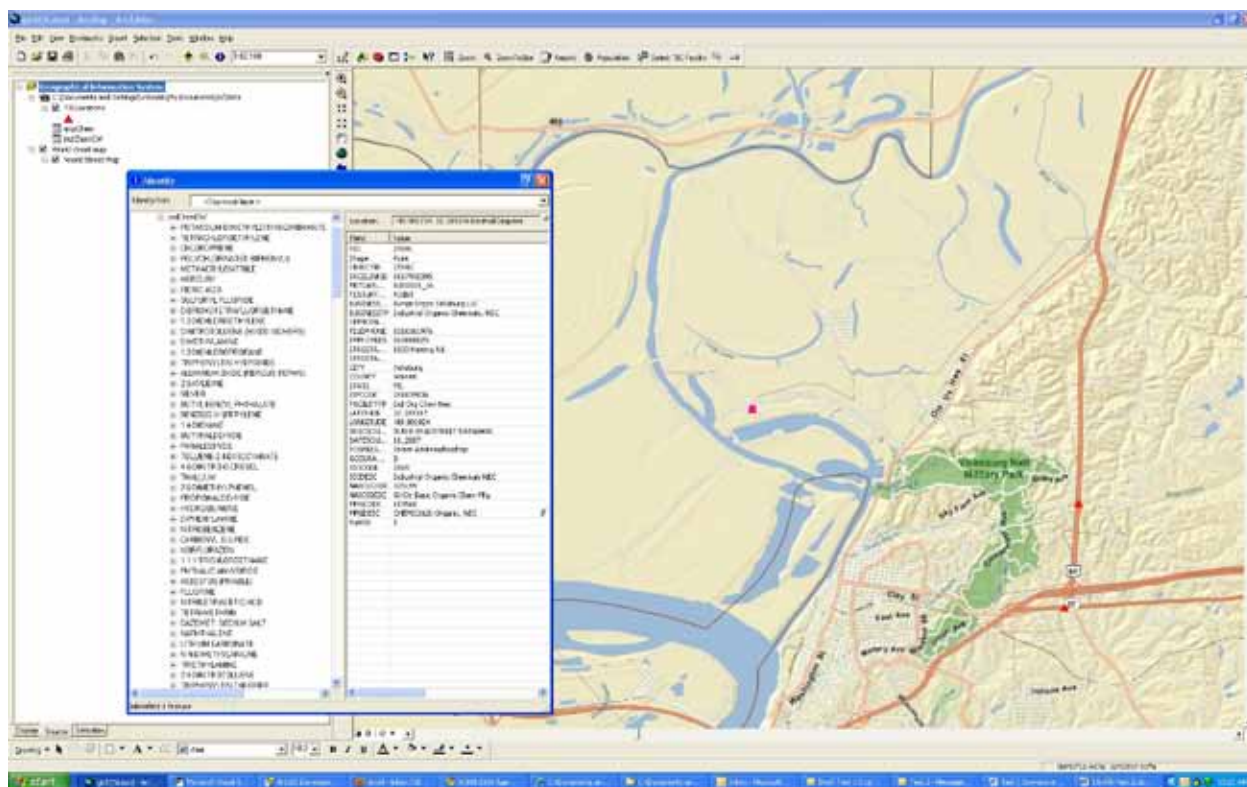
## Results

### Response time

The response time was 1 hr 35 minutes. This response time is consistent with the test team's goal of providing a response within 2 hr.

### Facility location and identification

Figure 14 is a location map for the facility. The facility and code NAICS 311222 were correctly identified as soybean processing. The actual facility was the Bunge-Ergon ethanol production facility in Vicksburg, MS.



**Figure 14. Location map for Test 2. The pink dot is the location of the facility – in the Port of Vicksburg.**

### Chemical identification

The database identified 26 TICs (Table 10). These included hydrocarbons, alcohols, ketones, metals, and toxic gases.

The 2009 Tier II report (reporting from 1 January to 31 December 2008) submitted by Bunge-Ergon Vicksburg, LLC was used to verify the chemicals identified by the test database. This report listed five chemicals, four of which were in the test database. The database correctly identified two of

Table 10. Estimated chemical list for Test 2.

CHEMICAL
BENZO(G H I)PERYLENE
N-HEXANE
LEAD
MERCURY
SULFURIC ACID
SODIUM HYDROXIDE (SOLUTION)
PHOSPHORIC ACID
METHANOL
ACETONE
ZINC (FUME OR DUST)
NITRIC ACID
HYDROCHLORIC ACID
NICKEL
CHLORINE
AMMONIA
NAPHTHALENE
BIPHENYL
PROPYLENE
ASBESTOS (FRIABLE)
POLYCHLORINATED BIPHENYLS
MANGANESE
ARSENIC
COPPER
SODIUM NITRITE
SELENIUM
AMMONIUM SULFATE (SOLUTION)

these four chemicals (Table 11). Of the remaining two chemicals, the database identified very similar chemicals – ammonia vs. aqueous ammonia and methanol vs. ethanol (Table 12).

Methanol and ethanol have very similar hazard characteristics. However, comparing ammonia to aqueous ammonia reveals some subtle, but possibly important, differences in hazard assessment, with the aqueous ammonia having a greater effect in terms of health, but the gaseous form having a slight flammability.

Table 11. Correctly identified chemicals for Test 2 simulation.

CAS	Chemical Name	Inventory (Max Daily lbs)	Inventory (Ave Daily lbs)
7664-93-9	Sulfuric Acid	100,000-999,999	10,000-99,999
1310-73-2	Sodium Hydroxide	100,000-999,999	100,000-999,999

Table 12. Comparison of database predicted chemicals with actual chemicals.

A. Database									
CAS	Chemical Name	State	Exposure Route	NFPA H/F/I			ITF40 (HSC/HPC/RR)		
7664-41-7	Ammonia	G	Inhalation	3	1	0	Negligible	Frequent	Moderate
67-56-1	Methanol	L	Oral	3	3	0	Critical	Likely	High
B. Actual									
CAS	Chemical Name	State	Exposure Route	NFPA H/F/I			ITF40 (HSC/HPC/RR)		
1336-21-6	Aqueous Ammonia	L	Oral	4			Catastrophic	Occasional	High
64-17-5	Ethanol	L	Oral	2	3	0	Critical	Likely	High

### *Most probable chemicals*

Table 13 summarizes the chemicals predicted to be most likely at the site. Two of these, sulfuric acid and sodium hydroxide, were found at the site. These were the second and third most likely chemicals, respectively. Ammonia was the eighth most likely chemical. As discussed in the section above, this chemical was similar, but not identical to, aqueous ammonia. The most likely chemical was n-hexane, which was actually not in the EPA report. However, this solvent chemical would likely serve the same role as ethanol, which was identified at the site.

Table 13. Ten most likely chemicals for Test 2.

%prob	CAS	Name
76.19	110543	N-HEXANE
38.10	7664939	SULFURIC ACID
26.19	1310732	SODIUM HYDROXIDE (SOLUTION)
25.00	7664382	PHOSPHORIC ACID
19.05	7647010	HYDROCHLORIC ACID
14.29	7440020	NICKEL
13.10	7782505	CHLORINE
11.90	7664417	AMMONIA
8.33	191242	BENZO(G H I)PERYLENE
7.14	7439921	LEAD

### *Most significant chemical hazards*

The test team ranked the chemicals based on fire and flammability. Four chemicals in the original 26 chemical list had flammability rankings of 3 or higher (Table 14). The database was also used to rank the chemicals on toxicity (NFPA Health) (Table 15). Sulfuric acid (high health rating), ammonia (3 in health and 1 in flammability), and methanol (3 in health and 3 in flammability) were the most significant hazards identified by the test team. However, a good case could be made for several other chemicals.

**Table 14. Ranking of Test 2 chemicals on fire/flammability: chemicals with NFPA-F ratings of 3 or higher.**

NFPA F	CAS	Chemical Name	%Prob
4	115071	PROPYLENE	1.19
3	110543	N-HEXANE	76.19
3	67561	METHANOL	2.38
3	67641	ACETONE	2.38

**Table 15. Ranking of Test 2 chemicals on toxicity/health: Chemicals with NFPA-H ratings of 3 or higher.**

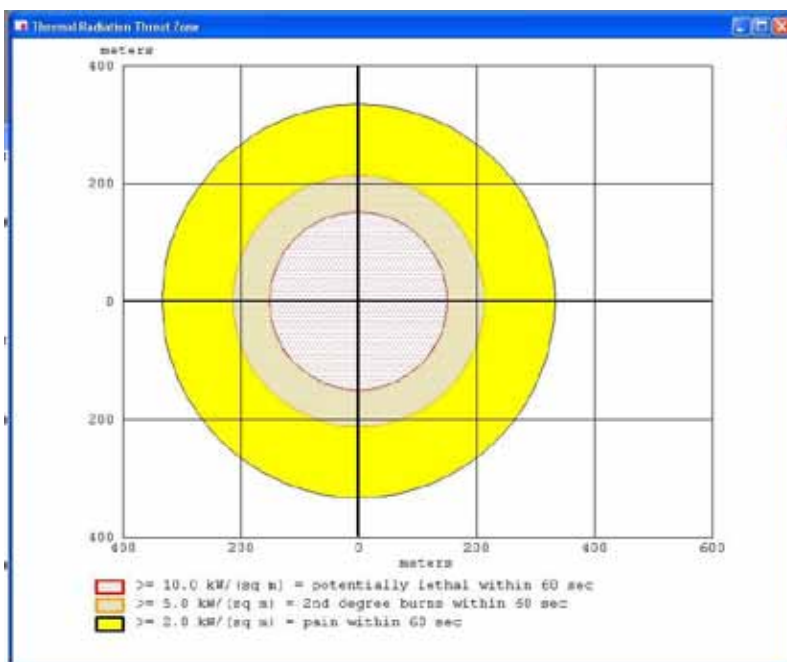
NFPA H	CAS	Name	%Prob
4	7664939	SULFURIC ACID	38.10
4	7647010	HYDROCHLORIC ACID	19.05
4	7782505	CHLORINE	13.10
4	7439976	MERCURY	4.76
4	7632000	SODIUM NITRITE	1.19
3	7664417	AMMONIA	11.90
3	67561	METHANOL	2.38
3	91203	NAPHTHALENE	1.19
3	92524	BIPHENYL	1.19

### *Release modeling*

A tank explosion was conducted using propylene, a gas with an NFPA flammability rating of “4.” Explosion of a 4000-gal cylindrical tank was assumed, with a diameter of 10 ft and a pressure of 75 psig. A relative humidity of 50% and a single story, metal building were also assumed.

ALOHA estimated a circular hazard pattern based on explosive energy (kwatts/m<sup>2</sup>) (Figure 15a). The diameter of the circle of impact (lethal within 60 seconds) was 400 m. In addition, the program estimated that a 30-m-diam fireball would be generated. Such an explosion would impact a significant portion of the harbor area and would potentially cut off traffic on the roads servicing the area (Figure 15b).

a. Size diagram of modeled explosion of a 4000-gal propylene tank.



b. Explosion of a 4000-gal propylene tank for the facility in Test 2.



Figure 15. Release modeling of an explosion of a propylene tank for Test 2.

*SOPs from ITF-40 63 most significant list*

Only one chemical was found with an SOP, hydrogen chloride, with this information derived from the SOP:

Hydrogen chloride – Its major concern is depression of pH, creating acidic issues. It can cause severe burns. When diluted, it may not be a great hazard, but could mobilize metals. It can be neutralized, but care must be used because strong bases can explosively react with strong acids

**Test 3 Summary****Scenario**

The Test 3 scenario assumes multiple sites or industrial processes exist within a 1-square-mile area. At least one railroad car tank containing 7664-93-9 sulfuric acid was hit by a mortar shell. Surface soil modeling of 7664-93-9 is requested. Figure 16 is the submission sheet for the test.

**Results***Response time*

Due to sickness and differences in schedules, the team worked separately in increments of 1.5 hr, 1.5 hr, and 1 hr, respectively. It probably would have taken approximately 2.5 hr of labor if the team had worked together. This response time exceeded the goal of providing a respond within 2 hr.

*Facility identification*

The NAICS code: 336111 (automotive manufacturing) and 336112 (light truck and utility vehicle manufacturing) were correct. The test team's database only contained information for 336111, automotive manufacturing. However, it was reasonable to assume that the processes for automotive manufacturing and light truck manufacturing would be about the same. The 2008 Tier II report submitted by the Nissan North America, Inc., Canton Manufacturing Facility was used to verify the chemicals identified by the test team.

Test <u>3</u> Scenario		
Date: <u>10/26/09</u> Time Submitted: <u>0915 CDT</u> Time Team Response Received: _____		
Team Initial Start Time: _____ Time Team Response Submitted: _____		
<u>Industrial Processes or Process Name:</u>  NAICS: 336111 and 336112	<u>TIC Release Mechanism:</u> The enemy has vacated this area after a fierce aerial assault. This assault resulted in numerous fires at least 8 hours ago.	<u>Atmospheric conditions:</u> Temperature: 100 ° F Wind Speed: 15 Wind Direction: SW Precipitation: None
<u>Suspected TICs:</u> At least one railroad car tank containing 7664-93-9 sulfuric acid was hit by a mortar shell. Surface soil modeling of 7664-93-9 is requested.	<u>Other Industrial Chemicals:</u>	<u>Other Potential Hazards:</u>  Flammable liquids are on site. Drought conditions exist.
<u>Site Designation:</u> Single Site: _____ Multiple Sites: _____ General Location*: <u>X</u>	Site A*: Longitude: 90° 04' 45.69" W Latitude: 32° 35' 17.42" N Elevation:	Site B: Longitude: Latitude: Elevation:
<u>Other Characteristics:</u> Assume multiple sites or industrial processes exist within a 1-square-mile area. The area of concern includes a half-mile distance in all four directions from the coordinates provided above.		
* See the <u>Other Characteristics</u> section. For example, the U.S. Army may wish to occupy a 2-square-mile area centered at the location provided by Site A*.		

Figure 16. Scenario description for Test 3.

Figure 17 is a location map. The skull and crossbones symbol marks the location of the facility located by the test team. However, the facility was apparently mislocated due to a discrepancy of the longitude and latitude provided. Perhaps the location of an office was provided, although technically the plant location should be reported in the Tier II report. The pink dot shows the actual location of the plant.



Figure 17. Location map for Test 3. Skull and crossbones symbol marks the identified location. The pink dot is the actual location.

#### *Chemical identification*

The test team identified 91 potential TICs (Table 16). Fifty-three chemicals were listed on the Tier II report from Nissan. Of the 53, 30 were actually in the program database (Table 17). The database correctly identified 15 of the 30 remaining chemicals (Table 18).

#### *Ten most likely chemicals*

The test team identified the 10 most probable chemicals (Table 19), six of which were included in the Nissan Tier II report.

#### *Most significant chemical hazards*

Tables 20 and 21 show the highest-rated chemical hazards for fire/flammability and for toxicity/health for the chemicals identified in Table 19. The test team chose the top three risks to be sulfuric acid (4 health risk and identified by the scenario as a high probability of release), benzene (4 health risk, and 3 flammability), and vinylidene chloride (4 fire risk, 3 health) (Table 22). Strong cases could be made for other chemicals as well.

Table 16. Estimated chemical list for Test 3.

CAS	Name	%probability	NFFA F	NFFA H	history	TP40 priority	Notes
50000	FORMALDEHYDE	15.75	2	3	3	y	
51796	URETHANE	3.75	2	2	2		
54198	FORMIC ACID	5.00	1	2	2		
67551	METHANOL	47.50	3	3	3		
67530	ISOPROPYL ALCOHOL (MAI	8.25	3	2	3		
67541	ACETONE	42.50	3	2	3		
71353	N-BUTYL ALCOHOL	47.50	3	2	3		
71432	BENZENE	43.75	3	4	3		
71556	1,1,1-TRICHLOROETHANE	22.50	1	2	2		
74573	CHLOROMETHANE	1.25	4	2	2		
75014	VINYL CHLORIDE	2.50	4	2	3	y	
75092	DICHLOROMETHANE	33.75	1	3	3		
75354	VINYLDENE CHLORIDE	1.25	4	3	3		
75550	TERT-BUTYL ALCOHOL	1.25	3	2	3		
75718	DICHLORODIFLUOROMETH	37.50			1		
78131	Freon 113	5.00					
78922	SEC-BUTYL ALCOHOL	3.75	3	2	2		
78938	METHYL ETHYL KETONE	52.50	3	2	2		
79015	TRICHLOROETHYLENE	5.00	1	1	3		
80057	4,4'-ISOPROPYLDENEDIPH	10.00	3	2	2		
80525	METHYL METHACRYLATE	1.25	3	2	2		
84552	DIETHYL PHTHALATE	1.25	1	2	2		
84742	DIBUTYL PHTHALATE	5.00	1	2	2		
85537	BUTYL BENZYL PHTHALATE	5.75	1	2	2		
91203	NAPHTHALENE	22.50	2	3	3		
92524	BIPHENYL	1.25	1	3	3		
94380	BENZOYL PEROXIDE	1.25					OK
95475	O-XYLENE	1.25	3	1	1		
95487	O-CRESOL	1.25	2	3	3		
95535	1,2,4-TRIMETHYLBENZENE	45.00	3	1	1		
96525	CUMENE	21.25	3	2	2		
100414	ETHYLBENZENE	48.75	3	2	2		
100425	STYRENE	5.75	3	3	3		
101536	METHYLENEBIS(PHENYL)	11.25					not in data
103231	BIS(2-ETHYLHEXYL) ADIPAT	5.00	1	1	1		
105579	2,4-DIMETHYLPHENOL	1.25					
106445	P-CRESOL	1.25	2	3	3		
106596	EPICHLOROHYDRIN	1.25	3	3	3		
107211	ETHYLENEGLYCOL	45.25	1	3	3		
108101	METHYL ISOBUTYL KETON	51.25	3	2	1		
108333	M-XYLENE	1.25	3	1	1		
108394	M-CRESOL	1.25	2	2	2		
108553	TOLUENE	55.75	3	3	3		
108952	PHENOL	1.25	2	4	3	y	
109384	2-METHOXYETHANOL	2.50	2	2	2		
110543	N-HEXANE	33.75	3	1	3		
110605	2-ETHOXYETHANOL	8.25	2	2	2		
110627	CYCLOHEXANE	40.00	3	2	3		
111422	DIETHANOLAMINE	3.75	1	2	2		
115071	PROPYLENE	7.50	4	1	3	y	
117517	DI(2-ETHYLHEXYL) PHTHAL	12.50		2	1		
117840	NDIOCTYL PHTHALATE	3.75	1	1	1		
120321	1,2,4-TRICHLOROBENZENE	1.25	1	3	3		
122394	DIPHENYLAMINE	1.25		3	3		
123911	1,4-DIOXANE	2.50	3	2	2		
127164	TETRACHLOROETHYLENE	2.50	0	2	2		
128030	POTASSIUM DIMETHYLDITHIO	1.25			3		
128041	SODIUM DIMETHYLDITHIOK	1.25			3		
131113	DIMETHYL PHTHALATE	2.50	1	2	2		
141322	BUTYL ACRYLATE	3.75	2	3	3		
159242	BENZO(G,H)PERYLENE	22.50					
872504	N-METHYL-2-PYRROLIDON	27.50	2	2	2		
1153195	DECABROMODIPHENYL OX	3.75					
1310732	SODIUM HYDROXIDE (SOL)	37.50	0	2	3		
1319773	CRESOL (MIXED ISOMERS)	1.25		3	3		
1330207	XYLENE (MIXED ISOMERS)	55.25	3	2	3		NFFA code
1344281	ALUMINUM OXIDE (FIBROU	1.25					
1534044	METHYL TERT-BUTYL ETH	35.25		2	3		
7429905	ALUMINUM (FUME OR DUST	7.50					
7439921	LEAD	20.00					
7439955	MANGANESE	13.75					
7439978	MERCURY	1.25		4			
7440020	NICKEL	22.50					
7440224	SILVER	1.25					

Table 17. Chemicals listed on Nissan Tier II report and the TIC-Master database.

CAS_Text	Chemical_Name	State	Inventory (Max Daily lb)	Inventory (Ave Daily lb)
7664-93-9	Sulfuric acid	L	10,000-99,999	10,000-99,999
107-21-1	Ethane-1,2-diol	L	100,000-999,999	100,000-999,999
75-65-0	2-methylpropan-2-ol	L	10,000-99,999	10,000-99,999
1310-73-2	Sodium hydroxide	S	10,000-99,999	10,000-99,999
8008-20-6	Kerosene	L	10,000-99,999	10,000-99,999
763-69-9	Ethyl 3-ethoxypropionate	L	10,000-99,999	10,000-99,999
7705-08-0	Iron trichloride	L	100,000-999,999	100,000-999,999
67-63-0	Propan-2-ol	L	10,000-99,999	10,000-99,999
78-93-3	Butanone	L	1,000-9,999	1,000-9,999
123-86-4	n-butyl acetate	L	100,000-999,999	100,000-999,999
811-97-2	1,1,1,2-Tetrafluoroethane	G	10,000-99,999	10,000-99,999
5131-66-8	2-Propanol, 1-butoxy-(PGnBE)	L	100,000-999,999	100,000-999,999
67-56-1	Methanol	L	10,000-99,999	10,000-99,999
1305-78-8	Calcium oxide	S	100,000-999,999	10,000-99,999
497-19-8	Sodium carbonate	S	10,000-99,999	1,000-9,999
112-80-1	Oleic acid	L	10,000-99,999	10,000-99,999
1310-58-3	Potassium hydroxide	S	10,000-99,999	10,000-99,999
7697-37-2	Nitric acid	L	10,000-99,999	10,000-99,999
95-63-6	1,2,4-trimethylbenzene	L		
112-25-4	2-hexyloxyethanol	L	10,000-99,999	10,000-99,999
7632-00-0	Sodium nitrite	S	10,000-99,999	10,000-99,999
7631-99-4	Sodium nitrate	S	10,000-99,999	10,000-99,999
108-10-1	4-methylpentan-2-one	L	10,000-99,999	10,000-99,999
124-38-9	Carbon dioxide	G	10,000-99,999	10,000-99,999
7727-37-9	Nitrogen, compressed	G	10,000-99,999	10,000-99,999
6846-50-0	1-isopropyl-2,2-dimethyltrimethylene diisobutyrate			
67-64-1	Acetone	L	1,000-9,999	1,000-9,999
7647-01-0	Hydrogen chloride	G	100,000-999,999	10,000-99,999
872-50-4	1-methyl-2-pyrrolidone	L	10,000-99,999	10,000-99,999
306-83-2	Ethane, 2,2-dichloro-1,1,1-trifluoro-	L	10,000-99,999	10,000-99,999

blue font : Chemicals on the ITF-40 Top 63 list

State Symbols: S = Solid, L = Liquid, G = Gas.

**Table 18. Fifteen chemicals identified by the test team appear in the Tier II report and TIC-Master database.**

CAS_Text	Chemical Identified by the Test Team
67-56-1	Methanol
67-63-0	Propan-2-ol
67-64-1	Acetone
71-36-3	Butan-1-ol (n-butyl alcohol)
75-65-0	2-methylpropan-2-ol
78-93-3	Butanone (methyl ethyl ketone)
95-63-6	1,2,4-trimethylbenzene
107-21-1	Ethane-1,2-diol (ethylene glycol)
108-10-1	4-methylpentan-2-one (methyl isobutyl ketone)
872-50-4	1-methyl-2-pyrrolidone
1310-73-2	Sodium hydroxide
7632-00-0	Sodium nitrite
7647-01-0	Hydrogen chloride
7664-93-9	Sulfuric acid
7697-37-2	Nitric acid

**Table 19. Most probable chemicals for Test 3. Starred Items were included in the actual Tier II report.**

%prob	CAS	Name
58.75	108883	TOLUENE
56.25	1330207	XYLENE (MIXED ISOMERS)
52.50	78933	METHYL ETHYL KETONE
51.25	108101	METHYL ISOBUTYL KETONE
48.75	100414	ETHYLBENZENE
47.50	67561	METHANOL
47.50	71363	N-BUTYL ALCOHOL
47.50	7664939	SULFURIC ACID
46.25	107211	ETHYLENE GLYCOL
46.25	7664382	PHOSPHORIC ACID

Note: 37 chemicals with 10% or greater probability

Table 20. Chemical hazard list for Test 3 based on fire/flammability (NFPA-F). All chemicals with rankings of 3 or greater.

NFPA F	CAS #	CHEMICAL NAME	%Prob	NFPA H
4	115071	PROPYLENE	7.50	1
4	75014	VINYL CHLORIDE	2.50	2
4	74873	CHLOROMETHANE	1.25	2
4	75354	VINYLDENE CHLORIDE	1.25	3
3	108883	TOLUENE	58.75	3
3	1330207	XYLENE (MIXED ISOMERS)	56.25	2
3	78933	METHYL ETHYL KETONE	52.50	2
3	108101	METHYL ISOBUTYL KETONE	51.25	2
3	100414	ETHYLBENZENE	48.75	2
3	67561	METHANOL	47.50	3
3	71363	N-BUTYL ALCOHOL	47.50	2
3	71432	BENZENE	43.75	4
3	67641	ACETONE	42.50	2
3	110827	CYCLOHEXANE	40.00	2
3	110543	N-HEXANE	33.75	1
3	98828	CUMENE	21.25	2
3	100425	STYRENE	8.75	3
3	67630	ISOPROPYL ALCOHOL	6.25	2
3	78922	SEC-BUTYL ALCOHOL	3.75	2
3	123911	1,4-DIOXANE	2.50	2
3	75650	TERT-BUTYL ALCOHOL	1.25	2
3	80626	METHYL METHACRYLATE	1.25	2
3	95476	O-XYLENE	1.25	1
3	106898	EPICHLOROHYDRIN	1.25	3
3	108383	M-XYLENE	1.25	1

Table 21. Chemical hazard list for Test 3 based on toxicity/health (NFPA-H). All chemicals with rankings of 3 or greater.

NFPA H	CAS	Chemical Name	% prob	NFPA F
4	7664939	SULFURIC ACID	47.50	0
4	71432	BENZENE	43.75	3
4	7632000	SODIUM NITRITE	43.75	i
4	7697372	NITRIC ACID	41.25	0
4	7647010	HYDROCHLORIC ACID	35.00	0
4	7782505	CHLORINE	6.25	0
4	7723140	PHOSPHORUS (YELLOW OR WHITE)	5.00	i
4	108952	PHENOL	1.25	2
4	7439976	MERCURY	1.25	i
4	7440439	CADMIUM	1.25	i
4	7664393	HYDROGEN FLUORIDE	1.25	0
3	108883	TOLUENE	58.75	3
3	67561	METHANOL	47.50	3
3	107211	ETHYLENE GLYCOL	46.25	1
3	75092	DICHLOROMETHANE	33.75	1
3	91203	NAPHTHALENE	22.50	2
3	50000	FORMALDEHYDE	18.75	2
3	100425	STYRENE	8.75	3
3	7664417	AMMONIA	5.00	1
3	141322	BUTYL ACRYLATE	3.75	2
3	75354	VINYLDENE CHLORIDE	1.25	4
3	92524	BIPHENYL	1.25	1
3	95487	O-CRESOL	1.25	2
3	106445	P-CRESOL	1.25	2
3	106898	EPICHLOROHYDRIN	1.25	3
3	120821	1 2 4-TRICHLOROBENZENE	1.25	1
3	122394	DIPHENYLAMINE	1.25	i
3	128041	SODIUM DIMETHYLDITHIOCARBAMATE	1.25	i
3	1319773	CRESOL (MIXED ISOMERS)	1.25	i

Table 22. Three most significant chemical hazards selected by the test team.

Chemicals	Health	Fire
Sulfuric Acid	4	0
Benzene	4	3
Vinylidene Chloride	4	3

### *Release modeling*

The scenario stated “At least one railroad car tank containing 7664-93-9 sulfuric acid was hit by a mortar shell. Surface soil modeling of 7664-93-9 is requested.” In this case, the ALOHA model’s limitations prevented modeling of the scenario presented. First, ALOHA does not include surface release modeling. However, it does allow for modeling of air contamination from a surface release; therefore efforts were focused on modeling the atmospheric effects. Secondly, ALOHA does not include sulfuric acid as one of its chemicals. The available acids were reviewed, including nitric acid, hydrochloric acid, hydrofluoric acid, and sulfurous acid. Sulfurous acid was investigated and found to be different than sulfuric acid in chemical behavior. Sulfurous acid also required inputs that were not available. HF acid, which is generally considered the worst of the strong acids, modeled as a worst-case scenario.

The model focused on the air effects of a release of 100 gal of HF, substantially less than a railcar. This was an oversight by the response team. The modeling assumed that the release was from 2 x 55 gallon drums spilled on the ground. The release included the weather conditions given in the start conditions. As seen, the surface spill release from a circular tank creates a plume-shaped area of influence that is greatly impacted by the wind (Figure 18a).

Figure 18b shows the modeled plume. The plume was rather narrow in diameter, less than 0.1 km. The 44-ppm portion of the plume was about 0.1 km long. However, the 1-ppm portion of the plume, which is great enough to affect the eyes and respiratory system, extended for a length of 1 km.

a. Map of the release of HF from the plant in Test 2.



b. Model of the size of the release of HF.

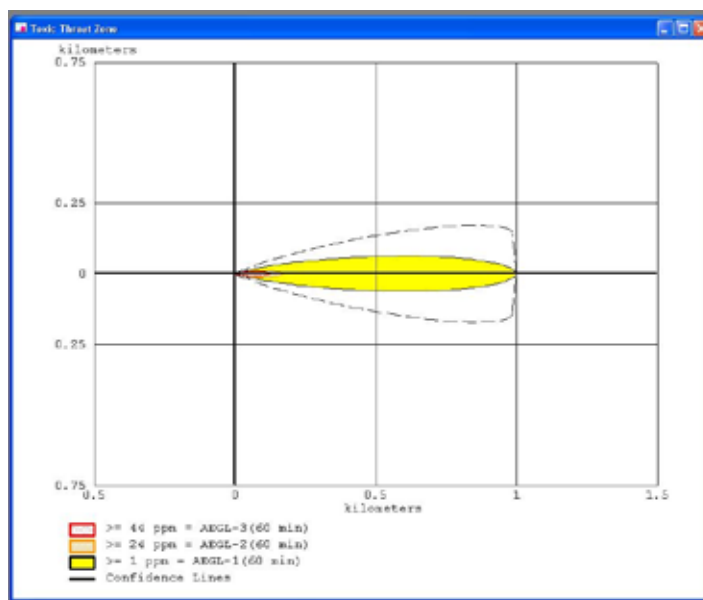


Figure 18. Release modeling supporting Test 3.

## Test 4 Summary

### Scenario

Coordinated suicide bombers detonated trucks at two facilities. Massive fires ensued. A partial chemical list is available for each facility. The commander needs to address community concerns about hydrogen sulfide (CAS # 7783-06-4) in a nearby waterway. Figure 19 is the test scenario submission sheet.

Test <u>4</u> Scenario		
Date: <u>11/2/09</u> Time Submitted: <u>0945 CST</u> Time Team Response Received: _____		
Team Initial Start Time: _____ Time Team Response Submitted: _____		
We estimate our working time on this to be on the order of 2.75 hours.		
<u>Industrial Processes or Process Name:</u> Site A: Petroleum Refineries Site B: Polish and Other Sanitation Good Manufacturing	<u>TIC Release Mechanism:</u> Coordinated suicide bombers drove trucks into each facility and detonated them 12 hours ago. Massive fires ensued.	<u>Atmospheric conditions:</u> Temperature: 24 deg F Wind Speed: 10 mph Wind Direction: East Precipitation: 0.25 inch per hour
<u>TICs:</u> Site A: Chemicals include 007647-01-0 Muriatic acid, 001310-73-2 Caustic, 8002-26-4 Distilled Toll Oil, and 7783-06-4 Hydrogen sulfide.  <b>Note: This is only a partial chemical list.</b>  Site B: Chemicals include 7664-93-9 Battery acid, 68608-26-4 Hostapur, 7439-92-1 Lead, 57-55-6 Propylene glycol, 1310-73-2 Sodium hydroxide (Caustic), and 67-63-0 Isopropyl alcohol-anhydrous.  <b>Note: This is only a partial chemical list.</b>		<u>Other Potential Hazards:</u> Site A: Other potential hazards on site include 64741-44-2 diesel, 8008-20-6 Kerosene, and unleaded gasoline.  Site B: Other potential hazards on site include 65-85-0 Benzoic acid, 8002-09-3 Terpene alcohol/hydrocarbons, and original pine sol.
<u>Site Designation:</u> Single Site: _____ Multiple Sites: <u>X</u> General Location*: _____	Site A*: Longitude: 90° 54' 21.20" W Latitude: 32° 23 19.63" N Elevation: _____	Site B: Longitude: 90° 09' 26.06" W Latitude: 32° 17' 03.09" N Elevation: _____
<u>Other Characteristics:</u> Site A: Can you address community concerns about 7783-06-4 Hydrogen sulfide concentrations in a nearby waterway?		
* See the <u>Other Characteristics</u> section. For example, the U.S. Army may wish to occupy a 2-square-mile area centered at the location provided by Site A*.		

Figure 19. Scenario description for Test 4.

## Results

### *Response time*

The estimated working time was 2.75 hr. This response time exceeded the goal of 2 hr.

### *Facility identification*

Site A referred to Ergon Refining, Inc., and was identified as a chemical refining industry. Site B referred to the Clorox Products Manufacturing Co. NAICS: 325612 and was identified by the TIC-Master database as cleaning and sanitation products manufacturing. Figure 20 is a site location map. The red date refers to Site A. Site B was located by the skull and crossbones symbol.



Figure 20. Location map for Test 4. The pink dot is Site A. Skull and crossbones was Site B.

### *Chemical identification*

A total of 160 possible chemicals were identified for the chemical refining operation (Table 23). Forty-four chemicals were actually identified in the EPA Tier II report for Ergon Refining, Inc. Of these, 11 were TICs found in the TIC-Master database. Four of these 11 were identified by the TIC-Master database. Tables 24 and 25 show the TICs identified and not identified, respectively.

Table 23. Full listing of chemicals predicted for chemical refining, with the first 30 ranked.

CHEMICAL	%Prob	Ranking	CHEMICAL	%Prob	Ranking
TOLUENE	79.65	1	1 2-DICHLOROETHANE	20.00	
BENZENE	78.95	2	CHROMIUM	19.30	
XYLENE (MIXED ISOMERS)	76.84	3	BIPHENYL	16.49	
ETHYLBENZENE	73.33	4	METHYL ETHYL KETONE	16.14	
CYCLOHEXANE	69.82	5	ANTHRACENE	15.09	
1 2 4-TRIMETHYLBENZENE	64.91	6	O-XYLENE	14.74	
NAPHTHALENE	62.11	7	PHENANTHRENE	14.39	
SULFURIC ACID	61.75	8	1 2-DIBROMOETHANE	14.39	
N-HEXANE	59.30	9	HYDROGEN CYANIDE	13.33	
SODIUM HYDROXIDE (SOLUTION)	58.25	10	P-XYLENE	13.33	
AMMONIA	55.44	11	M-XYLENE	13.33	
CHLORINE	53.68	12	CARBON TETRACHLORIDE	12.98	
PROPYLENE	50.18	13	2-METHOXYETHANOL	12.28	
BENZO(G H I)PERYLENE	48.77	14	VANADIUM (FUME OR DUST)	11.93	
METHYL TERT-BUTYL ETHER	48.42	15	COBALT	11.23	
ALUMINUM OXIDE (FIBROUS FORMS)	47.37	16	MANGANESE	10.88	
ETHYLENE	47.02	17	TERT-BUTYL ALCOHOL	9.47	
CUMENE	47.02	18	FORMALDEHYDE	8.07	
METHANOL	44.91	19	N-BUTYL ALCOHOL	8.07	
MOLYBDENUM TRIOXIDE	42.81	20	COPPER	8.07	
1 3-BUTADIENE	40.35	21	ASBESTOS (FRIABLE)	7.72	
PHENOL	39.30	22	BARIUM	7.72	
HYDROCHLORIC ACID	34.04	23	ZINC (FUME OR DUST)	7.37	
DIETHANOLAMINE	33.33	24	ANTIMONY	6.67	
PHOSPHORIC ACID	31.93	25	2 4-DIMETHYLPHENOL	6.32	
TETRACHLOROETHYLENE	31.23	26	ACETONE	5.61	
LEAD	30.88	27	TRICHLOROETHYLENE	4.91	
CARBONYL SULFIDE	28.07	28	METHYL ISOBUTYL KETONE	4.91	
CARBON DISULFIDE	26.67	29	CHLORODIFLUOROMETHANE	4.56	
ETHYLENE GLYCOL	26.32	30	DICHLORODIFLUOROMETHANE	4.21	
1 1 1-TRICHLOROETHANE	24.91		POLYCHLORINATED BIPHENYLS	4.21	
NICKEL	23.16		ARSENIC	4.21	
CRESOL (MIXED ISOMERS)	22.81		DICYCLOPENTADIENE	3.86	
HYDROGEN FLUORIDE	22.81		N-METHYL-2-PYRROLIDONE	3.86	
MERCURY	20.70		CADMIUM	3.86	
STYRENE	20.35		SELENIUM	3.86	

CHEMICAL	%Prob	Ranking	CHEMICAL	%Prob	Ranking
SODIUM NITRITE	3.51		ACRYLIC ACID	0.70	
DICHLOROMETHANE	3.16		METHYL METHACRYLATE	0.70	
SEC-BUTYL ALCOHOL	2.81		QUINOLINE	0.70	
BERYLLIUM	2.81		1 2 3-TRICHLOROPROPANE	0.70	
HYDRAZINE	2.46		ACETOPHENONE	0.70	
AMMONIUM SULFATE (SOLUTION)	2.46		P-CRESOL	0.70	
ACETALDEHYDE	2.11		P-PHENYLENEDIAMINE	0.70	
FREON 113	2.11		1 2-BUTYLENE OXIDE	0.70	
CHLORINE DIOXIDE	2.11		EPICHLOROHYDRIN	0.70	
ISOPROPYL ALCOHOL	1.75		ALLYL ALCOHOL	0.70	
TRICHLOROFLUOROMETHANE	1.75		PYRIDINE	0.70	
1 2-DICHLOROPROPANE	1.75		HEXACHLOROBENZENE	0.70	
BUTYL BENZYL PHTHALATE	1.75		HYDROQUINONE	0.70	
PHOSPHORUS (YELLOW OR WHITE)	1.75		BUTYRALDEHYDE	0.70	
ANILINE	1.40		DIMETHYLAMINE	0.70	
CHLOROFORM	1.40		1 2-DICHLOROETHYLENE	0.70	
ETHYLENE OXIDE	1.40		AMMONIUM NITRATE (SOLUTION)	0.70	
CHLOROBENZENE	1.40		THALLIUM	0.70	
1 2 4-TRICHLOROBENZENE	1.40		VINYL CHLORIDE	0.35	
ALUMINUM (FUME OR DUST)	1.40		BROMOFORM	0.35	
SILVER	1.40		ETHYLIDENE DICHLORIDE	0.35	
NITRIC ACID	1.40		BROMOTRIFLUOROMETHANE	0.35	
N N-DIMETHYLFORMAMIDE	1.05		DIMETHYL SULFATE	0.35	
CHLOROMETHANE	1.05		2 3-DICHLOROPROPENE	0.35	
ACETONITRILE	1.05		PERACETIC ACID	0.35	
PROPYLENE OXIDE	1.05		1 1 2 2-TETRACHLOROETHANE	0.35	
1 1 2-TRICHLOROETHANE	1.05		4 4'-ISOPROPYLIDENEDIPHENOL	0.35	
O-CRESOL	1.05		CUMENE HYDROPEROXIDE	0.35	
ACROLEIN	1.05		PHTHALIC ANHYDRIDE	0.35	
VINYL ACETATE	1.05		2 6-XYLIDINE	0.35	
M-CRESOL	1.05		TOLUENE-2 6-DIISOCYANATE	0.35	
DIBENZOFURAN	1.05		BENZIDINE	0.35	
1 3-DICHLOROPROPYLENE	1.05		1 2-DICHLOROBENZENE	0.35	
FLUORINE	1.05		2 4 5-TRICHLOROPHENOL	0.35	
FORMIC ACID	0.70		NITROBENZENE	0.35	
BROMOMETHANE	0.70		METHYLENEBIS(PHENYLISOCYANATE)	0.35	
CHLOROTRIFLUOROMETHANE	0.70		ACRYLONITRILE	0.35	

CHEMICAL	%Prob	Ranking	CHEMICAL	%Prob	Ranking
MALEIC ANHYDRIDE	0.35		SODIUM DIMETHYLDITHIOCARBAMATE	0.35	
1 3-PHENYLENEDIAMINE	0.35		NITRILOTRIACETIC ACID	0.35	
DI(2-ETHYLHEXYL) PHTHALATE	0.35		BROMOCHLORODIFLUOROMETHANE	0.35	
2 4-DICHLOROPHENOL	0.35		TOLUENE-2 4-DIISOCYANATE	0.35	
PROPIONALDEHYDE	0.35		TITANIUM TETRACHLORIDE	0.35	
1 4-DIOXANE	0.35		BROMINE	0.35	

Table 24. Correctly identified chemicals for both Ergon and Clorox.

Facility	CAS	Chemical Name	Inventory	
			Max daily lb	Ave daily lb
Ergon	67-64-1	Acetone	100-999	100-999
	1310-73-2	sodium hydroxide	10,000-99,999	10,000-99,999
	7647-01-0	hydrogen chloride	10,000-99,999	10,000-99,999
	108-88-3	Toluene	1,000-9,999	100-999
Clorox	7664-93-9	sulfuric acid	1,000-9,999	1,000-9,999
	67-63-0	propan-2-ol	100,000-999,999	100,000-999,999
	7439-92-1	Lead	10,000-99,999	10,000-99,999
	1310-73-2	sodium hydroxide	10,000-99,999	10,000-99,999

Table 25. TIC chemicals in the Tier II report for Ergon not Identified by TIC-Master.

CAS	Chemical Name	NFAH/F/I			ITF40 (HSC/HPC/RR)		
8052-42-4	Asphalt, [at or above its flash point]		1	0	Negligible	Likely	Low
8002-26-4	Tall oil	1			Negligible	Occasional	Low
7783-06-4	hydrogen sulfide	4	4	0	Catastrophic	Occasional	High
8008-20-6	Kerosene	2	2	0	Marginal	Likely	Moderate
7727-37-9	Nitrogen, compressed				Indeterminate	Frequent	Not Ranked
7664-38-2	phosphoric acid	2			Marginal	Likely	Moderate
16721-80-5	sodium hydrogen sulfide				Indeterminate	Occasional	Not Ranked

For the sanitation product industry, 113 chemicals were estimated (Table 26). The Tier II report for Clorox Products Manufacturing Co. estimated 23 chemicals. Of these, nine appear in the TIC database, and four of these chemicals were actually identified. Tables 24 and 27 show the TICs identified and not identified, respectively.

Table 26. Full listing of chemicals predicted for sanitation product manufacturing, with the first 30 ranked.

CHEMICAL NAME	%Prob	Rank	CHEMICAL NAME	%Prob	Rank
PHOSPHORIC ACID	41.56	1	FORMIC ACID	1.11	
HYDROCHLORIC ACID	24.67	2	N-BUTYL ALCOHOL	1.11	
SODIUM HYDROXIDE (SOLUTION)	22.89	3	METHYL ISOBUTYL KETONE	1.11	
1 1 1-TRICHLOROETHANE	21.33	4	PHENOL	1.11	
SULFURIC ACID	18.89	5	1 2-DICHLOROBENZENE	0.89	
METHANOL	17.33	6	CUMENE	0.89	
DICHLOROMETHANE	15.56	7	BUTYL ACRYLATE	0.89	
AMMONIA	13.33	8	CRESOL (MIXED ISOMERS)	0.89	
ETHYLENE GLYCOL	10.67	9	CHROMIUM	0.89	
NITRIC ACID	10.22	10	SEC-BUTYL ALCOHOL	0.67	
CHLORINE	10.00	11	PERACETIC ACID	0.67	
DIETHANOLAMINE	9.11	12	METHYL METHACRYLATE	0.67	
XYLENE (MIXED ISOMERS)	8.44	13	PHTHALIC ANHYDRIDE	0.67	
TOLUENE	6.89	14	BUTYL BENZYL PHTHALATE	0.67	
TETRACHLOROETHYLENE	6.89	15	2-ETHOXYETHANOL	0.67	
ISOPROPYL ALCOHOL	5.33	16	ETHYL ACRYLATE	0.67	
SODIUM NITRITE	5.33	17	MERCURY	0.67	
METHYL ETHYL KETONE	5.11	18	NICKEL	0.67	
1 2 4-TRIMETHYLBENZENE	5.11	19	THIOUREA	0.44	
ACETONE	4.67	20	BENZENE	0.44	
DIBUTYL PHTHALATE	4.67	21	TRICHLOROFLUOROMETHANE	0.44	
N-METHYL-2-PYRROLIDONE	4.67	22	DICHLORODIFLUOROMETHANE	0.44	
HYDROGEN FLUORIDE	4.67	23	ACRYLIC ACID	0.44	
TRICHLOROETHYLENE	4.44	24	CHLOROACETIC ACID	0.44	
FORMALDEHYDE	4.00	25	DIETHYL PHTHALATE	0.44	
FREON 113	3.78	26	M-XYLENE	0.44	
1 4-DICHLOROBENZENE	3.56	27	PROPYLENE	0.44	
NAPHTHALENE	2.67	28	HYDROQUINONE	0.44	
ALUMINUM OXIDE (FIBROUS FORMS)	2.67	29	1 4-DIOXANE	0.44	
2-PHENYLPHENOL	2.22	30	ALUMINUM (FUME OR DUST)	0.44	
N-HEXANE	2.22	31	MANGANESE	0.44	
1 1-DICHLORO-1-FLUOROETHANE	2.22		DICHLOROBENZENE (MIXED ISOMERS)	0.44	
STYRENE	1.78		PIPERONYL BUTOXIDE	0.22	
ETHYLBENZENE	1.56		URETHANE	0.22	
NITRILOTRIACETIC ACID	1.33		CARBON TETRACHLORIDE	0.22	

CHEMICAL NAME	%Prob	Rank	CHEMICAL NAME	%Prob	Rank
BIS(TRIBUTYLTIN) OXIDE	0.22		DIPHENYLAMINE	0.22	
CHLOROMETHANE	0.22		POTASSIUM DIMETHYLDITHIOCARBAMATE	0.22	
CARBON DISULFIDE	0.22		SODIUM DIMETHYLDITHIOCARBAMATE	0.22	
ETHYLENE OXIDE	0.22		CAPTAN	0.22	
DICYCLOPENTADIENE	0.22		2-MERCAPTOBENZOTHAZOLE	0.22	
CUMENE HYDROPEROXIDE	0.22		1 2-DICHLOROETHYLENE	0.22	
O-TOLUIDINE	0.22		C.I. BASIC GREEN 4	0.22	
ACETOPHENONE	0.22		TOLUENE-2 4-DIISOCYANATE	0.22	
BENZYL CHLORIDE	0.22		POLYCHLORINATED BIPHENYLS	0.22	
METHYLENEBIS(PHENYLISOCYANATE)	0.22		1-(3-CHLOROALLYL)-3 5 7-TRIAZA-1-AZONIAADAMANTANE CHLORIDE	0.22	
4 4'-METHYLENEDIANILINE	0.22		LEAD	0.22	
P-XYLENE	0.22		ANTIMONY	0.22	
ACRYLONITRILE	0.22		CADMIUM	0.22	
VINYL ACETATE	0.22		COPPER	0.22	
CYCLOHEXANOL	0.22		TETRAMETHRIN	0.22	
2-METHYLPYRIDINE	0.22		PHOSPHORUS (YELLOW OR WHITE)	0.22	
2-METHOXYETHANOL	0.22		TOLUENE DIISOCYANATE (MIXED ISOMERS)	0.22	
CYCLOHEXANE	0.22		PERMETHRIN	0.22	
PROPOXUR	0.22		TRICLOPYR TRIETHYLAMMONIUM SALT	0.22	
DI(2-ETHYLHEXYL) PHTHALATE	0.22		FENARIMOL	0.22	
N-DIOCTYL PHTHALATE	0.22		HYDRAMETHYLNON	0.22	
TRIETHYLAMINE	0.22				

Table 27. TIC chemicals in the Tier II report for Clorox not identified by TIC-Master.

CAS	Chemical Name	NFPA H/F/I			ITF40 (HSC/HPC/RR)		
68608-26-4	Sulfonic acids, petroleum, sodium salts				Indeterminate	Indeterminate	Not Ranked
8052-41-3	Stoddard solvent	2	2	0	Marginal	Indeterminate	Not Ranked
57-55-6	propane-1,2-diol	1	1	0	Negligible	Occasional	Low
151-21-3	sodium dodecyl sulphate	2			Marginal	Occasional	Moderate
61790-12-3	Fatty acids, tall-oil				Indeterminate	Indeterminate	Not Ranked
65-85-0	Benzoic acid	4	1	0	Catastrophic	Occasional	High

### *Most likely chemicals*

Tables 23 and 26 list the first 30 chemicals from each facility ranked in order of most likely. Reviewing the ranking of the first 30 chemicals for the chemical refining industry, it is interesting that even the 30th chemical has a percent probability over 25%. This seems to indicate that this type of facility would likely have a relatively high number of chemicals. In fact, the facility had 44 chemicals reported in its Tier II report.

### *Most significant chemical hazards*

Tables 28 and 29 show refinery chemicals with flammability and health indices of 4.

Tables 30 and 31 show that chemicals from cleaning products manufacturing had flammability (NFPA-F) and health (NFPA-H) indices of 4.

**Table 28. Refinery chemicals with flammability indices of 4.**

NFPA_F	CAS	Chemical Name	%Prob	NFPA_H
4	74908	HYDROGEN CYANIDE	13.33	4
4	302012	HYDRAZINE	2.46	4
4	75150	CARBON DISULFIDE	26.67	3
4	75218	ETHYLENE OXIDE	1.40	3
4	75569	PROPYLENE OXIDE	1.05	3
4	124403	DIMETHYLAMINE	0.70	3

**Table 29. Refinery chemicals with health indices of 4.**

NFPA_H	CAS	Name	%prob	NFPA_F
4	71432	BENZENE	78.95	3
4	74908	HYDROGEN CYANIDE	13.33	4
4	79107	ACRYLIC ACID	0.70	2
4	79210	PERACETIC ACID	0.35	2
4	91225	QUINOLINE	0.70	2
4	107028	ACROLEIN	1.05	3
4	107131	ACRYLONITRILE	0.35	3
4	107186	ALLYL ALCOHOL	0.70	3
4	108952	PHENOL	39.30	2
4	120832	2 4-DICHLOROPHENOL	0.35	1
4	123319	HYDROQUINONE	0.70	1
4	302012	HYDRAZINE	2.46	4

NFPA_H	CAS	Name	%prob	NFPA_F
4	7439976	MERCURY	20.70	i
4	7440439	CADMIUM	3.86	i
4	7632000	SODIUM NITRITE	3.51	i
4	7647010	HYDROCHLORIC ACID	34.04	0
4	7664393	HYDROGEN FLUORIDE	22.81	0
4	7664939	SULFURIC ACID	61.75	0
4	7697372	NITRIC ACID	1.40	0
4	7723140	PHOSPHORUS (YELLOW OR WHITE)	1.75	i
4	7726956	BROMINE	0.35	i
4	7782414	FLUORINE	1.05	0
4	7782505	CHLORINE	53.68	0

Table 30. Cleaning product chemicals with flammability indices of 4 or higher.

NFPA F	CAS	Name of Chemical	% Prob	NFPA H
4	75150	CARBON DISULFIDE	0.22	3
4	75218	ETHYLENE OXIDE	0.22	3
4	74873	CHLOROMETHANE	0.22	2
4	115071	PROPYLENE	0.44	1

Table 31. Cleaning product chemicals with health indices of 4 or higher (i indicates insufficient data).

NFPA H	CAS	Name	% Prob	NFPA F
4	56359	BIS(TRIBUTYLTIN) OXIDE	0.22	i
4	71432	BENZENE	0.44	3
4	79107	ACRYLIC ACID	0.44	2
4	79210	PERACETIC ACID	0.67	2
4	107131	ACRYLONITRILE	0.22	3
4	108952	PHENOL	1.11	2
4	123319	HYDROQUINONE	0.44	1
4	7439976	MERCURY	0.67	i
4	7440439	CADMIUM	0.22	i
4	7632000	SODIUM NITRITE	5.33	i
4	7647010	HYDROCHLORIC ACID	24.67	0
4	7664393	HYDROGEN FLUORIDE	4.67	0
4	7664939	SULFURIC ACID	18.89	0
4	7697372	NITRIC ACID	10.22	0
4	7723140	PHOSPHORUS (YELLOW OR WHITE)	0.22	i
4	7782505	CHLORINE	10.00	0

Finally, the following chemicals were given as part of the exercise:

- 8002-26-4 tall oil: F1, H1
- 7783-06-4 hydrogen sulfide: F4, H4
- 64741-44-2 diesel: not in database (F2, H0)
- 8008-20-6 kerosene: F2, H2
- 8006-61-9 gasoline: F3, H1
- 68608-26-4 hostapur: F1, H2
- 57-55-6 propylene glycol: F1, H1
- 65-85-0 benzoic acid: F1, H4
- 8002-09-3 terpene alcohol: not in database (F1, H1)

The three greatest hazards were determined:

1. Hydrogen cyanide, which had NFPA-H = 4 and NFPA-F = 4. Associated with refining.
2. Hydrogen sulfide, which also had an NFPA-H = 4 and an NFPA-F = 4.
3. Benzene: NFPA-H = 4 and NFPA-F = 3. Associated with both industries.

Once again, strong cases could be made for other chemicals.

Recommendations were made using the information in the database, the SOPs, and professional judgment. Fire protection equipment and breathing protection are recommended, particularly self-contained breathing apparatus for those who may have to enter the area while fires from the suicide bomber explosions are burning. For HCN releases, level A protection is strongly recommended, including use of SCBA, at least until testing verifies safe levels. For organic vapors (benzene), SCBA or organic vapor cartridges in air purifying respirators are recommended. Level B protection (Tyvek) should be sufficient, but Level A could be warranted for highly contaminated areas. Of course, this protection is provided on the order of SCBA and the need for test results might not be practical for a military operation. This would make other courses of action more attractive.

#### *Release modeling*

A model run was conducted evaluating the release of 200 gal of hydrogen cyanide (HCN) from the refinery. Figure 21 is an aerial photo showing the release.



Figure 21. Modeled release of HCN.

#### *Response to environmental question*

The test team provided a response to the environmental question, “Can you address community concerns about 7783-06-4 Hydrogen sulfide concentrations in a nearby waterway?” The response reads,

Hydrogen sulfide ( $\text{H}_2\text{S}$ ) generally exists as a gas. However, it can dissolve in water. It rarely will cause a health concern in water, but can make it unpalatable as it can impart a foul taste (rotten egg) in the water. Obviously, a terrorist attack would be an extraordinary circumstance that could affect normal access to utilities, such as drinking water. Because of the issues with  $\text{H}_2\text{S}$  and with other, even more toxic chemicals, we would recommend that those residents who are not evacuated be put on bottled potable water. If the waterway is contaminated, a cleanup would be instituted. For  $\text{H}_2\text{S}$ , which is a highly reduced chemical, the easiest removal mechanism would be aeration, which would drive the gas out of solution or oxidize the sulfide into sulfate forms.

## **Test 5 summary**

### **Scenario**

The Test 5 scenario assumes that U.S. Armed Forces want to select a location for an operation. Two sites are under consideration. Intelligence has confirmed the presence and volumes of the chemicals listed. Armed resistance is expected if either site is acquired or maintained. The test

team was asked to identify the preferred location and state the basis for the recommendation. The evaluation should include the potential impact to municipal facilities like parks, schools, fire stations, hospitals, etc. (Figure 22).

Test <u>5</u> Scenario		
Date: <u>12/03/09</u> Time Submitted: <u>0915 CST</u>		
<u>Industrial Processes or Process Name:</u>	<u>TIC Release Mechanism:</u> US Armed Forces want to select a location for an operation. Two sites are under consideration. Intelligence has confirmed the presence of the chemicals listed below. Armed resistance is expected if either site is acquired or maintained.	<u>Atmospheric conditions:</u> Temperature: 39 deg F Wind Speed: 15 mph Wind Direction: West Precipitation: None
<u>TICs:</u> Site A: Chemicals include 143-33-9 Sodium Cyanide (100-999 lbs), 151-50-8 Potassium Cyanide (1,000-9,999 lbs), and 7664-93-9 Sulfuric Acid (1,000-9,999 lbs). These volumes are the average daily amounts.  <b>Note: This is a complete chemical list.</b>  Site B: Chemicals include 7647-01-1 Muriatic Acid (1,000-9,999 lbs); 7664-93-9 Sulfuric Acid (1,000-9,999 lbs); 8006-64-2 steam distilled turpentine (1,000-9,999 lbs); 1310-73-2 sodium hydroxide (10,000-99,999 lbs); 25155-23-1 trixylenylphosphate (1,000-9,999 lbs); 8002-26-4 Tall oil (1,000-9,999 lbs); 7704-34-9 sulfur (1,000-9,999 lbs); and 7681-52-9 sodium hypochlorite (1,000-9,999 lbs). These volumes are the average daily amounts.  <b>Note: This is only a partial chemical list.</b>		<u>Other Potential Hazards:</u>   Site B: Other potential hazards include 64742-96-7 Diesel Fuel (1,000-9,999 lbs); and 8006-61-9 Gasoline (1,000-9,999 lbs).
<u>Site Designation:</u> Single Site: ____ Multiple Sites: <u>X</u> General Location*: ____	Site A*: Longitude: 90° 48' 32.31" W Latitude: 32° 22' 37.36" N Elevation:	Site B: Longitude: 90° 50' 13.86" W Latitude: 32° 19' 54.29" N Elevation:
<u>Other Characteristics:</u> Identify the preferred location and state the basis for the recommendation. Include in your evaluation the potential impact to municipal facilities like parks, schools, fire stations, hospitals, etc.		
* See the <u>Other Characteristics</u> section. For example, the U.S. Army may wish to occupy a 2-square-mile area centered at the location provided by Site A*.		

Figure 22. Scenario description for Test 5.

## Results

### *Response time*

The estimated working time was just over 5 hr.

### *Facility information*

This test was significantly different than the previous tests in two ways. First, the given locations did not correspond to the real facility. The Site A coordinates refer to the Pecan Ridge Apartment Complex at 2501 Culkin Road, Vicksburg, MS (Latitude (Lat) 32°22'37.36"N; Longitude (Long) 90°48'32.31"W). This complex is just east of River Region Hospital and Sherman Avenue Elementary School. The Site B coordinates refer to a steel facility at 1250 Highway 27, Vicksburg, MS (Lat 32°19'54.29"N; Long 90°50'13.86"W). A second departure was that an NAICS code was not provided. Rather, a listing of chemicals was provided for the two facilities. The team needed to determine the industry from the chemicals provided. The test team estimated that Site A was a metal working facility with an NAICS code ranging from 331000 to 333000. In particular, the team guessed that the facility was a rolled steel shape manufacturer (3331221). This response was accurate. The test team's response proposed that Site B was a paper processing facility: a paper mill (322121) or a paperboard mill (322130), based on the presence of Kemira ama-1750, which is a biocide specific to paper products and the combination of flotation agents, bleach, fuel, and acids/bases. This was also correct. Figure 23 is a location map.



Figure 23. Location map for Test 5.

### Chemical identification

Chemicals for this exercise were provided (see Figure 22). Table 32 summarizes the chemicals and their risks.

**Table 32. Chemicals from Sites A and B for Test 5.**

Site A		NFPA_H	NFPA_F	NFPA_I	Health Hazard	Flammability Hazard	Comments
24613-89-6	chromium 111						
64742-52-5	slushing oil						
1310-73-2	sodium hydroxide	3	0	1	X		
7632-00-0	sodium nitrate	4			X		this CAS is for sodium nitrite, nitrate is 7631-99-4, both ha
7664-93-9	sulfuric acid	4	0		X		
64742-48-9	vanishing oil	1	2	0			
Site B		NFPA_H	NFPA_F	NFPA_I	Health Hazard	Flammability Hazard	Comments
1310-73-2	sodium hydroxide	3	0	1	X		
7647-01-0	hydrogen chloride	4	0	1	X		
7664-93-9	sulphuric acid	4	0		X		
7681-52-9	sodium hypochlorite	1	0				
7704-34-9	sulphur, precipitated, subli	3	1	0	X		
8002-26-4	Tall oil	1	2	0			flotation agent?
8006-61-9	Gasoline	1	3	0		X	
8006-64-2	Turpentine, oil	1	3	0		X	
25155-23-1	trixylenylphosphate	1					fire resistant hydraulic fluid used in electrical devices
64742-96-7	diesel fuel	0	2	0			
111-30-8	Kemira ama 1750						biocide
1305-78-8	calcium oxide (lime)	1	0	1			
497-19-8	sodium carbonate	2					
16721-80-5	sodium hydrosulfide	3	2	0	X		

blue indicates changes or new data to add to database

### Chemical hazards

Site A has two chemicals with a health risk of 4 and one with a health risk of 3:

- Sodium nitrate
- Sulfuric acid

Site A does not have any chemicals with flammability of 3 or higher.

Site B also has two chemicals with a health risk of 4 and three chemicals with a health risk of 3:

- Hydrogen chloride (4)
- Sulfuric acid (4)
- Sodium hydroxide (3)
- Sulfur, precipitated (3)
- Sodium hydrosulfide (3)

Site B had no chemicals with a flammability hazard of 4, but had two chemicals with flammability hazards of 3:

- Gasoline
- Turpentine oil

#### *Release modeling*

Figure 24 shows the results of a simulation of sulfuric acid releases from both facilities. The volume of release from both sites was 99,999 lb. Wind velocity and direction were given in the exercise.

Observations from the GIS data:

#### Site A:

1. The wind is blowing from the west; therefore, the resulting plume will most likely drift eastward, and continue to follow the dominant wind direction.
2. From the predicted plume resulting at Site A, the population potentially to be affected is approximately 1,638 people (using the census feature provided by the GIS system).
3. One school is present within the predicted plume (see green marker). No other facilities were identified.

#### Site B:

1. The wind is blowing from the west; therefore, the resulting plume will most likely drift eastward, and continue to follow the dominant wind direction.
2. From the predicted plume resulting at Site A, the population potentially to be affected is approximately 888 people.
3. None of the critical infrastructure (school, hospital, fire station, or police) is found in the predicted plume for site B (Figure 24a). However, based on the map, one can state that if the wind direction changes from "from west" to "from south," then three schools just northeast of Site B are affected, and these are quite close, closer in fact than the school affected in Site A.

#### *Best site to perform operation*

***Analysis of chemical inventory.*** Based on the chemical inventory provided, Site B appears to have the most hazards. It has more chemicals, a higher total quantity of chemicals, more chemicals with health hazards of 3 or 4, and more chemicals with flammability hazards. This would make Site A a better choice to conduct the operation.

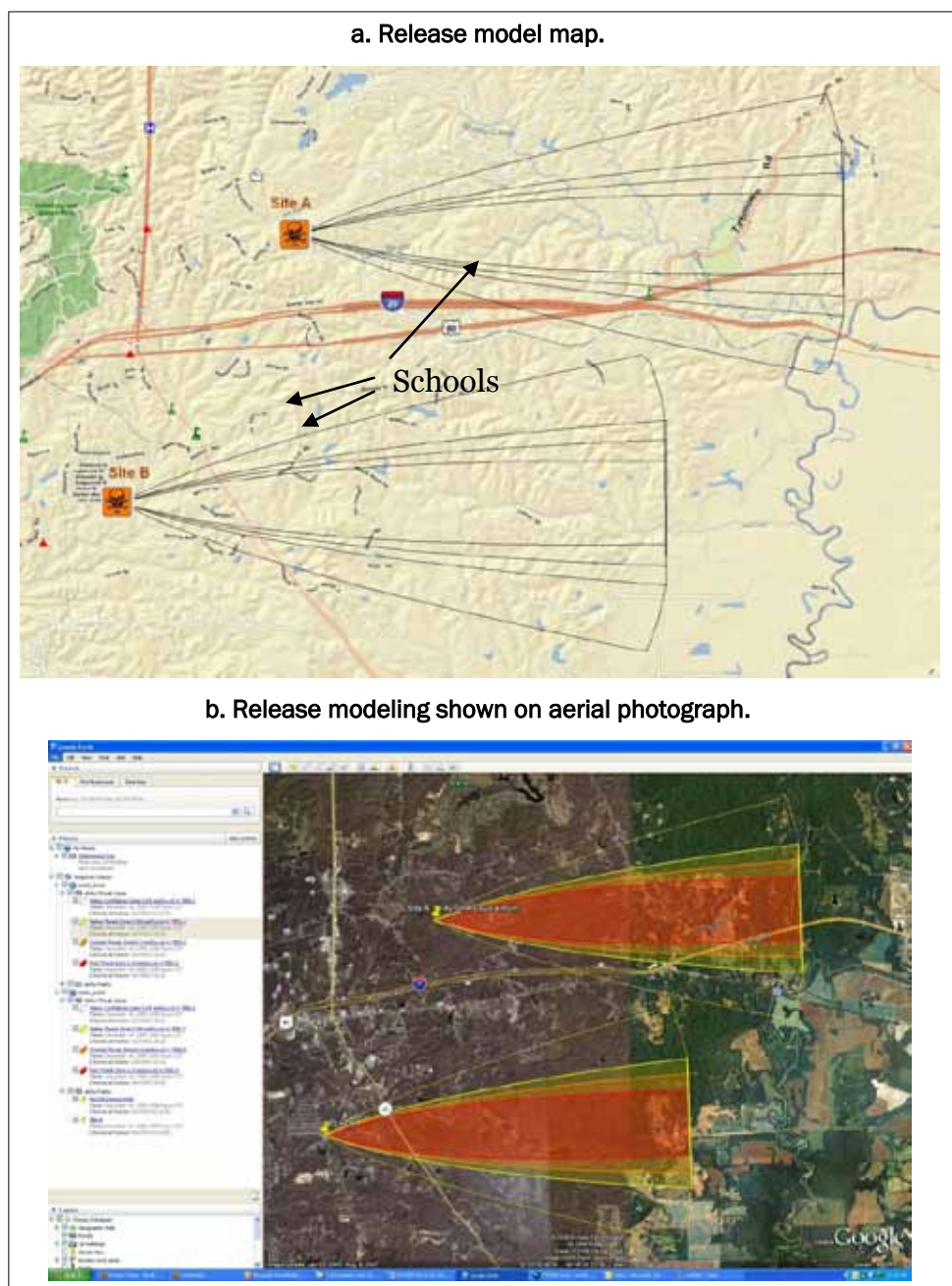


Figure 24. Maps of releases from two facilities from Test 5 scenario.

**Assessment of release scenario.** Site B would affect fewer people than Site A and would not affect any schools, compared to one school affected by Site A. However, this can change if the wind direction changes. The plume from Site A would cross Interstate 20, a major transportation highway. The plume from Site B would cross highway 27, which is not nearly as busy.

**Final assessment.** Overall, Site B is recommended as the best choice to conduct the operation. If a release should occur, Site B should affect a

smaller population and should not impact schools. There are some caveats; there are two schools close to site B that could be affected if the wind should change direction and Site B involves more chemicals that could be problematic. Chemical information and modeling can greatly aid in choosing the site.

## 5 Discussion

### Response time

The response time of 2 hr was only met for one of the tests: Test 2. Response times ranged from 2 to more than 5 hr. The overall average response time was close to 3 hr. Response times were affected by the fact that the test team included participants from three different locations and that, due to conflicting schedules, they were not always able to work together at the same time. Of course, similar issues would likely occur in using the TIC-Master program during an actual scenario.

### Facility identification

For Tests 1-4, the NAICS code was provided. In each case, the proper facility type was derived from the code, which is entirely expected. For Test 5, complete chemical lists were provided for two facilities, but the type of facility was not. The team was able to determine the facility. However, for both facilities, there were distinctive chemicals that greatly aided the team in identification. It is clear that determining chemical facilities from a partial list is challenging due to the fact that most chemicals can be used for a wide range of industrial processes.

### Chemical prediction

#### Comparison of total number of chemicals

Table 33 summarizes the chemicals predicted for Tests 1-4 (in Test 5, the chemical list was provided). First, it is clear that the database provided more chemicals than those actually found at the site, ranging from a factor of slightly less than 2 (for Test 3) to more than 5 (Test 2). This reflects variability of chemicals used by similar industrial processes by different facilities. Analyzing possible chemicals from a particular industrial facility based on chemicals used by similar industries will likely lead to this overestimation. The challenging aspect of this technique is that the decision maker may have to deal with extraneous information that makes hazard estimation more cumbersome.

Table 33. Scenario summary.

Scenario	# TICs Predicted	Tier II Report	Tier II TICs (In TMaster DB)	Predicted TICs Using TMaster DB
1	156	58	35	15/35
2	26	5	4	2/4
3	91	53	30	15/30
4 Site A	158	44	11	4/11
4 Site B	113	23	9	4/9

### TIC chemical predictions versus chemicals actually reported

Table 33 indicates that the total chemicals reported in the Tier II reports of the various industries were more than those commonly found in the database. As discussed in Chapter 2, the database developed for this program was inclusive of TICS defined by the combined lists of the ITF-40 (which included chemicals listed in ITF-25), and EPA TRI. Chemicals in common with the TIC-Master database can be definitively called TICs. However, TICs do not encompass all hazardous chemicals. The Tier II report had some hazardous chemicals not defined as TICs, as was expected. Expanding the list to incorporate other chemicals would probably not be too difficult. Additional chemical lists are readily available. The VFD indicated, for example, that programs like the Computer Aided Management of Emergency Operations (CAMEO - which focuses on releases, but does not have TIC-Master's predictive effort) lists 5000 to 7000 chemicals (National Oceanic and Atmospheric Administration (NOAA), undated). However, expanding the chemical list could also result in a similar increase of total chemicals reported, which was up to a factor of 5 (see Chapter 5). It is not clear if expanding the chemical would actually improve the ability to assess the risk at a given site, at the expense of information overload. A data gap analysis could be useful to determine if the addition of more chemicals would improve the system performance (Dortch et al. 2005).

Links could easily be developed to existing databases that would allow for a wider range of chemicals as well as allow users to check chemical data and recommendations from various sources. Some useful databases include:

- Wireless Information System for Emergency Responders (WISER) (<http://wiser.nlm.nih.gov/>)
- CAMEO (<http://cameochemicals.noaa.gov/>)

- Chemical Biological Response Aid (CoBRA)  
(<http://www.defensegroupinc.com/cobra/>)
- ALOHA (<http://www.chemicalspill.org/OffSite/aloha.html>).

Of these, ALOHA was the only database linked to TIC-Master, but the other databases could also be linked.

Table 33 also compares the defined TICs found at the site with those predicted by the TIC-Master program. The program identified about 50% of the correct TICs. This must result from deficiencies in the industry-chemical crosswalk. The relationships used to develop the crosswalk were found using the EPA TRI database. As mentioned in Chapter 2, this database uses a more limited list (less than 600 chemicals) compared to the combined list used in the TIC-Master database (more than 2000). Furthermore, the TRI focuses on released chemicals. It is possible that industries may have chemicals that were not released, so they are not reported on the TRI searches.

The TRI database was the best tool found to develop industry-chemical relationships. However, a better tool could be available within the next 2 to 3 years. As mentioned in Chapter 1, Tier II reporting has been primarily a paper process. However, the EPA is now strongly encouraging electronic submission (USEPA 2010). This will likely lead to digitized records of the Tier II reports that can be searched and sorted. Such a database would likely improve an ICCW. However, it would also likely increase the total number of chemicals estimated, once again increasing the potential for information overload. One advantage of the existing TRI database is that it focuses on released chemicals, which are probably of most importance.

### **Top 10 chemical comparison**

For scenarios 1-4, the test team predicted the top 10 chemicals on site. In Scenario 2, two of the top 10 most probable chemicals were included in the Tier II Report. However, in this test, there were only four total chemicals, and a third top 10 chemical (i.e., ammonia) was very similar to another Tier II chemical (i.e., aqueous ammonia). For the other tests, the top 10 chemicals were represented 50% to 80% in the Tier II Reports. Using the ranking feature to identify most likely chemicals appears to be a very useful feature for focusing risk assessment.

## Assessment of hazards/risks

Testing lacked a systematic means of comparing the risks from the predicted list to the actual. However, by looking over the chemicals, reviewing the hazard information, and applying professional judgment, it was determined that the risks derived from the chemicals found in the predicted chemicals completely encompassed the risks from the actual chemicals in all cases. A systematic screening tool, possibly based on the principles of decision science, could be useful to allow a first screening of the chemicals identified, which can be numerous for certain industrial processes.

As noted in Table 1 (Chapter 2), the health hazard data used during the tests was derived initially from the ITF-40 data. However, some of these data were replaced by NFPA health rankings, especially in the cases where a refereed NFPA ranking was available and it had a value greater than ITF-40. The health hazard listings for the tests were categorically labeled as NFPA, which is misleading. Although these values are frequently the same, they are not derived with the same criteria, so that the test results may involve mixed health rankings that were used to assess chemical risks. It is believed that this did not affect the performance of the TIC-Master to assess the hazards/risks because the test team used the health rankings in the TIC-Master database to assign health rankings to actual chemicals for comparison. However, the next version of the database should correct this issue by being consistent with health hazard definition and labeling. Table 12 (Chapter 4) provides examples of the inconsistency: the health ranking for ammonia is “3,” which is an example of a higher refereed NFPA code replacing the ITF-40 value (1); methanol also is ranked “3,” which is an example of an ITF-40 value that is greater than a refereed NFPA value (1) and so remained in the database; aqueous ammonia is represented by “4,” which is an example of a chemical that does not have a refereed NFPA ranking so the ITF-40 value remained in the database; and the health ranking for ethanol is “2,” which is an example of a chemical with a refereed NFPA ranking matching the ITF-40 ranking.

## Release modeling

ALOHA was chosen as the modeling program after discussions with the Vicksburg Fire Department. The program is freely available and is designed to be user-friendly. The program was found to work well for most simple explosive and atmospheric releases.

However, ALOHA had some limitations. It is simple to use because it has preset release scenarios, such as tank sizes and configurations. It is not easy to model situations beyond these scenarios, nor is it easy to model multiple-component releases.

Perhaps the greatest limitation of ALOHA is that it focuses solely on atmospheric releases. Several of the scenarios involved facilities near streams and rivers. It would have been helpful to determine whether contaminant release would impact these water bodies and, if so, the effect.

The good news is that it is entirely possible to add release modeling programs to the GIS system. In fact, in many cases, these various programs can be used together as a basis of comparison. CAMEO is a program developed by the EPA that is capable of airborne release modeling. According to the VFD, they typically attempt to model a release using ALOHA first. If this proves insufficient, they will then use CAMEO. CAMEO also has capabilities to assist industries in tracking chemicals and using this information to prepare Tier II reports. However, CAMEO does not have surface release capabilities.

The Hazard Prediction and Assessment Capability model (HPAC) was developed by the Oak Ridge National Laboratory, and it couples atmospheric and surface depositional modeling capability, which would make it a good fit for this application. GeoRAMS, a program developed to estimate the effect of hazardous chemical spills, has four models integrated for air, soil, surface water, and pipe rupture (ERDC-EL 2005).

The Joint Effects Model (JEM, <http://www.jpeocbd.osd.mil/packs/Default.aspx?pg=1701>) is a suite of models being developed for the DoD that will combine the best models for air, explosive, and surface water release of chemical weapons, nuclear and radiological agents, and TICs. This model will be an obvious choice to integrate with systems like the TIC-Master.

Surface water modeling programs commonly used by the Corps of Engineers include CASCade 2 Dimensional SEDiment (CASC2D-SED) (Johnson et al. 2008) and Gridded Surface/Subsurface Hydrologic Analysis (GSSHA). Both of these are distributed sources models and are capable of achieving excellent results for modeling contaminant movement in the environment. However, these models require a higher level of input data

and require longer run times than HPAC, which is specifically designed for emergency situations.

In terms of the data, a drawback is the lack of quantity estimation capability. Therefore, the quantity for each release had to be estimated, unless it was provided by the test condition (equivalent to the intelligence available). One guideline is that a given chemical could be assumed to be available in quantities greater than the reporting threshold (generally 1000 kg, although this can vary). Quantity estimation is discussed in the section that follows.

## **Quantity estimation**

TIC-Master has the ability to estimate chemicals related to industrial processes. However, the TIC-Master system does not currently have the capability to estimate chemical quantities. One assumption could be made if a chemical is identified - it is likely at a quantity greater than its reporting limit. Other than that, no quantity information is provided.

Simply knowing the type of industry is not enough information to estimate quantity. However, first order quantification may be possible if some details of the facility are known. Information such as the size of the facility footprint, information on shipping capacity of the facility, or the number and size of certain critical chemical storage tanks could allow for order of magnitude estimates of chemical quantities. Developing these relationships requires a concerted effort to work through a range of industries, each of which may have different criteria that could be used to estimate quantity.

## **Identification of industry**

For the first four exercises, either the NAICS code was provided or the industry name was provided, meaning that TIC-Master was used to determine the industry or NAICS code, respectively. Obviously, these are easily accomplished tasks. In test 5, the team was tasked with identifying two industries from a list of chemicals from each facility. The task was accomplished, but it was challenging. Most chemicals were found to be of limited value in identifying an industry, or even in greatly narrowing down the possibilities, because they are so widely used. However, some industries have distinctive chemicals that could be very useful for quickly identifying the industry.

Improved algorithms to identify industries from complete or partial chemical lists would be valuable. Identification of distinctive chemicals could be a key piece. However, evaluation of multiple chemicals could be valuable to narrow down possible industries.

The GIS system could also be a useful tool for industry identification. Tools could be developed to evaluate facility footprints obtained from aerial imagery and compare them to a library of footprints to look for similarities. Such a system would require considerable development to design, construct, and evaluate.

### **Standard operating procedures (SOPs)**

SOPs had a limited role in the testing of the TIC-Master program. The SOPs were used to evaluate three chemicals in Test 1 (page 35), one in Test 2 (page 43), and provided information to address the environmental question in Test 4 (page 63). It is clear that the information provided was quite valuable in assessing risks and promoting mitigation. However, the effort to produce these SOP reports was substantial. Creating these reports for all the chemicals in ITF-40 would be costly. Future efforts might be better spent in other system enhancements.

### **System modification recommendations based on test results**

After evaluating the results of five tests, concepts on improving TIC-Master for further use were developed. Recommendations resulting from the test observations are:

- Develop a chemicals-to-industry crosswalk to improve estimation of industries based on partial chemical lists. This should focus on the use of industry-distinctive chemicals whenever available.
- Make the system easier to use in general.
- Make it easier to add sites and define and redefine industry use.
- Increase the chemical database to contain other hazardous chemicals that are not defined as TICs.
- Develop a means of estimating volume of chemicals. This process might be challenging, but could use relationships based on facility size, number of storage tanks, etc.
- All TICs are a concern. However, some are particularly dangerous. Right now, the list is presented alphabetically. It would be useful to develop an alarm when the most hazardous materials are identified.

- Address the importance of the health hazard rankings by defining and identifying the primary methods for determining the rankings and provide hazard ranking data for more chemicals, in particular for the TRI-based chemicals that comprise the chemical prediction subset.
- ALOHA is adequate for air modeling. Groundwater modeling would require too much data and probably would not be a factor to consider in the IPB process. However, a surface release model would enhance capabilities. A logical step would link the system to the new Army Joint Effects Model (JEM).
- Weave in identification of areas of concern, including topographic factors (such as streams, lakes), transportation factors (i.e., key roads, railroads), dense residential areas, and schools, hospitals, fire departments, etc. Test 5 touched on this by identifying schools and allowing searches for population counts in the plume area. However, it would be useful to automatically provide more details of these features.

### **Feedback from possible user organizations (MANSCEN, the Army Intelligence Center, Army Geospatial Center, Program Manager for Army Geospatial Data Model)**

Various meetings were carried out with the Army Maneuver Support Center of Excellence (MSCoE), the Army Intelligence Center (AIC), the Army Geospatial Center (AGC) and the Program Manager for the Army Geospatial Data Model. MSCoE, AIG, and AGC were considered possible end users. AGC and PMAGDM were possible hosts for the program, or at least elements of the program. A fact sheet has been prepared to provide interested parties with information on the program (Appendix D).

#### **Army Intelligence Center**

The AIC was briefed on 8 April 2009 at Fort Huachuca, AZ. Kevin Enright of the AIC attended the project review on 27 May 2009. AIC identified the program as providing unique and important capabilities that can be used by the intelligence community. The AIC agreed with the assessment that the Global War on Terror, Operation Enduring Freedom, and Operation Iraqi Freedom indicate that operations in urban environments may be unavoidable, which brings into play issues with chemicals from industrial facilities. The AIC developed an analogy where chemical hazards could be analogous to geographical hazards, affecting the mobility of forces where they are encountered. Overall, the AIC was interested in and supported the development of TIC-Master. A support email by Mr. Enright stated "I do

believe that what you are planning will be helpful for maneuver commanders by showing them a new aspect of the battlespace and its effect on friendly and enemy movement" (Appendix E). However, the AIC indicated that although TIC-Master output would be used in the IPB process, it would likely be operated by other units, such as engineers, chemical officers, or geospatial officers.

A key portion of the discussion focused on whether the system would benefit from classified sources. Without getting into specifics, classified information on various industrial facilities throughout the world is available. Tapping into these sources could enhance the value of the system. However, there are severe limitations to using a classified system that would hamper the effectiveness of TIC-Master if these sources were utilized. First, a secured area would have to be set up in the field to facilitate its use. Second, many National Guard and Reserve units do not have SIPRnet (the DoD internet that contains secret and top secret sources) resources to use or train with the system. Third, only the United Kingdom and Australia currently have access to the US SIPRnet system. Many coalitions include countries beyond these. Because of these limitations, it was decided to keep the system unclassified at this time. The AIC also indicated that a system that is unclassified to a point, but can draw upon classified information, is actually quite cumbersome.

As an unclassified system, it would still be important for proper security to be maintained. Fortunately, the Department of Defense and the Army have greatly enhanced their unclassified security capabilities. Use of passwords and controlled access cards (CAC) could provide sufficient security for unclassified use.

The AIC indicated that using ARCGIS as the program platform was a good choice and that the methods used would be easily integrated into Army operations.

### **Feedback from MSCoE**

The MSCoE was determined to be a potential end user of TIC-Master program. Several interactions were conducted with MSCoE. On 27 April 2008, a workshop was conducted at Vicksburg, MS. MSCoE experts participated both in person and via teleconference. On 22 April 2009, a briefing was conducted to MSCoE at Fort Leonard Wood (FLW). On 08 June 2010, a phone meeting was conducted with the Experimentation

Branch of MSCoE. On 15 June 2010, a phone briefing was conducted with various MSCoE units. And on 20-21 July 2010, the TIC-Master team conducted a series of meetings at FLW to demonstrate the TIC-Master program. The following conclusions were obtained from these meetings:

- The TIC-Master system offers unique capabilities in assessing chemicals and risks associated with industrial facilities.
- This capability cross cuts multiple Army capabilities (engineer, chemical, intelligence, geospatial). However, of these organizations, the engineer units would be most likely to use such a program.
- The system has potential value for battle plan preparation as it was originally conceived. However, its greatest value may in assessing base camp locations.

Although the Corps of Engineers would most likely utilize the program for planning, response to releases from an acute perspective is the responsibility of the Chemical Corps (Mauroni 2002). So, these units may also find the program very useful.

The following improvement areas were identified:

- Provide better and more accessible information on environmental fate, management, and response.
- Provide more information on personal protective equipment considerations. Particularly useful for the Army would be an assessment of whether standard issue protective equipment is suitable for protection.
- Provide links to resources for response to release of the contaminants. A particular resource discussed was The Emergency Response Guidebook (U.S. Department of Transportation (USDOT) 2008), which is available in electronic format. It would be even more useful if this resource could be automatically linked to specific chemicals. Links could also be provided to internet sources on the chemicals.
- Provide chemical lists for illegal facilities (i.e., drug manufacturing, bomb making).
- Provide multiple chemical and plume release capabilities.

Interestingly, a briefing to the Department of Public Works (DPW) generated substantial interest in the program. The DPW was interested in the concept of modifying the program to manage chemical risks at the Fort itself. The GIS and release modeling capabilities were determined as useful

tools for both planning and reactive situations. A white paper has been prepared proposing a small program to modify the existing program for this application (Appendix F). Such a project could be a good segue into developing civilian applications (Chapter 5).

Overall, MSCoE determined that a product based on the TIC-Master platform could be very useful to several mission areas. Consequently, they provided a formal endorsement memorandum supporting the project (Appendix G).

### **Interactions with the Army Geospatial Center**

Briefings were carried out at the Army Geospatial Center (AGC) at Fort Belvoir (Alexandria, VA) on 24 February 2009. Of particular interest was a briefing with the Program Manager of the Army Geospatial Data Model (PMAGDM), Kevin Backe. The AGDM is a collection of geospatial data whose maintainance is critical for the Army that is not maintained on other geospatial systems. The PMAGDM was not interested in the TIC-Master program as a whole, but rather focused on determining if the data model for the system was useful for the AGDM and if the data were in a format that could be imported. After briefing the PMAGDM and allowing Mr. Backe to review the data model, it was determined that the information contained in the system, particularly the ICCW, was unique and should be incorporated into the AGDM.

### **Adaptations for civilian use**

As the project progressed, the topic of adaptation for civilian use was frequently discussed. Civilian organizations who contributed feedback included the VFD, the Mississippi Department of Environmental Quality, MEMA, disaster response groups from the Corps of Engineers Vicksburg District, and homeland security researchers at Applied Research Associates, Inc. The obvious difference with civilian applications is that the capability to estimate chemicals would not be required, as chemical information from EPA Tier 2 reports is generally available. However, features such as the ability to identify TIC facilities within an area, identify populations, roads, and various public facilities in the area, and to model releases could be of great value in various civilian applications. On a smaller scale, a system could be useful for a municipality or even a given industrial complex with multiple sites in a given area to track and manage their hazardous material. The release modeling features could be useful for training exercises as well

as in an actual release response incident. TICs and other hazardous materials could be used by terrorists as a means of attack on the United States (Bennett 2003). From a police or homeland security perspective, the system could be useful for training and for scenario exercises to improve preparedness to terrorist or natural disaster incidents.

A key issue is importing chemical data into the system. Currently, EPA Tier II reports are not available electronically. However, as mentioned in Chapter 5, this is expected to change within the next 2 to 3 years. Once these data are available electronically, it should be possible to mine it for precise chemical information at a given facility.

Until that time, a stop gap measure could be the use of the existing TRI database, which, as discussed in Chapter 5, has some limitations, but does focus on chemicals most likely to be released. Another option would be to manually input information from EPA Tier II reports. This is obviously a cumbersome approach, but might be feasible for limited applications.

The ability to use a GIS system to locate and identify industrial facilities could be very beneficial during natural disaster response. Following Hurricane Katrina, many areas had nearly complete loss of signs, including road signs and business identification signs. This, as well as extensive facility damage, made it difficult to identify industrial facilities. The TIC-Master system could be a valuable tool for these applications.

## 6 Conclusions

This work unit produced an integrated GIS/database system called the TIC-Master program. The system was evaluated under five different scenarios. The response times ranged from 2 to 5 hr, averaging about 3 hr. The program was able to correctly identify the industries from the NAICS codes and manufacturing processes provided in the scenarios and was able to successfully predict two industries from the chemical lists provided in Scenario 5. TIC-Master listed 2 to 5 times more TICs than found in the EPA Tier II report for the facilities. In general, TIC-Master precisely identified about 50% of TIC chemicals actually reported. However, the chemical risks appeared to be encompassed by the chemicals that were identified. Narrowing the chemical list down by chemical frequency (in this case, using a 10 most likely list), helped to focus the predicted results. ALOHA was an adequate model for explosions and atmospheric releases; however, other release modeling programs could be added to provide additional capabilities, including surface spills.

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## **Appendix A: Chemical Database for the TIC-Master Program**

To access the chemical database for the TIC-Master Program, contact Dr. Victor Medina at:

[Victor.F.Medina@usace.army.mil](mailto:Victor.F.Medina@usace.army.mil)

## Appendix B: Instructions for Using the EPA TRI Database to Create ICCW

Instructions to fill crosswalk with data for one industrial six-digit code:

- A. Perform custom query of U.S. Environmental Protection Agency (USEPA) Toxic Release Inventory (TRI) database
  1. Go to customized TRI query at <http://www.epa.gov/enviro/html/tris/adhoc.html>
  2. Select Facility information
  3. Select Form R Reporting information
  4. Select Submission NAICS Codes
  5. Click the Step 2 button
  6. Select tri\_facility
  7. Select v\_tri\_reporting\_form
  8. Select v\_tri\_submission\_naics
  9. Click the Step 3 button
  10. In TRI\_FACILITY, select TRI Facility ID (and state abbreviation, if desired)
  11. In V\_TRI\_REPORTING\_FORM, select Chemical Name, Trichemical ID, and CAS chem name
  12. In V\_TRI\_SUBMISSION\_NAICS, select NAICS Codes
  13. Click the Step 4 button
  14. In the NAICS Code section, enter the NAICS code in the Search Value box, and select TRI CHEM ID
- B. Save this file to disc and then open up in EXCEL. Resave the file as "Naics#####TRIconsum#.xls" where the first #s are for the code and the last number is for the number query that has been performed on that code (usually 1).
- C. Perform the sequence below to manipulate data in this new EXCEL file for entry into the crosswalk:
  1. Copy CAS and chem name to two new columns, then select both columns and sort based on CAS (rename CAS sort) (when selecting the two columns, do not include the mixture numbers and mixture names, only the chemicals with a CAS (these may be used at a later time)).

2. Select CAS sort, select menu item Data / Pivot table and chart, drop CAS sort into Rows, then into Drop Data Items. This gives a new sheet with all chemical CAS's plus the number of entries for each. Call the new sheet CAS frequency.
3. Do the same thing as step 3 for the first column of facility IDs in the original sheet to get the total number of facilities reporting, name the new sheet Facility Count, type that number in a cell at the top of the sheet for future reference (# total facilities).
4. On the CAS frequency sheet, copy both columns to new columns, in another column type in a function for the first row to calculate %probability that a facility will report a certain chemical (e.g.,  $=100 \times D2 / \# \text{facilities}$ , where #facilities is the integer value obtained from step 3). Calculate this for the entire chemical list by selecting that formula cell and stretching it down the entire column to include all cells that are in rows with chemicals (release the mouse and it will automatically fill the cells). Name that new row %prob facility.
5. Now create a matching row of chemical names to correspond to the trimmed-down CAS names. Create a new sheet and call it "Filtered list of names," copy the sorted column of names (next to CAS sort) from the original sheet, and paste it in this new sheet. Select the column and select from the menu Data/Filter/Advanced Filter/and then check the box for unique records only. This will replace the list of names with a filtered list that will list each name only once.
6. Go to the CAS frequency sheet, copy the CAS column and paste the copy into a new column to the right of the % prob facility column, call it CAS, copy the column created in step 5 to the column to the right of CAS and call it sorted CAS name, copy the %prob facility column and paste special - value only- to the column next to sorted CAS name and call it %prob facility value only. Review the three columns to ensure that the number of entries match and that the names and CAS's also match up. Two names could be assigned to the same CAS, which would result in an extra name entry in the sorted names.
7. Now copy the three columns of data to the crosswalk: count the number of rows (chemicals in the CAS sort column), and insert that many rows into the crosswalk after the appropriate code in the two columns titled NAICS\_Code and NAICS\_Title. Select the number of cells to insert at the point of insertion and choose from menu insert/cells/shift cells down). Paste that code and name to all the cells just created (copy two cells, select all new cells, paste). Back in CAS Frequency, copy the three columns of data (be careful to only select the rows to copy), and paste

into the crossover next to the newly created cells. Convert the new text to Arial font, and verify that the sheets are cleaned up, saved, and named properly for future reference.

## **Appendix C: Standard Operating Procedures (SOPs) Used in This Study**

### **Acronyms and abbreviations used in this appendix**

atm: Atmosphere(s)

BCF: Bioconcentration Factor

C: Centigrade

cm: centimeter(s)

cu: cubic

deg: degree(s)

EPA: United States Environmental Protection Agency

EPA RfC: EPA reference concentration

EPA RfD: EPA reference dose

FW: Fresh Water

h: hour

hc: Human Health Carcinogen

IDLH: Immediately Dangerous to Life and Health

kg: Kilogram

Koc: Organic carbon adsorption coefficient

Kow: Octanol Water Coefficient

l: liter

m: meter

MCL: Maximum Contaminant Level

m: meter(s)

mg: milligram(s)

mm: millimeter(s)

mol: mole(s)

Pa: Pascal(s)

ppb: parts per billion

ppm: parts per million

PQL: Practical Quantification Limit

PRG: Preliminary Remediation Goal

UTS: Universal Treatment Standards

ug: microgram

**Acrolein 107-02-8****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** Acrolein emitted to air reacts primarily with photochemically generated hydroxyl radicals in the troposphere (Ghilarducci & Tjeerdema, 1995). Minor processes include direct photolysis, reaction with nitrate radicals, and reaction with ozone (Atkinson et al., 1987; Haag et al., 1988a; Howard, 1989; BUA, 1994). Acrolein has been detected in rainwater, indicating that it may be removed by wet deposition (Grosjean & Wright, 1983). The calculated atmospheric half-life of acrolein, based on rate constants for hydroxyl radical reaction, is between 3.4 and 33.7 h (Atkinson, 1985; Edney et al., 1986b; Haag et al., 1988a; Howard, 1989; Howard et al., 1991; BUA, 1994). The overall reactivity-based half-life of acrolein in air, as estimated by Mackay et al. (1995), is less than 10 h. Based on these short estimated half-lives, acrolein is not a candidate for long-range atmospheric transport.

**Water:** Acrolein is removed from surface water primarily by reversible hydration, biodegradation by acclimatized microorganisms, and volatilization (Irwin, 1987; Haag et al., 1988b; Howard, 1989; ATSDR, 1990; Springborn Laboratories, 1993). In groundwater, acrolein is removed by anaerobic biodegradation and hydrolysis (Chou & Spanggord, 1990a). The overall reactivity-based half-life of acrolein in surface water is estimated to be between 30 and 100 h (Mackay et al., 1995). In groundwater, half-lives of 11 days and 336–1344 h (14–56 days) are estimated based on aerobic and anaerobic degradation, respectively (Howard et al., 1991). Observed dissipation half-lives of acrolein applied as a herbicide in irrigation range from 7.3 to 10.2 h (Jacobson & Gresham, 1991a, b, c; Nordone et al., 1996a). The relatively short observed half-lives of acrolein in surface waters make long-range aquatic transport unlikely. However, groundwater contamination could result in contaminant exposures.

**Sediment:** In sediment/water systems, acrolein undergoes hydrolysis, self-oxidation, and biodegradation. Experimental half-lives of 7.6 h and 10 days were determined for aerobic and anaerobic conditions, respectively (Smith et al., 1995). An overall reactivity-based half-life is estimated by Mackay et al. (1995) to be between 100 and 300 h. Because of its low organic carbon/water partition coefficient ( $K_{oc}$ ) and high water solubility, acrolein is not expected to significantly adsorb to suspended solids or sediments, nor are these suspended solids or sediments expected to significantly absorb acrolein from water (Irwin, 1988; Howard, 1989).

**Soil:** In the terrestrial environment, acrolein undergoes biodegradation, hydrolysis, volatilization, and irreversible sorption to soil (Irwin, 1988; Howard, 1989; Chou & Spanggord, 1990b). These processes are expected to significantly decrease the high infiltration rate of acrolein estimated from its low experimental  $K_{oc}$  (Irwin, 1988). The overall reactivity-based half-life of acrolein in soil is estimated to be between 30 and 100 h (Mackay et al., 1995).

**Biota:** Based on the high water solubility, low octanol/ water partition coefficient ( $K_{ow}$ ), and high reactivity of acrolein, uptake by organisms is predicted to be low. A bioconcentration factor (BCF) of 344 and a half-life of greater than 7 days were reported for acrolein in bluegill (*Lepomis macrochirus*) following exposure to acrolein at a mean concentration of 13  $\mu\text{g/l}$  for a 28-day period (Barrows et al., 1980). However, these values may be overestimates, as the total  $^{14}\text{C}$  measured in the fish may have included metabolites. A lower BCF of 0.6 was estimated using the linear regression equation of Veith et al. (1980) and a log  $K_{ow}$  of 0.01 for acrolein. Acrolein was not detected in the tissues of fish and shellfish sampled 1 day after a second exposure to [ $^{14}\text{C}$ ]acrolein in water (0.02 and 0.1 mg/litre for the first and second exposures, respectively) over a 1-week period. The presence of metabolites indicates that these species were able to rapidly metabolize acrolein and its residues (Nordone et al., 1998). Based on these results and the low reported BCFs, acrolein is unlikely to bioaccumulate or bioconcentrate significantly in aquatic organisms (Howard, 1989; ATSDR, 1990; DFO, 1995; Nordone et al., 1996b). Absorption of acrolein by terrestrial plants is poor (WSSA, 1983).

**Environmental partitioning :** Fugacity modeling was conducted to characterize key reaction, intercompartment, and advection (movement out of a system) pathways for acrolein and its overall distribution in the environment. A steady-state, non-equilibrium model (Level III fugacity model) was run using the methods developed by Mackay (1991) and Mackay & Paterson (1991). Assumptions, input parameters, and results are presented in Mackay et al. (1995) and summarized here. Values for input parameters were as follows: molecular mass, 56.06; melting point, 186.95  $^{\circ}\text{C}$ ; water solubility, 208 g/l; vapour pressure, 36.5 kPa at 20  $^{\circ}\text{C}$ ; log  $K_{ow}$ , 0.01; Henry's law constant, 9.8 Pa@m<sup>3</sup>/mol; half-life in air, 5 h; half-life in water, 55 h; half-life in soil, 55 h; half-life in sediments, 170 h. Modeling was based on an assumed default emission rate of 1000 kg/h into a region of 100 000 km<sup>2</sup>, which includes a surface water area (20 m deep) of 10 000 km<sup>2</sup>. The height of the atmosphere was set at 1000 m. Sediments and soils were assumed to have an organic carbon content of 4% and 2% and a depth of 1 cm and 10 cm, respectively. The estimated percent distribution predicted by this model is not affected by the assumed emission rate. Results of the modeling indicate that acrolein behaves differently depending on the medium to which it is released. Generally, when acrolein is continuously discharged into a specific medium, most of it can be expected to remain in

that medium. For example, if discharged into air, almost all of it will exist in the atmosphere, with very small amounts in soil and water. The same applies for discharge to water and soil (Mackay et al., 1995). These predicted distributions suggest that acrolein does not tend to partition from one compartment to another. It could also be possible that when acrolein does partition to another compartment, its persistence in that second compartment is so short that little remains there.

## Regulatory

MCL (Drinking Water): None

Universal Treatment Standards (UTS): 0.29 mg/l<sup>2</sup> in wastewater standards

Controlling Water Quality Criterion for Surface Water:

Basis: Aquatic Toxicity Criterion Concentration: 21 ug/L Criterion Quantitation Limit: 5 ug/L

EPA Rf D: 0.0005 mg/kg-day

EPA RfC: 0.000002 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
0.1	0.345	0.21	0.042

Soil PQL: 0.5 mg/kg

## Remediation

Due to its high reactivity, acrolein does not tend to persist in the environment, and its intercompartmental movement is small. Therefore, there does not appear to be much information on its treatment. For short-term spills, adsorbent could be an effective approach. If soil or groundwater should occur, and clean-up is deemed necessary, acrolein does appear to be susceptible to bioremediation.

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**Acrylonitrile      107-13-1****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** Acrylonitrile emitted to air reacts primarily with photochemically generated hydroxyl radicals in the troposphere (Atkinson et al., 1982; Edney et al., 1982; Munshi et al., 1989; US DHHS, 1990; Bunce, 1996). The atmospheric half-life, based on hydroxyl radical reaction rate constants, is calculated to be between 4 and 189 hours (h) (Callahan et al., 1979; Cupitt, 1980; Edney et al., 1982; Howard, 1989; Grosjean, 1990b; Kelly et al., 1994). Modeling of environmental partitioning is based on a mean half-life for acrylonitrile in air of 55 h. The reaction of acrylonitrile with ozone and nitrate is slow, because of the absence of chlorine and bromine atoms in the molecule, and is not likely to constitute a major route of degradation (Bunce, 1996). The reaction of hydroxyl radicals with acrylonitrile yields formaldehyde and, to a lesser extent, formic acid, formyl cyanide, carbon monoxide, and hydrogen cyanide (Edney et al., 1982; Spicer et al., 1985; Munshi et al., 1989; Grosjean, 1990a).

**Water:** Acrylonitrile in water can be biodegraded by acclimatized microorganisms or volatilized (Going et al., 1979). In water, half-lives of 30–552 h are estimated based on aqueous aerobic biodegradation (Ludzack et al., 1961; Going et al., 1979; Howard et al., 1991). Modeling of environmental partitioning based on a mean half-life for acrylonitrile in water of 170 h (7 days). The half-life based on volatilization is 1–6 days (Howard et al., 1991). The hydrolysis of acrylonitrile is slow, with half-lives under acidic and basic conditions of 13 and 188 years, respectively (Ellington et al., 1987). Acrylonitrile has an initial inhibitory effect on activated sludge systems and other microbial populations and does not meet the criteria of Organization for Economic Co-operation and Development (OECD) Test Method 301C for ready biodegradability (Chemicals Inspection and Testing Institute of Japan, 1992; AN Group, 1996; BASF AG, 1996). However, acrylonitrile will be extensively degraded (95–100%) following a short acclimation period if emitted to wastewater treatment plants (Tabak et al., 1980; Kincannon et al., 1983; Freeman & Schroy, 1984; Watson, 1993).

**Soil and sediment:** Acrylonitrile is biodegraded in a variety of surface soils (Donberg et al., 1992) and by isolated strains of soil bacteria and fungi (Wenzhong et al., 1991). Concentrations of acrylonitrile up to 100 mg/kg were degraded in under 2 days (Donberg et al., 1992). Similar breakdown by microbial populations present in sediment is likely (DMER & AEL, 1996; EC, 2000). Results of experimental studies (Zhang et al., 1990) or soil sorption coefficients calculated by quantitative structure–activity relationships (Koch & Nagel, 1988; Walton et al., 1992) or based on water solubility (Kenaga, 1980) indicate

little potential for adsorption of acrylonitrile to soil or sediments. A half-life of acrylonitrile in soil of 6–7 days has been reported (Howard et al., 1991; Donberg et al., 1992). Based on biodegradability and the soil partition coefficient, the half-life of acrylonitrile in soil was classified in the category of 300 days (EC, 2000). Modeling of environmental partitioning is based on a mean half-life for acrylonitrile in soil of 170 h (7 days). The half-life in the oxic zone of sediment can be assumed to be similar.

**Biota:** Bioaccumulation of acrylonitrile is not anticipated, given experimentally derived values of the octanol/water partition coefficient (log Kow) ranging from 0.92 to 1.2 (mean 0.25) (Tanii & Hashimoto, 1984; Sangster, 1989) and a log bioconcentration factor (log BCF) of 0 calculated from the water solubility of acrylonitrile (EC, 2000). Log BCF values were 0.48–1.68 in bluegill (*Lepomis macrochirus*) (Barrows et al., 1980) and rainbow trout (*Oncorhynchus mykiss*) (Lech et al., 1995). The experimentally derived log BCF of 1.68 reported by Barrows et al. (1980) in whole-body tissue of bluegill may be an overestimate, due to uptake of <sup>14</sup>C-labelled degradation products in addition to acrylonitrile and to cyanoethylation of macromolecules (EC, 2000).

**Environmental partitioning:** Fugacity modeling was conducted to characterize key reaction, intercompartment, and advection (movement out of a compartment) pathways for acrylonitrile and its overall distribution in the environment. A steady state, non-equilibrium model (Level III fugacity model) was run using the methods developed by Mackay (1991) and Mackay & Paterson (1991). Values for input parameters were as follows: molecular mass, 53.06 g/mol; water solubility, 75.5 g/litre; vapour pressure, 11.0 kPa; log Kow, 0.25; Henry's law constant, 11 Pa@m<sup>3</sup>/mol; half-life in air, 55 h; half-life in water, 170 h; half-life in soil, 170 h; half life in sediments, 550 h. Modeling was based on an assumed default emission rate of 1000 kg/h into a region of 100 000 km<sup>2</sup>, which includes a surface water area (20 m deep) of 10 000 km<sup>2</sup>. The height of the atmosphere was set at 1000 m. Sediments and soils were assumed to have an organic carbon content of 4% and 2% and a depth of 1 cm and 10 cm, respectively. The estimated percent distribution predicted by this model is not affected by the assumed emission rate. Modeling indicates that when acrylonitrile is continuously discharged into a specific medium, most of it (84–97%) can be expected to be present in that medium (DMER & AEL, 1996). More specifically, Level III fugacity modeling by DMER & AEL (1996) predicts that: when released into air, the distribution of mass is 92.8% in air, 6.4% in water, 0.8% in soil, and 0.0% in sediment; when released into water, the distribution of mass is 2.5% in air, 97.3% in water, 0.0% in soil, and 0.1% in sediment; and when released into soil, the distribution of mass is 4.4% in air, 11.9% in water, 83.7% in soil, and 0.0% in sediment. The major removal mechanisms in air, water, and soil are reaction within the medium and, to a lesser degree, advection and volatilization. Abiotic and biotic degradation in various compartments result in low persistence overall and little if any

bioaccumulation. Owing to the paucity of data on concentrations of acrylonitrile in environmental media, fugacity modeling with version 4 of the ChemCAN3 model (Mackay et al., 1995) was also conducted with the conservative assumption that all known releases in 1996 (Environment Canada, 1997) in Canada occurred in southern Ontario. Release to air was considered to be approximately 19 tonnes per year, with simultaneous release to water of 0.53 tonnes per year. Since the half-life of acrylonitrile in air is the major determinant of its fate in the environment, the model was run using the minimum, median, and maximum half-life values (4, 55, and 189 h) under summer, winter, and year-round conditions. Modeling predicted distribution primarily to air (41.9–78.1%) and water (21.6–57.9%).

## Regulatory

MCL (Drinking Water): None

UTS: 0.24 mg/L<sup>2</sup> in wastewater standards and 84 mg/kg<sup>3</sup> non wastewater standards

Soil PQL: 0.5 mg/kg

EPA Rf D: 0.001 mg/kg/d provisional

EPA Rf C: 2 µg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
0.21	0.51	0.028	0.039

Fresh Water (FW) Criteria: Human Health: 0.051 ug/L (hc) Saline Water Criteria: 0.25 ug/L (hc) 'Human health carcinogen'

## Remediation

Land spill: Dig a pit, pond, lagoon, or holding area to contain liquid or solid material. /SRP: If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash, cement powder, sawdust, or commercial sorbents. Apply "universal" gelling agent to immobilize spill.

Air spill: Apply water spray or mist to knock down vapors. Combustion products include corrosive or toxic vapors.

Water spill: Use natural barriers or oil spill control booms to limit spill travel. Use surface

active agent (eg, detergent, soaps, or alcohols) if approved by EPA. Inject "universal" gelling agent to solidify encircled spill and increase effectiveness of booms. Add calcium hypochlorite. If dissolved, in region of 10 ppm or greater concentration, apply activated carbon at ten times the spilled amount. Use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates (Association of American Railroads).

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**Allyl Alcohol      107-18-6****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Soil:** An estimated Koc value of 1.3, determined from a structure estimation method (Meylan et al. 1992), indicates that allyl alcohol is expected to have very high mobility in soil (Swann et al. 1983). Leaching of allyl alcohol from soil surface to deeper layers was found to rank in order of sand > sandy loam > humus sand (Scheunert et al. 1981). The percent leached after 2 days and 400 ml of water with sand containing 0.51% and 2.89% organic matter were 100 and 83.3%, respectively (Scheunert et al. 1981). Freundlich adsorption coefficients for allyl alcohol in Texas soil (pH 7.8, 3.25% organic matter) and Mississippi soil (pH 4.8, <1% organic matter) were  $4.5 \times 10^{-3}$  and  $3.3 \times 10^{-4}$ , respectively (Loehr 1989). Volatilization of allyl alcohol from moist soil surfaces is expected to be an important fate process given a Henry's Law constant of  $4.99 \times 10^{-6}$  atm-cu m/mole (Scheunert et al. 1981). The potential for volatilization of allyl alcohol from dry soil surfaces may exist based upon a vapor pressure of 26.1 mm Hg (Daubert and Danner 1989). Half-lives of 10.2 and 9.5 days at 20 deg C were found for allyl alcohol with Texas soil (pH 7.8, 3.25% organic matter) and Mississippi soil (pH 4.8, <1% organic matter), respectively (Loehr 1989).

**Water:** An estimated Koc value of 1.3, determined from a structure estimation method (Meylan et al 1992), indicates that allyl alcohol is not expected to adsorb to suspended solids and sediment (Swann et al. 1983). Volatilization from water surfaces is expected (Lyman et al.1990) based upon a Henry's Law constant of  $4.99 \times 10^{-6}$  atm-cu m/mole (Hine and Mookerjee 1975). Using this Henry's Law constant and an estimation method (Lyman et al.1990), volatilization half-lives for a model river and model lake are 5.7 and 44 days, respectively. According to a classification scheme (Franke et al.1994), an estimated BCF of 3.1, from its log Kow of 0.17 (Hansch et al 1995) and a regression-derived equation (Meylan et al. 1999), suggests the potential for bioconcentration in aquatic organisms is low. Allyl alcohol is not expected to undergo hydrolysis in the environment due to the lack of hydrolyzable functional groups (Lyman et al.1990). In a biodegradation study, allyl alcohol (25 ppm) was found to degrade 100 and 60% in marine and river water after 3 days at 30 deg C, respectively (Kondo et al. 1988).

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), allyl alcohol, which has a vapor pressure of 26.1 mm Hg at 25 deg C (Daubert and Danner 1989), is expected to exist solely as a vapor in the ambient atmosphere. Vapor-phase allyl alcohol is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals; the half-life for this reaction in air is

estimated to be 13 hours, calculated from its rate constant of  $3.0 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C that was derived using a structure estimation method (Meylan and Howard 1993). The rate constant for the vapor-phase reaction of allyl alcohol with ozone has been reported as  $14.4 \times 10^{-18}$  cu cm/molecule-sec at 12-22 deg C (Grosjean and Grosjean 1994). This corresponds to an atmospheric half-life of about 0.8 days, at an atmospheric concentration of  $7 \times 10^{11}$  ozone molecules per cu cm (Atkinson and Carter 1984). While allyl alcohol in water was found to degrade to carbon dioxide upon exposure to UV light from a high pressure mercury-vapor lamp through a quartz tube (Knoevenagel and Himmelreich 1976); no data were found regarding the direct photolysis of allyl alcohol by sunlight in air.

**Biota:** An estimated BCF of 3.1 was calculated for allyl alcohol (SRC), using a log Kow of 0.17 (Hansch et al. 1995) and a regression-derived equation (Meylan et al. 1999). According to a classification scheme (Franke et al. 1994), this BCF suggests the potential for bioconcentration in aquatic organisms is low (SRC). The uptake of  $^{14}\text{C}$ -allyl alcohol residues by lettuce and carrots was investigated in the greenhouse (Scheunert et al. 1981). Uptake of residues was higher by carrots than by lettuce, and higher by lettuce roots than by lettuce tops (Scheunert et al. 1981). No bioaccumulation was observed.

**Environmental Partitioning:** Using a structure estimation method based on molecular connectivity indices (Meylan et al. 1992), the Koc for allyl alcohol can be estimated to be 1.3 (SRC). This estimated Koc value suggests that allyl alcohol is expected to have very high mobility in soil (Swann et al. 1983). Leaching of allyl alcohol from soil surface to deeper layers was found to rank in order of sand > sandy loam > humus sand (Scheunert et al. 1981). The percent leached after 2 days and 400 ml of water with sand containing 0.51% and 2.89% organic matter were 100 and 83.3%, respectively (3). Freundlich adsorption coefficients for allyl alcohol in Texas soil (pH 7.8, 3.25% organic matter) and Mississippi soil (pH 4.8, <1% organic matter) were  $4.5 \times 10^{-3}$  and  $3.3 \times 10^{-4}$  respectively (Loehr 1989).

Allyl alcohol is not persistent in the environment. It oxidizes at a moderate rate in the atmosphere and there is also some potential for photolytic degradation. Allyl alcohol is expected to biodegrade under aerobic and anaerobic conditions. It is not susceptible to hydrolysis. Fugacity modeling indicates that over 95% of allyl alcohol in the environment partitions to water. Very little allyl alcohol will adhere to soil or sediment and it is not expected to bioaccumulate in food chains.

## Regulatory

MCL (Drinking Water): None

Universal Treatment Standards (UTS): None

EPA Rf D: 0.0005 mg/kg-day

EPA RfC: 0.001 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
310	4400	18	180

## Remediation

Land spill: Dig a pit, pond, lagoon, or holding area to contain liquid or solid material. /SRP: If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner./ Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquids with fly ash, cement powder, sawdust, or commercial sorbents. Apply "universal" gelling agent to immobilize spill.

Water spill: If dissolved, in region of 10 ppm or greater concentration apply activated carbon at ten times the spilled amount. Use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates.

Air spill: Apply water spray or mist to knock down vapors (Association of American Railroads) cited in HSDB.

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## Carbon disulfide 75-15-0

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** In air, carbon disulfide is primarily degraded through photo-oxidation by reactions with hydroxyl (OH) radicals and by a secondary route involving triplet oxygen ( $O(^3P)$ ). With a hydroxyl radical concentration of  $5 \times 10^5$  radicals/cm<sup>3</sup>, a half-life of about 5.5-15 days is calculated from rate constants between  $1.1 \times 10^{-12}$  and  $2.9 \times 10^{-12}$  cm<sup>3</sup>/molecule per second (BUA, 1993). Wine et al. (1981) likewise estimated that photo-oxidation in the troposphere results in a half-life in air of 7-14 days. Reaction products include carbonyl sulfide (COS) and sulfur dioxide (SO<sub>2</sub>). Carbonyl sulfide has a much longer lifetime (2 years) than carbon disulfide in the atmosphere.

Photolysis of carbon disulfide by radiation at wavelengths above 290 nm occurs in the troposphere. An atmospheric lifetime of 11 days (half-life of 7.7 days) was calculated assuming 12 h of sunlight (Peyton *et al.*, 1976). Wood and Heicklen (1971) demonstrated that direct photolysis of carbon disulfide at 313 nm produces reaction products similar to those of the photo-oxidation reaction - that is, carbon monoxide (CO), carbonyl sulfide, sulfur dioxide plus an unidentified polymeric material. Wet deposition from the atmosphere is probably a minor removal process, because carbon disulfide is interacted only weakly with water (Lovejoy, 1989).

**Water:** With a Henry's law constant of 1748 Pa m<sup>3</sup>/mol at 20°C and a vapor pressure of 48.2 kPa at 20°C, the major fate process for carbon disulfide released into water is volatilization, with a half-life ranging between 11 minutes in water (saturated solution) and 2.6 hours in a model river (Peyton *et al.*, 1976; Howard, 1989). Carbon disulfide is resistant to hydrolysis in water within the biological pH range (4-10), with a hydrolysis half-life extrapolated to pH 9 of 1.1 years (Peyton *et al.*, 1976). Its predicted rate of biodegradation in water is negligible compared with its rate of volatilization from surface water (ATSDR, 1996). The mean degradation half-life used for fugacity modeling by DMER and AEL (1996) of 5500 hours (7.4 months) was based on the estimate of biodegradation half-life by Abrams *et al.* (1975).

**Sediment:** Owing to its low affinity for sorption to organic substances (organic carbon/water partition coefficient [ $\log K_{oc}$ ] = 1.79), very little carbon disulfide is likely to partition to or remain in sediment. One study indicated that the soil/sediment microorganism *Thiobacillus thiorapax* (grown aerobically, incubated anaerobically) was able to metabolize carbon disulfide to produce carbonyl sulfide and hydrogen sulfide (Smith and Kelly, 1988). Thus, some biodegradation is expected to occur. The estimated

mean reactivity half-life used for fugacity modeling was 5500 hours (7.4 months), based on the estimate of biodegradation half-life by Abrams *et al.* (1975).

**Soils:** No estimates of a half-life for carbon disulfide in soil were identified in the literature. Aerobic degradation of carbon disulfide has been observed with a strain of *Thiobacillus thiorapus*. This particular strain was able to hydrolytically oxidize carbon disulfide sequentially to carbonyl sulfide and hydrogen sulfide; all the carbon was released as carbon dioxide, followed by oxidation of the sulfide to sulfate (Smith and Kelly, 1988). For soil, DMER and AEL (1996) used a mean degradation half-life of 5500 hours for their fugacity modeling, based on the estimate of biodegradation half-life by Abrams *et al.* (1975). In the natural environment, carbon disulfide is highly mobile in soil ( $\log K_{oc} = 1.79$ ) and is subject to rapid volatilization, so it is unlikely to remain in soil long enough to undergo significant biodegradation.

**Biota:** Carbon disulfide is expected to have little or no tendency to bioaccumulate or biomagnify in biota, owing to its relatively low octanol/water partition coefficient ( $\log K_{ow}$ ) value (2.14) and rapid metabolism in most animals (Beauchamp *et al.*, 1983).

**Environmental Partitioning:** Fugacity modeling was carried out to provide an overview of key reaction, intercompartment and advection (movement out of a system) pathways for carbon disulfide and its overall distribution in the environment (DMER and AEL, 1996). A steady-state, non-equilibrium EQC model (Level III fugacity modeling) was run using the methods developed by Mackay (1991) and Mackay and Paterson (1991). Values for input parameters were as follows: molecular weight, 76.1 g/mol; water solubility, 2100 mg/L; vapor pressure, 48 210 Pa;  $\log K_{ow}$ , 2.14; Henry's law constant, 1748 Pa m<sup>3</sup>/mol; half-life in air, 170 hours; half-life in water, soil and sediment, 5500 hours. Modeling was based on an assumed default emission rate of 1000 kg/hour into a region of 100 000 km<sup>2</sup>, which includes a 10 000-km<sup>2</sup> area of surface water (20 m deep). The height of the atmosphere is 1000 m. Sediments and soils have an organic carbon content of 4% and 2% and a depth of 1 cm and 10 cm, respectively. The estimated percent distribution predicted by this model is not affected by the assumed emission rate.

Modeling indicates that carbon disulfide partitions differently depending on the medium to which it is released. For example, if emitted into air, 99.8% of the carbon disulfide is present in air; if emitted into soil, the fraction in air is reduced to 73%, with most of the rest in soil. When carbon disulfide is released to water, it is present primarily in water (85%) and, to a lesser extent, in air (15%) (DMER and AEL, 1996). Thus, while the predicted distributions suggest that little intermedia transport will occur when carbon disulfide is discharged to air, release to each of soil and (to a lesser extent) water has the potential for substantial transport of carbon disulfide to air.

## Regulatory

MCL (Drinking Water): None

GPS - Groundwater protection standard (DWEL): 3,650 ug/L

Universal Treatment Standards (UTS): 3.8 mg/l<sub>2</sub> in wastewater standards and 4.8mg/L TCLP

Controlling Water Quality Criterion for Surface Water: (WDEQ)

Basis: Aquatic Toxicity Criterion Concentration: 21 ug/L Criterion Quantitation Limit: 5 ug/L

EPA Rf D: 0.1 mg/kg-day

EPA RfC: 0.7 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
360	720	730	1000

## Remediation

Remove all ignition sources. Ventilate area of spill or leak. For small quantities, absorb on paper towels and evaporate in a safe place (such as a fume hood). Allow sufficient time for evaporating of vapors to completely clear the hood ductwork. Burn the paper in a suitable location away from combustible materials. Large quantities can be reclaimed or collected and atomized in a suitable combustion chamber equipped with an appropriate effluent gas cleaning device. Carbon disulfide should not be allowed to enter a confined space, such as a sewer, because of the possibility of an explosion.

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## Chloroethylene (Vinyl Chloride) 75-01-4

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), vinyl chloride, which has a vapor pressure of 2,980 mm Hg at 25 deg C (Daubert and Danner 1989), is expected to exist solely as a gas in the ambient atmosphere. Gas-phase vinyl chloride is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 55 hours, calculated from its rate constant of  $6.96 \times 10^{-12}$  cu cm/molecule-sec at 25 deg C (Atkinson 1989). Vinyl chloride is not expected to undergo considerable direct photolysis since this compound does not absorb light appreciably in the environmental UV spectrum.

**Water:** Based on a classification scheme (Swann et al 1983), an estimated Koc value of 57(SRC), calculated from a water solubility of 2,700 mg/l (Riddick et al 1986) and a regression derived equation (Lyman et al 1990) indicates that vinyl chloride is not expected to adsorb to suspended solids and sediment in water (SRC). Volatilization from water surfaces is expected to occur rapidly (Lyman et al 1990) based upon a Henry's Law constant of 0.0278 atm-cu m/mole (Gossett 1987). Using this Henry's Law constant and an estimation method (Lyman et al 1990), volatilization half-lives for a model river and model lake are 1 hour and 3 days, respectively (SRC). According to a classification scheme (Franke et al 1994), a BCF value of less than 10 measured in fish (Freitag et al 1985), suggests bioconcentration in aquatic organisms is low(SRC). The biodegradation half-life of vinyl chloride in aerobic and anaerobic waters was reported as 28 and 110 days, respectively (Capel and Larson 1995). Hydrolysis is not expected to be an important environmental fate process based on hydrolysis half-lives of 9.91 years (pH = 7, 25 deg C) and 10.7 years (pH = 7, 10 deg C). Vinyl chloride may undergo indirect photolysis in natural waters when photosensitizers such as humic material are available (Mill 1999).

**Soil/Sediment:** An estimated Koc value of 57(SRC), calculated from a water solubility of 2,700 mg/l (Riddick et al 1986) and a regression derived equation (Lyman et al 1990) indicates that vinyl chloride is expected to have high mobility in soil (Swann et al 1983). Volatilization of vinyl chloride from moist soil surfaces is expected to be an important fate process (SRC) given a Henry's Law constant of 0.0278 atm-cu m/mole (Gossett 1987). Vinyl chloride may volatilize from dry soil surfaces based on a vapor pressure of 2,780 mm Hg at 25 deg C (Daubert and Danner 1989). The volatilization half-life of vinyl chloride was estimated as 0.2 days when incorporated in a soil at a depth of 1 cm and 0.5 days at a depth of 10 cm (Jury et al 1984).

**Biota:** Some /data indicated/ that vinyl chloride is too readily volatilized to undergo bioaccumulation, except perhaps in the most extreme exposure conditions. Studies on five bacterial, three fungal, and two single organism cultures from natural aquatic systems did not show bioaccumulation to be an appreciable process (Callahan et al 1979). The BCF value of vinyl chloride in golden ide fish was reported as less than 10 (Freitag et al 1985). The BCF value of vinyl chloride in green algae was reported as 40. According to a classification scheme (Franke et al 1994), the BCF data suggest that bioconcentration in aquatic organisms is low (SRC).

## Regulations

MCL (Drinking Water): 0.002 mg/L

GPS - Groundwater protection standard (DWEL): 0.002 mg/L

Universal Treatment Standards (UTS): 0.27 mg/l<sub>2</sub> in wastewater standards and 6.0 mg/kg<sup>3</sup>

Groundwater Water Quality Criterion : 0.27ug/L

PQL: 1ug/L

EPA Rf D: 0.002 mg/kg-day

EPA RfC: 0.1 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
0.15	0.83	0.22	0.041

## Remediation

**Land Spill:** Construct barriers to contain spill. Absorb small amounts of spill with natural or synthetic sorbents, shovel into containers with covers.

**Water Spill:** Contain contaminated water with dams or natural barriers. (Environment Canada)

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## **Chlorohydrin (3-Chloro-1,2-Propanediol) 96-24-2**

### **ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), 3-chloro-1,2-dihydroxypropane, which has a measured vapor pressure of 0.2 mm Hg at 20 deg C (Lande 1980), will exist solely as a vapor in the ambient atmosphere. Vapor-phase 3-chloro-1,2-dihydroxypropane is degraded in the atmosphere by reaction with photochemically produced hydroxyl radicals(SRC); the half-life for this reaction in air is estimated to be about 2 days (Meylan and Howard 1993).

**Water:** Based on a recommended classification scheme (Lyman et al.1990), an estimated Koc value of 1(SRC), determined from structure estimation method (Meylan et al. 1992), indicates that 3-chloro-1,2-dihydroxypropane should not adsorb to suspended solids and sediment in water (SRC). 3-Chloro-1, 2-dihydroxypropane is not expected to volatilize from water surfaces (Lyman et al. 1990, SRC) based on an estimated Henry's Law constant of  $6.1 \times 10^{-8}$  atm-cu m/mole (SRC), developed using a fragment constant estimation method (Meylan and Howard 1991). According to a classification scheme (Franke et al. 1994), an estimated BCF value of 0.2 (Lyman et al. 1990, SRC), from an estimated log Kow (Meylan and Howard 1995, SRC), suggests that bioconcentration in aquatic organisms is low (SRC). Data regarding the aerobic biodegradation of 3-chloro-1,2-dihydroxypropane in water are inconclusive; in screening tests, this compound reached 0 (Niemi et al. 1987) to 1% (Bridie et al. 1979) of the theoretical BOD in 5 days and 68% of the theoretical BOD in 14 days (Chemicals Inspection and Testing Institute). Under anaerobic conditions, 3-chloro-1,2-dihydroxypropane may be resistant to biodegradation; using an acclimated culture, only 26% of the initial 3-chloro-1,2-dihydroxypropane concentration was utilized in 90 days (Chou et al. 1987).

**Soil:** An estimated Koc value of 1 (SRC), determined from a structure estimation method (Meylan et al. 1992), indicates that 3-chloro-1,2-dihydroxypropane will have very high mobility in soil (Swann et al. 1983). Volatilization of 3-chloro-1,2-dihydroxypropane is not expected to be significant from moist soil surfaces (SRC) given an estimated Henry's Law constant of  $6.1 \times 10^{-8}$  atm-cu m/mole (SRC), using a fragment constant estimation method (Meylan and Howard 1991). Data regarding the aerobic biodegradation of 3-chloro-1, 2-dihydroxypropane in soil are variable; in screening tests, this compound reached 0 (Niemi et al. 1987) to 1% (Bridie et al. 1979) of the theoretical BOD in 5 days and 68% of the theoretical BOD in 14 days (Chemicals Inspection and Testing Institute). Under anaerobic conditions, 3-chloro-1, 2-dihydroxypropane may be resistant to

biodegradation. Using an acclimated culture, only 26% of the initial 3-chloro-1,2-dihydroxypropane concentration was utilized in 90 days (Chou et al. 1978).

**Biota:** An estimated BCF value of 0.2 was calculated for 3-chloro-1,2-dihydroxypropane (SRC), using an estimated log Kow of -0.53 (Meylan and Howard) and a recommended regression-derived equation (Lyman et al. 1990). According to a classification scheme (Franke et al. 1994), this BCF value suggests that bioconcentration in aquatic organisms is low.

## Regulatory

MCL (Drinking Water): none

Universal Treatment Standards (UTS): none

EPA RfD: discontinued May 1995

EPA RfC: discontinued May 1995

IDLH: 10 mg/m<sup>3</sup>

## Remediation

Criteria for land treatment or burial (sanitary landfill) disposal practices are subject to significant revision. Prior to implementing land disposal of waste residue (including waste sludge), consult with environmental regulatory agencies for guidance on acceptable disposal practices.

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**Cyanogen Chloride (CK) 506-77-4****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** Because hydrogen cyanide is a gas and has a relatively slow degradation rate in air, the atmosphere will be the ultimate sink for this compound. Almost all of the hydrogen cyanide released to the atmosphere remains in the lower altitudes (troposphere); only 2% of tropospheric hydrogen cyanide is transferred to the stratosphere (Cicerone and Zellner 1983). Cyanide has the potential to be transported over long distances from its emission source. Despite higher water solubility at saturated pressure, the removal of hydrogen cyanide by rainwater appears to be a negligible partitioning pathway (Cicerone and Zellner 1983). Because hydrogen cyanide is a gas, its removal from air by dry deposition is also likely to be negligible. However, metal cyanide particles, particularly water-soluble cyanide particles, are expected to be removed from the air by both wet and dry deposition.

**Water:** Volatilization and sorption are the two physical processes that contribute to the loss of cyanide from water. At  $\text{pH} < 9.2$ , most of the free cyanide in solution should exist as hydrogen cyanide, a volatile cyanide form (EPA 1978). On the basis of Henry's law constant and the volatility characteristics associated with various ranges of Henry's law constant (Thomas 1982), volatilization is a significant and probably dominant fate process for hydrogen cyanide in surface water (EPA 1992). The most common alkali metal cyanides (e.g., sodium and potassium cyanide) may also be lost from surface water primarily through volatilization; whereas, the sparingly soluble metal cyanides such as copper (I) cyanide are removed from water predominantly by sedimentation and biodegradation (EPA 1992). Variations in the volatilization rate are expected because this process is affected by several parameters including temperature, pH, wind speed, and cyanide concentration (EPA 1979). EPA (1979) summarized the unpublished results of a laboratory study that indicated that the volatilization half-life of hydrogen cyanide from solutions at concentrations of 25–200  $\mu\text{g/L}$  ranged from 22 to 110 hours. First-order kinetics were observed. In outdoor experiments with moderate winds, the rate of hydrogen cyanide loss increased by a factor of 2–2.5. In a study to evaluate the effect of cyanide on biochemical oxidation, there was a 50% loss of 6 ppm (mg/L) cyanide in river water kept in open biochemical oxygen demand bottles (without aeration) at pH 7.4 within  $\approx 10$  days. When the bottles were aerated (rate of aeration not given), 50% loss occurred in  $\approx 10$  hours. The kinetics of the rate of loss due to volatilization was not rigorously investigated. The volatilization rate was pH-dependent, with the rate faster at a lower pH. Data indicated that cyanide volatilization is a more important fate process than cyanide loss due to chemical and biodegradation reactions (Raef et al. 1977a). Because

volatilization is not an important fate process for cyanide in groundwater, cyanide would be expected to persist for considerably longer periods of time in underground aquifers than in surface water.

**Sediment:** Cyanides are sorbed by various natural media, including clays (Cruz et al. 1974), biological solids (Raef et al. 1977b), and sediments (EPA 1979). However, additional data are necessary to assess the significance of cyanide sorption to suspended solids and sediments in water. Hydrogen cyanide and the alkali metal cyanides are not likely to be strongly sorbed onto sediments and suspended solids because of their high water solubility. Soluble metal cyanides may show somewhat stronger sorption than hydrogen cyanide, with the extent of sorption increasing with decreasing pH and increasing iron oxide, clay, and organic material contents of sediment and suspended solids (EPA 1979). However, sorption is probably insignificant even for metal cyanides when compared to volatilization and biodegradation (EPA 1979, 1992).

**Soil:** Volatilization of hydrogen cyanide would be a significant loss mechanism for cyanides from soil surfaces at a pH <9.2. Cyanides are fairly mobile in soil. Mobility is lowest in soils with low pH and high concentrations of free iron oxides, positively charged particles, and clays (e.g., chlorite, kaolin, gibbsite), and highest in soils with high pH, high concentrations of free CaCO<sub>3</sub> and negatively charged particles, and low clay content (EPA 1979). Although cyanide has a low soil sorption capability, it is usually not detected in groundwater, probably because of fixation by trace metals through complexation or transformation by soil microorganisms (EPA 1978). In soils where cyanide levels are high enough to be toxic to microorganisms (i.e., landfills, spills), this compound may leach into groundwater (EPA 1984). Also, leaching of cyanide into a shallow aquifer can occur, as demonstrated by the high concentration of cyanide (1,200 µg/L) in groundwater sampled from the Biscayne Aquifer in Dade County, Florida, which lies below a solid waste site (Myers 1983).

**Biota:** There are no data available to indicate that simple metal cyanides and hydrogen cyanide bioconcentrate in aquatic organisms (EPA 1979, 1980, 1985, 1992).

Bioconcentration factors (BCFs) of 0.73 and 1.62 can be calculated for hydrogen cyanide, using the equation of Veith et al. (1979) for the BCF of a chemical in whole fish (log BCF, 0.85; log Kow, -0.70) and the log Kow values. Similarly, the calculated BCF for sodium cyanide is 0.47. There is some evidence that certain metal cyanide complexes bioaccumulate in aquatic organisms. Fish from water with soluble silver and copper cyanide complexes were found to have metal cyanides in their tissues at concentrations ranging up to 168 and 304 µg/g, respectively (wet or dry weight not specified) (EPA 1979). It is difficult to evaluate the toxicologic significance of bioaccumulation of metal cyanide complexes because these compounds are much less toxic than soluble hydrogen cyanide, sodium cyanide, or potassium cyanide (EPA 1992). There is no evidence of

biomagnification of cyanides in the food chain (EPA 1978). Accumulation of cyanide in food webs is not expected, considering the rapid detoxification of cyanide by most species and the lethal effects of large doses of cyanide (EPA 1978).

**Environmental partitioning:** No information could be found in the available literature on the transport and partitioning of cyanogen chloride in the environment, or its partitioning coefficients (K<sub>oc</sub>, K<sub>ow</sub>) or Henry's law constants. Like cyanogen, cyanogen chloride is a highly volatile gas. Therefore, it would be expected that volatilization from water and soil would be a primary route of environmental partitioning for both cyanogen and cyanogen chloride.

## Regulatory

MCL (Drinking Water) for cyanide: 0.2 mg/L

Universal Treatment Standards (UTS): None established

Controlling Water Quality Criterion for Surface Water:

Basis: Aquatic Toxicity Criterion Concentration: ug/L Criterion Quantitation Limit: ug/L

EPA Rf D: 25.3 mg/kg-day

EPA RfC:  $3 \times 10^{-3}$  mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
160	540	180	300

Volatile Organic Cmpd. Measured using Method 8260: Permit limit 2000 ug/L

## Remediation

Land spill: Dig a pit, pond, lagoon, or holding area to contain liquid or solid material.

SRP: If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash or cement powder. /Cyanogen chloride, inhibited/

Water spill: Add dilute caustic soda. Add calcium hypochlorite. Adjust pH to neutral (pH 7). Cyanogen chloride, inhibited.

Air spill: Apply water spray or mist to knock down vapors. Vapor knock down water is corrosive or toxic and should be diked for containment (Association of American Railroads).

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## **Dimethylamine 124-40-3**

### **ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 19881), dimethylamine, which has a vapor pressure of 1,520 mm Hg at 25 deg C (Daubert and Danner 1985), is expected to exist solely as a gas in the ambient atmosphere. Gas-phase dimethylamine is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals(SRC); the half-life for this reaction in air is estimated to be 6 hours (SRC), calculated from its rate constant of  $6.54 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C (Atkinson 1989). Dimethylamine does not contain chromophores that absorb at wavelengths >290 nm (Lyman et al 1990) and therefore is not expected to be susceptible to direct photolysis by sunlight (SRC).

**Water:** A Koc value of 508 in lake sediment (von Oepen et al 1991), indicates that dimethylamine is expected to adsorb to suspended solids and sediment (Swann et al 1983). A pKa of 10.73 (Perrin 1965) indicates diemthylamine will exist almost entirely in the cation form at pH values of 5 to 9 and therefore volatilization from water surfaces is not expected to be an important fate process (Doucette 2000). According to a classification scheme (Franke et al 1994), an estimated BCF of 3 (SRC), from its log Kow of -0.38 (Hansch et al 1995) and a regression-derived equation (Meylan et al 1999), suggests the potential for bioconcentration in aquatic organisms is low (SRC). Dimethylamine is expected to biodegrade in water surfaces based on a half-life of 1.6 days in Vistula River water (Warsaw, Poland) following a 0.3 day lag period (Dojlido 1979).

**Soil/Sediment:** Based on a classification scheme (Swann et al 1983), an average Koc value of 434.9 calculated from data of 5 soils (Rao and Davidson 1982), indicates that dimethylamine is expected to have moderate mobility in soil (SRC). The pKa of dimethylamine is 10.73 (Perrin 1965), indicating that this compound will exist almost entirely in the cation form in the environment; volatilization of dimethylamine from moist soil surfaces is not expected to be an important fate process given its cationic state (SRC). Dimethylamine is expected to volatilize from dry soil surfaces (SRC) based upon a vapor pressure of 1,520 mm Hg (Daubert and Danner 1985). Dimethylamine was biodegraded 69-89% in three Saskatchewan soils during a 7 day incubation period (Smith and Aubin 1992), suggesting that biodegradation may be an important environmental fate process in soil (SRC).

**Biota:** An estimated BCF of 3 was calculated in fish for dimethylamine (SRC), using a log Kow of -0.38 (Hansch et al 1995) and a regression-derived equation (Meylan et al; 1999).

According to a classification scheme (Franke et al 1994), this BCF suggests the potential for bioconcentration in aquatic organisms is low (SRC).

**Environmental Partitioning:** The rate constant for the vapor-phase reaction of dimethylamine with photochemically-produced hydroxyl radicals has been measured as  $6.54 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C (Atkinson 1989). This corresponds to an atmospheric half-life of about 6 h at an atmospheric concentration of  $5 \times 10^5$  hydroxyl radicals per cu cm (Atkinson 1989). Dimethylamine is not expected to undergo hydrolysis in the environment due to the lack of functional groups that hydrolyze under environmental conditions. Dimethylamine does not contain chromophores that absorb at wavelengths >290 nm (Lyman et al 1990) and therefore is not expected to be susceptible to direct photolysis by sunlight.

## Regulations

MCL (Drinking Water): None

GPS - Groundwater protection standard (DWEL): none

Universal Treatment Standards (UTS): none

Groundwater Water Quality Criterion: none

LAND BAN HEALTH BASED LIMITS: 0.005 mg/L

EPA RfD: none

EPA RfC: none

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
0.067	0.25	0.021	0.035

## Remediation

Land spill: Dig a pit, pond, lagoon, holding area to contain liquid or solid material. /SRP: If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike the surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash, cement powder, or commercial sorbents. Neutralize with sodium bisulfate ( $\text{NaHSO}_4$ )

Water spill: Add sodium bisulfate ( $\text{NaHSO}_4$ ). If dissolved, in region of 10 ppm or greater concentration, apply activated carbon at ten times the spilled amount. Use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates (Association of American Railroads).

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## Formaldehyde 50-00-0

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), formaldehyde, which has an estimated vapor pressure of 3,890 mm Hg at 25 deg C (Boublik et al 1984), will exist in the gas phase in the ambient atmosphere (SRC). Gas-phase formaldehyde is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is 41 hrs (SRC), calculated from its rate constant of  $9.4 \times 10^{-12}$  cu cm/molecule-sec at 25 deg C (Kwok and Atkinson 1994). The hydroxy radical initiated oxidation of formaldehyde also occurs in cloud droplets to form formic acid, a component of acid rain (Kwok and Atkinson 1994). Formaldehyde absorbs ultraviolet radiation at wavelengths of  $>360$  nm (Su et al 19795); therefore, formaldehyde may directly photolyze in sunlight (SRC). Formaldehyde has a half-life of 6 hrs in simulated sunlight (Su et al 1979). The predicted half-life of formaldehyde due to photolysis in the lower atmosphere is 1.6 hrs at a solar zenith of 40 degrees (Calvert et al 1972). Formaldehyde reacts with the  $\text{NO}_3$  radical by H-atom abstraction with a half-life of 12 days (assuming an average  $\text{NO}_3$  radical concentration of  $2 \times 10^9$ /cu cm) (Atkinson et al 1984).

**Water:** An estimated Koc value of 37 (SRC) determined from a log Kow of 0.35 (Hansch et al 1995) and a regression-derived equation (Lyman et al 1990), indicates that formaldehyde is not expected to adsorb to suspended solids and sediment (Swann et al 1983). Volatilization from water surfaces is not expected (Lyman et al 1990) based upon a Henry's Law constant of  $3.4 \times 10^{-7}$  atm-cu m/mole (Betterton and Hoffmann 1988). According to a classification scheme (Franke et al 1994), an estimated BCF of 3 (SRC), from its log Kow (Hansch et al 1995) and a regression-derived equation (Meylan et al 1999), suggests the potential for bioconcentration in aquatic organisms is low (SRC). Formaldehyde readily biodegrades under both aerobic and anaerobic conditions in the environment (SRC). Formaldehyde in aqueous effluent was degraded by activated sludge and sewage in 48-72 hr. In a die-away test using water from a stagnant lake, degradation was complete in 30 and 40 h under aerobic and anaerobic conditions, respectively (Kitchens et al 1976).

**Soil/Sediment:** Based on a classification scheme (Swann et al 1983), an estimated Koc value of 37 (SRC), determined from a log Kow of 0.35 (Hansch et al 1995) and a regression-derived equation (Lyman et al 1990), indicates that formaldehyde is expected to have very high mobility in soil (SRC). Volatilization of formaldehyde from moist soil surfaces is not expected to be an important fate process (SRC) given a Henry's Law

constant of  $3.4 \times 10^{-7}$  atm-cu m/mole (Betterton and Hoffmann 1988). Volatilization of formaldehyde from dry soil surfaces will occur because it is a gas (Boublik et al 1984). Formaldehyde readily biodegrades under both aqueous aerobic and anaerobic conditions in the environment which suggests that these processes may be important in soil (SRC). Formaldehyde in aqueous effluent was degraded by activated sludge and sewage in 48-72 hr. In a die-away test using water from a stagnant lake, degradation was complete in 30 and 40 hrs under aerobic and anaerobic conditions, respectively (Kitchens et al 1976).

**Biota:** An estimated BCF of 3 was calculated for formaldehyde (SRC), using a log Kow of 0.35 (Hansch et al 1995) and a regression-derived equation (Meylan et al 1999). According to a classification scheme (Franke et al 1994), this BCF suggests the potential for bioconcentration in aquatic organisms is low (SRC). Experiments performed on a variety of fish and shrimp show no bioconcentration of formaldehyde (Hose and Lightner 1980, Sills and, Allen 1979).

## Regulations

MCL (Drinking Water): 0.7 ug/L

Universal Treatment Standards (UTS): none

Groundwater Water Quality Criterion : 100 ug/L

PQL: 5.0 µg/g

IDLH: 20 ppm

EPA RfD: 0.2 mg/kg-day

EPA RfC: 0.0098 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
9200	100000	0.15	5500

## Remediation

Use fluorocarbon water spray, Cellosize and Hycar to diminish vapors. Sodium carbonate, ammonium hydroxide, or sodium sulfite can neutralize the spill (Prager 1995).

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**Hydrazine                      302-01-2****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** Hydrazine and dimethylhydrazines degrade rapidly in air through reactions with ozone, hydroxyl (OH) radicals, and nitrogen dioxide (WHO 1987). The reaction of hydrazine and 1,1-dimethylhydrazine with ozone is probably the major fate of these chemicals in the atmosphere. The reaction rate constant for hydrazine, derived from its decay rate in the presence of excess ozone, was about  $3 \times 10^{-17}$  cm<sup>3</sup> molecule-s and for 1,1-dimethylhydrazine the rate was greater than  $1 \times 10^{-15}$  cm<sup>3</sup> molecule-s (Atkinson and Carter 1984). Major reaction products were hydrogen peroxide for the hydrazine reaction and dimethylnitrosamine (about 60%) for the 1,1-dimethylhydrazine reaction. Estimated atmospheric half-lives ranged from less than 10 minutes for hydrazine during an ozone pollution episode to less than 2 hours under usual conditions, with a half-life about one-tenth that time for 1,1-dimethylhydrazine (Tuazon et al. 1981). Reported results of additional studies indicate a reaction rate constant for hydrazine of  $2.5 \times 10^{-16}$  cm<sup>3</sup> molecule-s, resulting in an estimated half-life of less than 1 min (Stone 1989). The reported measured rate constant for reaction of hydrazine with atmospheric hydroxyl (OH) radicals producing ammonia and nitrogen gas was  $6.1 \times 10^{-11}$  cm<sup>3</sup> molecule-s (Harris et al. 1979). The rate constant for 1,1-dimethylhydrazine was not measured since the chemical decomposed rapidly in the test system, but the value was estimated at  $5 \times 10^{-11}$  cm<sup>3</sup> molecule-s. Assuming an average OH radical concentration of about 106 molecule/cm<sup>3</sup>, the tropospheric half-lives of both chemicals due to reaction with OH were estimated to be about 3 hours. The half-lives are expected to range from less than 1 hour in polluted urban air to 3-6 hours in less polluted atmospheres (Tuazon et al. 1981). Hydrazine and 1,1-dimethylhydrazine react rapidly with nitrogen oxides in both the light and dark, with a half-life of about 2 hours for hydrazine and less than 10 minutes for 1,1-dimethylhydrazine (Pitts et al. 1980). Hydrazine and 1,1-dimethylhydrazine may also be removed from the atmosphere by autooxidation. In a dark reaction chamber, the approximate half-lives of hydrazine ranged from 1.8 to 5 hours, with the lower value measured at higher humidity. Reported values for 1,1-dimethylhydrazine under similar conditions were 5.9-9 hours. Surface interactions are important in controlling the rates of these reactions (Stone 1989). Although data were not located for 1,2-dimethylhydrazine, this chemical is expected to be degraded in the atmosphere by undergoing the same reactions as hydrazine and 1,1-dimethylhydrazine, although the rate and extent of degradation may be different.

**Water:** Hydrazine and 1,1-dimethylhydrazine degrade in aqueous systems, but the rate of degradation is dependent on specific aquatic environmental factors, including pH,

hardness, temperature, oxygen concentration, and the presence of organic matter and metal ions (Moliner and Street 1989 and WHO 1987). Oxidation and biodegradation are the primary removal mechanisms. Reaction of hydrazine with dissolved oxygen is catalyzed by metal ions, particularly copper (EPA 1984). The reaction rate is strongly influenced by pH; degradation proceeds more rapidly in alkaline solutions. Hydrazine is rapidly removed from polluted waters, with less than one-third of the original concentration remaining in dirty river water after 2 hours (Slonim and Gisclard 1976). More than 90% of the hydrazine added to pond or chlorinated, filtered county water disappeared after 1 day. However, chlorinated, filtered, and softened city water contained almost the original amount of hydrazine after 4 days. Organic matter in the water and hardness were reported to be the major factors in the differing rates of degradation.

The primary reaction pathway for hydrazine degradation in water produces nitrogen gas and water (Moliner and Street 1989). In oxygen-deficient waters or in the presence of metal ions which serve as catalysts, ammonia may also be produced. The reaction of 1,1-dimethylhydrazine with dissolved oxygen in water may proceed by a process catalyzed by copper ions or by an uncatalyzed reaction (Banerjee et al. 1984). The products include dimethylnitrosamine, formaldehyde, dimethylamine, and other related chemicals. Dimethylnitrosamine did not form in dilute solutions, which might be encountered in ambient waters, but was reported in concentrated solutions, which could be present in the vicinity of spills (EPA 1984). The reported half-life of 1,1-dimethylhydrazine in ponds and seawaters ranged from 10 to 14 days, presumably because of reaction with oxygen and other free radicals (EPA 1984). Biodegradation may be a significant removal process at low hydrazine concentrations in ambient waters, but at higher concentrations the chemical is toxic to microorganisms. In the presence of bacterial cells, more than 90% of the hydrazine was degraded in six water samples containing 11 µg/mL of the chemical within 2 hours (Ou and Street 1987). Lower degradation rates were reported with increasing hydrazine concentrations. No degradation was reported for incubation of these waters without bacteria. Additional studies indicate that hydrazine and 1,1-dimethylhydrazine are toxic to bacterial populations. Concentrations of hydrazine and 1,1-dimethylhydrazine that reduced bacterial metabolism by 50% ranged from 14.6 to 145 mg/L and from 19.2 to 9,060 mg/L, respectively (Kane and Williamson 1983). Thus, biological treatment would not be useful for spills of these chemicals into the aquatic environment.

**Sediment and Soil:** Hydrazine appears to degrade more rapidly in soil than in water, with oxidation and biodegradation as the main removal processes. Hydrazine applied to nonsterile Arredondo soil (fine sand) at concentrations of 10, 100, and 500 µg/g was completely degraded in 1.5 hours, 1 day, and 8 days, respectively (Ou and Street 1987). In this study, comparison to degradation rates in sterile soils indicated that autoxidation

appeared to be the major factor contributing to disappearance of the chemical, but the study authors attributed about 20% of removal to biodegradation. Several heterotrophic soil bacteria were reported to degrade hydrazine, indicating that microbial degradation may contribute to removal of the chemical from soil (Ou 1987).

**Biota:** In a study of hydrazine in aqueous systems, the chemical was reported to be absorbed by guppies from a 0.5 µg/L solution (Slonim and Gisclard 1976). After 96 hours, the hydrazine concentration in fish was 144 µg/g, indicating a moderate tendency to bioconcentrate. However, the bioconcentration of hydrazine and dimethylhydrazines is not expected to be important in aquatic systems because of the rapid degradation of these chemicals in water as well as their low octanol-water partition coefficients.

**Environmental partitioning:** Hydrazine or dimethylhydrazines released to water or soil may volatilize into air or sorb onto soil. These chemicals have low vapor pressures and are miscible in water. Therefore, volatilization is not expected to be an important removal process. Reported evaporation rates from aqueous solutions under laboratory conditions were 0.49 mg/cm<sup>2</sup> minute for hydrazine and 13 mg/cm<sup>2</sup> minute for 1,1-dimethylhydrazine (EPA 1984a). The significance of these values to environmental conditions is unknown. Data from other studies indicate that volatilization of these chemicals from water increases with higher concentrations of the chemical and in the presence of sunlight (due to increased temperature of the hydrazine pool). Based on air dispersion modeling, volatilization of hydrazine from surface soil following a spill is expected to be sufficient (16-100 mg/cm<sup>2</sup> hour) to generate a short-term ambient air concentration of 4 mg/m<sup>3</sup> up to 2 km downwind of the spill under worst-case meteorological conditions (MacNaughton et al. 1981). Degradation of hydrazine would likely reduce the concentration within several hours. Atmospheric transport of hydrazine or dimethylhydrazines may occur, but transport will be limited by the high reactivity of the chemicals in the atmosphere. No data were located on deposition of hydrazine or dimethylhydrazines from air to water or soil, but deposition would also be limited by their high reactivity. Hydrazine undergoes complex interactions with soils, including both reversible physical-sorption and irreversible chemisorption to colloids (Mansell et al. 1988). In a study on the adsorption and leaching characteristics of hydrazine fuels, no adsorption of 1,1-dimethylhydrazine was observed on sand, with almost 100% of the chemical leaching with water (Braun and Zirrolli 1983). In three other soils, adsorption ranged from 26% to 80%. No correlation between adsorption and soil organic content or pH was observed. The mechanisms of attenuation in soil materials were not reported. However, reported results of additional hydrazine adsorption studies with clays and soils indicate that adsorption may be correlated with soil organic matter and clay content and is highly dependent on pH; hydrazine appears to be adsorbed by different mechanisms under acidic and alkaline conditions (Moliner and Street 1989).

## Regulatory

UTS: None established

IDLH: 50 ppm

EPA Rf D: None established

EPA RfC: None established

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
0.16	0.82	0.00039	0.022

## Remediation

Hydrazine has been disposed of by dilution with water to form at least a 400 g/liter solution, followed by neutralization with dilute sulfuric acid and drainage into a sewer with abundant water (IRPTC, 1985). However, it should be noted that even very dilute solutions of 0.1 mg/liter can be toxic for aquatic life. Alternatively, hydrazine has been burnt in an open pit after the addition of a hydrocarbon solvent (IRPTC, 1985). A better procedure is to dilute with abundant water and then oxidize the diluted solution (to below 20 g/liter) with hydrogen peroxide, calcium hypochlorite, or sodium hypochlorite before draining into a sewer (NEPSS, 1975).

Hydrazine vapor emissions can be controlled by scrubbing, using water as the scrubbing liquid, or by the direct flame of catalytic incineration (Gordon & Lewandowski, 1980). Hydrazine sulfate, a commonly-used derivative, may be disposed of by incineration (IRPTC, 1985). Hydrazine degradation is very rapid in the environment.

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## Hydrogen Chloride (Hydrochloric Acid) 7647-01-0

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Hydrogen chloride is removed from air by wet deposition as chloride salts with an atmospheric lifetime of 1-5 days (Kao 1994).

**Water:** If released to water, hydrogen chloride dissociates readily in water to chloride and hydronium ions, decreasing the pH of the water (ATSDR 2002). A Henry's law constant of  $2.04 \times 10^6$  mol/L atm ( $4.90 \times 10^{-10}$  cu m atm/mol) has been reported for hydrochloric acid (Brimblecombe and Clegg 1989). This Henry's Law constant indicates that hydrochloric acid is expected to be essentially nonvolatile from water surfaces (Lyman et al 1990).

**Soil/Sediment:** If released to soil, hydrogen chloride will evaporate from dry soil surfaces and dissociate into chloride and hydronium ions in moist soil (ATSDR 2002).

### Regulations

MCL (Drinking Water): None (completely disassociates)

Groundwater Water Quality Criterion: 0.2 ug/L

IDLH: 50 ppm

EPA RfD: 0.0057 mg/kg-day

EPA RfC: 0.02 mg/m<sup>3</sup>

PRG: residential soil mg/kg    industrial soil mg/kg    ambient air ug/m<sup>3</sup>    tap water ug/L

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### Remediation

**Land spill:** Dig a pit, pond, lagoon, holding area to contain liquid or solid material. If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed polyurethane or foamed concrete. Absorb bulk liquid with fly ash or cement powder. Neutralize with agricultural lime (CaO), crushed limestone (CaCO<sub>3</sub>) or sodium

bicarbonate ( $\text{NaHCO}_3$ ). Water spill: Neutralize with agricultural lime ( $\text{CaO}$ ), crushed limestone ( $\text{CaCO}_3$ ) or sodium bicarbonate ( $\text{NaHCO}_3$ ).

Air release: Apply water spray or mist to knock down vapors. Vapor knockdown water is corrosive or toxic and should be diked for containment (Association of American Railroads).

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## Hydrogen Cyanide 74-90-8

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Most cyanide in the atmosphere exists almost entirely as hydrogen cyanide gas, although small amounts of metal cyanides may be present as particulate matter in the air (EPA 1984). Hydrogen cyanide is very resistant to photolysis at wavelengths of normal sunlight (EPA 1979). The most important reaction of hydrogen cyanide in air is the reaction with photochemically-generated hydroxyl radicals and subsequent rapid oxidation to carbon monoxide (CO) and nitric oxide (NO); photolysis and reaction with ozone are not important transformation processes, and reaction with singlet oxygen ( $O^1D$ ) is not a significant transformation process except at stratospheric altitudes where singlet oxygen is present in significant concentrations (Cicerone and Zellner 1983). The rate of hydroxyl radical reaction with hydrogen cyanide in the atmosphere depends on the altitude, and the rate of the reaction is at least an order of magnitude faster at lower tropospheric altitudes (0–8 km) than at upper tropospheric altitudes (10–12 km) (Cicerone and Zellner 1983). Based on a reaction rate constant of  $3 \times 10^{-14} \text{ cm}^3/(\text{molecule} \cdot \text{sec})$  at  $25^\circ\text{C}$  (Fritz et al. 1982) and assuming an average hydroxyl radical concentration of  $5 \times 10^5 \text{ molecules/cm}^3$ , the residence time for the reaction of hydrogen cyanide vapor with hydroxyl radicals in the atmosphere is  $\approx 2$  years. This value compares well with the atmospheric residence time derived by Cicerone and Zellner (1983) of approximately 2.5 years, with a range of 1.3–5.0 years, depending on the hydroxyl radical concentrations assumed. Using the equation  $t_{1/2} = 0.693\tau$  for converting residence time ( $\tau$ ) to half-life ( $t_{1/2}$ ) (Lyman 1982) and an estimated atmospheric residence time for hydrogen cyanide of 2–3 years, and assuming first-order kinetics for the reaction of hydrogen cyanide with hydroxyl radicals, an atmospheric half-life of 1.4–2.9 years can be calculated for hydrogen cyanide.

Cyanogen is reactive and does not persist in the environment unchanged (EPA 1978). Cyanogen reacts slowly with water to yield hydrogen cyanide and cyanic acid (HOCN) among other products (EPA 1979) and this hydrolysis reaction may be a possible atmospheric degradation pathway. Cyanogen has also been shown to react with hydroxyl radicals in the gas phase (Atkinson 1989). Based on a rate constant of  $2.5 \times 10^{-15} \text{ cm}^3/(\text{molecule} \cdot \text{sec})$  at  $27^\circ\text{C}$  and assuming an average hydroxyl radical concentration of  $5 \times 10^5 \text{ molecules/cm}^3$ , the residence time for the reaction of hydrogen cyanide vapor with hydroxyl radicals in the atmosphere is  $\approx 25$  years. Therefore, the reaction of cyanogen with photochemically-induced hydroxyl radicals will not play a significant role in the degradation of this compound in air.

No specific information was found in the available literature on the transformation and degradation of cyanogen chloride or thiocyanates in air. However, cyanogen chloride has been shown to undergo slow hydrolysis in neutral aqueous solution (rate constant at pH 7 of  $6.45 \times 10^{-5} \text{ mol}^{-1} \text{sec}^{-1}$ ) (U.S. Army 1989). Therefore, hydrolysis of this compound may be a possible atmospheric degradation pathway in air.

**Water:** Cyanide occurs most commonly as hydrogen cyanide in water, although it can also occur as the cyanide ion, alkali and alkaline earth metal cyanides (potassium cyanide, sodium cyanide, calcium cyanide), relatively stable metalocyanide complexes (ferricyanide complex  $[\text{Fe}(\text{CN})_6]^{3-}$ ), moderately stable metalocyanide complexes (complex nickel and copper cyanide), or easily decomposable metalocyanide complexes (zinc cyanide  $[\text{Zn}(\text{CN})_2]$ , cadmium cyanide  $[\text{Cd}(\text{CN})_2]$ ). The environmental fate of these cyanide compounds varies widely (EPA 1979).

In water, hydrogen cyanide and cyanide ion exist in equilibrium with their relative concentrations primarily dependent on pH and temperature. At pH  $< 8$ ,  $> 93\%$  of the free cyanide in water will exist as undissociated hydrogen cyanide (EPA 1978). Hydrogen cyanide can be hydrolyzed to formamide, which is subsequently hydrolyzed to ammonium and formate ions (EPA 1979).

However, the relatively slow rates of hydrolysis reported for hydrogen cyanide in acidic solution (Kriebel and McNally 1929; Kriebel and Peiker 1933) and of cyanides under alkaline conditions (Wiegand and Tremelling 1972) indicate that hydrolysis is not competitive with volatilization and biodegradation for removal of free cyanide from ambient waters (EPA 1979).

The significance of photolysis in the fate of cyanides in water has not been fully investigated. Hydrogen cyanide and cyanide ions in aqueous solution have been found to be very resistant to photolysis by natural sunlight, except under heterogeneous photocatalytic conditions (EPA 1979; Frank and Bard 1977). Photocatalytic oxidation may not be significant in natural waters, however, because of significant light reduction at increasingly greater depths (EPA 1992). In clear water or at water surfaces, some metalocyanides, such as ferrocyanides and ferricyanides, may decompose to the cyanide ion by photodissociation and subsequently form hydrogen cyanide. For example, diurnal changes in free cyanide concentrations in the drainage from spent precious metal ore heaps were found to maximize around mid-day due to the photodissociation of iron and cobalt cyanocomplexes (Johnson et al. 2002). Because of adsorption of ferrocyanide onto soil surfaces and sediment of surface waters, and light scattering in turbid waters in the field, the rate of free cyanide formation from the photolysis of ferrocyanide in runoff and surface water from washout of ferrocyanide in de-icing salt will be slower than from laboratory photolysis with clean water (EPA 1979).

**Sediment and Soil:** Analogous to the fate of cyanides in water, it is predicted that the fate of cyanides in soil would be dependent on cyanide concentrations, pH, temperature, metal content, concentration of microbes, availability of nutrients, and acclimation of microbes. Cyanide may occur as hydrogen cyanide, alkali metal salts, or as immobile metalocyanide complexes. In soil, cyanide present at low concentrations would biodegrade under aerobic conditions with the initial formation of ammonia, which would be converted to nitrite and nitrate in the presence of nitrifying bacteria. Under anaerobic conditions, the cyanides ion will denitrify to gaseous nitrogen (Richards and Shieh 1989). Upper limits of 200 and 2 ppm (mg/kg  $\text{CN}^-$ ) have been reported for uninhibited aerobic and anaerobic biodegradation of cyanide in soil respectively (Fueller 1985); however, these limits have not been confirmed in other studies (Thomas and Lester 1993). Cyanide ions in soil are not involved in oxidation-reduction reactions but may undergo complexation reactions with metal ions in soil (EPA 1978).

**Biota:** There are no data available to indicate that simple metal cyanides and hydrogen cyanide bioconcentrate in aquatic organisms (EPA 1979, 1985, 1992). Bioconcentration factors (BCFs) of 0.73 and 1.62 can be calculated for hydrogen cyanide, using the equation of Veith et al. (1979) for the BCF of a chemical in whole fish ( $\log \text{BCF}$ , 0.85;  $\log \text{Kow}$ , -0.70) and the  $\log \text{Kow}$  values. Similarly, the calculated BCF for sodium cyanide is 0.47. There is some evidence that certain metal cyanide complexes bioaccumulate in aquatic organisms. Fish from water with soluble silver and copper cyanide complexes were found to have metal cyanides in their tissues at concentrations ranging up to 168 and 304  $\mu\text{g/g}$ , respectively (wet or dry weight not specified) (EPA 1979). It is difficult to evaluate the toxicologic significance of bioaccumulation of metal cyanide complexes because these compounds are much less toxic than soluble hydrogen cyanide, sodium cyanide, or potassium cyanide (EPA 1992). There is no evidence of biomagnification of cyanides in the food chain (EPA 1978). Accumulation of cyanide in food webs is not expected, considering the rapid detoxification of cyanide by most species and the lethal effects of large doses of cyanide (EPA 1978).

**Environmental Partitioning:** Because hydrogen cyanide is a gas and has a relatively slow degradation rate in air, the atmosphere will be the ultimate sink for this compound. Almost all of

the hydrogen cyanide released to the atmosphere remains in the lower altitudes (troposphere); only 2% of tropospheric hydrogen cyanide is transferred to the stratosphere (Cicerone and Zellner 1983). Cyanide has the potential to be transported over long distances from its emission source. Despite higher water solubility at saturated pressure, the removal of hydrogen cyanide by rainwater appears to be a negligible partitioning pathway (Cicerone and Zellner 1983). Because hydrogen cyanide is a gas, its removal from air by dry deposition is also likely to be negligible. However, metal cyanide particles, particularly water-soluble cyanide particles, are expected to be removed from the air by both wet and dry deposition.

Volatilization of hydrogen cyanide would be a significant loss mechanism for cyanides from soil surfaces at a pH <9.2. Cyanides are fairly mobile in soil. Mobility is lowest in soils with low pH and high concentrations of free iron oxides, positively charged particles, and clays (e.g., chlorite, kaolin, gibbsite), and highest in soils with high pH, high concentrations of free CaCO<sub>3</sub> and negatively charged particles, and low clay content (EPA 1979). Although cyanide has a low soil sorption capability, it is usually not detected in groundwater, probably because of fixation by trace metals through complexation or transformation by soil microorganisms (EPA 1978). In soils, where cyanide levels are high enough to be toxic to microorganisms (i.e., landfills, spills), this compound may leach into groundwater (EPA 1984). Also, leaching of cyanide into a shallow aquifer can occur, as demonstrated by the high concentration of cyanide (1,200 µg/L) in groundwater sampled from the Biscayne Aquifer in Dade County, Florida, which lies below a solid waste site (Myers 1983).

## Regulatory

Drinking Water Standards: MCLG: 0.2 mg/L (cyanide)

RfD: 0.02 mg/kg/day

RfC: 0.003 mg/m<sup>3</sup>

UTS: None established

Soil PQL: 3 ug/L

IDLH: 25 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
11	35	3.1	6.2

## Remediation

Cyanides are not persistent in water or soil. Cyanides may accumulate in bottom sediments, but residues are generally as low as <1 mg/kg even near polluting sources. Majority of an accidental release of cyanide is volatilized to the atmosphere where it is quickly diluted and degraded by ultra violet. Other factors, such as biological oxidation, precipitation and the effects of sunlight also contribute to cyanide degradation. There is

no evidence of bioaccumulation in the food chain, and hence, secondary poisoning does not occur.

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## Hydrogen Fluoride (HF) 7664-39-3

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Hydrogen fluoride is the most abundant gaseous fluoride released into the atmosphere. It reacts with many materials both in vapor and in aerosols. For example, hydrogen fluoride reacts with silica, forming silicon tetrafluoride. However, no information was found on the reactions of hydrogen fluoride with common atmospheric species or estimates of its overall atmospheric half-life. The predominant mode of degradation of inorganic fluorides in the air is hydrolysis. Silicon tetrafluoride, a major industrial pollutant, reacts with water vapor in air to form hydrated silica and fluorosilicic acid. Sulfur hexafluoride, a gaseous dielectric for electrical and electronic equipment, reacts with water at elevated temperatures ( $>850\text{ }^{\circ}\text{C}$ ) to form sulfuric acid and hydrogen fluoride (Guo et al. 2001). Molecular fluorine hydrolyzes to form hydrogen fluoride and oxygen. Hydrolysis of uranium hexafluoride, which is used in nuclear power applications, also produces hydrogen fluoride as well as nonvolatile uranyl fluoride. These compounds are then removed from the atmosphere by condensation or nucleation processes (NAS 1971). Fluorides emitted by industries in particulate matter are stable compounds that do not readily hydrolyze.

**Water:** Contrary to traditional thought, hydrogen fluoride, a very weak acid in dilute solution, is dissociated in solution, but forms tight ion pairs  $\text{F}\cdots\text{H}^+-\text{OH}_2$ , unique to  $\text{F}^-$ , which reduce the thermodynamic activity coefficient of  $\text{H}_3\text{O}^+$  (Cotton et al. 1999). In natural water, fluoride ions form strong complexes with aluminum, and fluorine chemistry in water is largely regulated by aluminum concentration and pH. Below pH 5, fluorine is almost entirely complexed with aluminum and consequently, the concentration of free  $\text{F}^-$  is low. As the pH increases, Al-OH complexes dominate over Al-F complexes and the free level increases. The dominant Al-F complex at pH  $<5$  is  $\text{AlF}_2^+$  (Skjelkvale 1994). In the absence of aluminum, dissolved fluorides are usually present as free  $\text{F}^-$  at neutral pH (Bell et al. 1970). As the pH decreases, the proportion of  $\text{F}^-$  decreases, while  $\text{HF}_2^-$  and undissociated hydrogen fluoride increase. Levels of undissociated hydrogen fluoride also increase in concentrated solutions. Fluorine can form stable complexes with calcium and magnesium, which are present in sea water. Using the stability constants valid for sea water, 51.0% of fluorine will be present as free  $\text{F}^-$ , 47.0% as  $\text{MgF}^+$ , and 2.0% as  $\text{CaF}^+$  (Stumm and Morgan 1981). Calcium carbonate precipitation dominates the removal of dissolved fluoride from sea water. Fluoride is incorporated into the calcium salt structure and is removed from solution when the latter precipitates.

**Sediment and Soil:** Fluoride occurs in soil as a variety of minerals and complexes with aluminum, iron, and calcium. At low pH, aluminum complexes,  $\text{AlF}_3$ ,  $\text{AlF}_2^+$ , and  $\text{AlF}_2^+$ , are the dominant dissolved species, and the availability of these soluble complexes increases with decreasing pH (Fung et al. 1999; Shacklette et al. 1974). This explains why more acidic soils have both higher water-soluble fluoride and higher extractable aluminum levels. In certain soils in which calcium is present mostly as calcium fluoride and in which there is sufficient alumina, fluoride is fixed by the formation of relatively insoluble aluminum fluorosilicate,  $\text{Al}_2(\text{SiF}_6)_3$  (Brewer 1966).

**Transport and Partitioning:** In the atmosphere, gaseous hydrogen fluoride is absorbed by atmospheric water (rain, clouds, fog, snow) forming an aerosol or fog of aqueous hydrofluoric acid. It will be removed from the atmosphere primarily by wet deposition (including rainout or in-cloud scavenging and washout or below-cloud scavenging). Particulate fluorides are similarly removed from the atmosphere and deposited on land or surface water by wet and dry deposition. Atmospheric precipitation weathers crustal rocks and soil, but dissolves out very little fluoride; most of the fluoride mobilized during weathering is bound to solids such as clays. Upon reaching bodies of water, fluorides gravitate to the sediment (Carpenter 1969). Most of the fluorides in the oceans are received from rivers; a lesser amount comes from atmospheric deposition. Losses occur in aerosols to the atmosphere and incorporation into the tissue of aquatic organisms. Fluorides have been shown to accumulate in some marine aquatic organisms.

## Regulatory

MCL: 4 mg/L (fluoride)

MCL (Drinking Water): None

Universal Treatment Standards (UTS): 0.29 mg/l<sup>2</sup> in wastewater standards

Controlling Water Quality Criterion for Surface Water:

Basis: Aquatic Toxicity Criterion Concentration: 21 ug/L Criterion Quantitation Limit: 5 ug/L

EPA RfD: 0.0005 mg/kg-day

EPA RfC: 0.000002 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
0.1	0.345	0.21	0.042

## Remediation

Reaction with excess lime followed by lagooning and either recovery or landfill disposal of the separated calcium fluoride. Alternatively, hydrogen can be recovered and recycled in many cases. In accordance with 40CFR165 recommendations for the disposal pesticides and pesticide containers. Must be disposed properly by following package label directions or by contacting your state pesticide or environmental control agency or by contacting your regional EPA office. Consult with environmental regulatory agencies for guidance on acceptable disposal practices. . Generators of waste containing this contaminant ( $\geq 100$  kg/mo) must conform to EPA regulations governing storage, transportation, treatment, and waste disposal (Sittig 2002).

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**Mercuric Chloride      7487-94-7****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** Mercury is released to the environment by both natural processes (e.g., volcanic activity and weathering of mercury-containing rocks) and anthropogenic sources. Anthropogenic releases are primarily to the atmosphere. The primary form of atmospheric mercury, metallic mercury vapor ( $\text{Hg}^0$ ), is oxidized by ozone to other forms (e.g.,  $\text{Hg}^{+2}$ ) and is removed from the atmosphere by precipitation (Brosset 1991). The oxidation/reduction of mercury with dissolved ozone, hydrogen peroxide, hypochlorite entities, or organoperoxy compounds or radicals may also occur in the atmosphere (Schroeder et al. 1991). The overall residence time of elemental mercury in the atmosphere has been estimated to be 6 days to 2 years, although in clouds, a fast oxidation reaction on the order of hours may occur between elemental mercury and ozone. Some mercury compounds, such as mercuric sulfide, are quite stable in the atmosphere as a result of their binding to particles in the aerosol phase (Lindqvist 1991). Other mercury compounds, such as mercuric hydroxide ( $\text{Hg}[\text{OH}]_2$ ), which may be found in the aqueous phase of the atmosphere (e.g., rain), are rapidly reduced to monovalent mercury in sunlight (Munthe and McElroy 1992). The main atmospheric transformation process for organomercurials appears to be photolysis (EPA 1984; Johnson and Bramen 1974).

**Water:** The most important transformation process in the environmental fate of mercury in surface waters is biotransformation. Photolysis of organomercurials may also occur in surface waters, but the significance of this process in relation to biotransformation is not clear (Callahan et al. 1979). Any form of mercury entering surface waters can be microbially converted to methylmercuric ions, given favorable conditions. Sulfur-reducing bacteria are responsible for most of the mercury methylation in the environment (Gilmour and Henry 1991), with anaerobic conditions favoring their activity (Regnell and Tunlid 1991). Yeasts, such as *Candida albicans* and *Saccharomyces cerevisiae*, whose growth is favored by low pH conditions, are able to methylate mercury and are also able to reduce ionic mercury to elemental mercury (Yannai et al. 1991). Methyl cobalamine compounds produced by bacterial synthesis appear to be involved in the nonenzymatic methylation of inorganic mercury ions (Regnell and Tunlid 1991). The rate of methylmercury formation by this process is largely determined by the concentration of methyl cobalamine compounds, inorganic mercuric ions, and the oxygen concentration of the water, with the rate increasing as the conditions become anaerobic. Volatile elemental mercury may be formed through the demethylation of methylmercury or the reduction of inorganic mercury, with anaerobic conditions again favoring the demethylation of the methylmercury (Barkay et al. 1989; Callahan et al. 1979; Regnell and Tunlid 1991). Increased dissolved organic carbon levels reduce methylation of mercury in

the water column (Gilmour and Henry 1991), possibly as a result of the binding of free mercury ions to the dissolved organic carbon at low pH, thus reducing their availability for methylation, or the dissolved organic carbon may inhibit the methylating bacteria (Miskimmin et al. 1992). Alternatively, low pH favors the methylation of mercury in the water column, particularly in acid deposition lakes, while inhibiting its demethylation (Gilmour and Henry 1991). It has also been shown that the methylation rate is not affected by addition of sulfate in softwater lakes (Kerry et al. 1991).

**Sediment and Soil:** Mercury compounds in soils may undergo the same chemical and biological transformations described for surface waters. Mercuric mercury usually forms various complexes with chloride and hydroxide ions in soils; the specific complexes formed depend on the pH, salt content, and composition of the soil solution. Formation and degradation of organic mercurials in soils appear to be mediated by the same types of microbial processes occurring in surface waters and may also occur through abiotic processes (Andersson 1979). Elevated levels of chloride ions reduce methylation of mercury in river sediments, sludge, and soil (Olson et al. 1991), although increased levels of organic carbon and sulfate ions increase methylation in sediments (Gilmour and Henry 1991). In freshwater and estuarine ecosystems, the presence of chloride ions (0.02 M) may accelerate the release of mercury from sediments (Wang et al. 1991).

**Biota:** Bioconcentration factors of 10,000 and 40,000 have been obtained for mercuric chloride and methylmercury with oyster (EPA 1986).

## Environmental partitioning

The natural global bio-geochemical cycling of mercury is characterized by degassing of the element from soils and surface waters, followed by atmospheric transport, deposition of mercury back to land and surface waters, and sorption of the compound to soil or sediment particulates. Mercury deposited on land and open water is in part revolatilized back into the atmosphere. This emission, deposition, and revolatilization create difficulties in tracing the movement of mercury to its sources (WHO 1990). Particulate-bound mercury can be converted to insoluble mercury sulfide and precipitated or bioconverted into more volatile or soluble forms that re-enter the atmosphere or are bioaccumulated in aquatic and terrestrial food chains (EPA 1984b).

Mercury is transformed in the environment by biotic and abiotic oxidation and reduction, bioconversion of inorganic and organic forms, and photolysis of organomercurials. Inorganic mercury can be methylated by microorganisms indigenous to soils, fresh water, and salt water. This process is mediated by various microbial populations under both aerobic and anaerobic conditions. The most probable mechanism for this reaction involves the nonenzymatic methylation of mercuric mercury ions by methylcobalamine compounds produced as a result of bacterial synthesis. Mercury forms stable complexes with organic compounds. Monoalkyl mercury

compounds (e.g., methylmercuric chloride) are relatively soluble; however, the solubility of methylmercury is decreased with increasing dissolved organic carbon content, indicating that it is bound by organic matter in water (Miskimmin 1991). Dialkyl mercury compounds (e.g., dimethylmercury) are relatively insoluble (Callahan et al. 1979; EPA 1984). Dimethylmercury is volatile, although it makes up less than 3% of the dissolved gaseous mercury found in water (Andersson et al. 1990; Vandal et al. 1991).

## Regulatory

MCL (Drinking Water): 0.002 mg/L. (Hg)

Universal Treatment Standards (UTS) for Hg others: 0.15mg/l<sup>2</sup> in wastewater standards

0.025mg/L TCLP in nonwastewater standards

EPA RfD: 0.0003 mg/k/day (U.S. EPA, 1995)

EPA RfC:           mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
6.1	88	0.31	3.6

Soil PQL (mercury): 0.1mg/kg

## Remediation

Mercury removal from waste water can be accomplished by these processes: BMS process; Chlorine is added to the waste water, oxidizing any mercury present to the ionic state. The BMS adsorbent (an activated carbon concentrate of sulfur compound on its surface) is used to collect ionic mercury. The spent adsorbent is then distilled to recover the mercury, leaving a carbon residue for reuse or disposal. TMR IMAC Process; Waste water is fed into a reactor, whereby a slight excess of chlorine is maintained, oxidizing any mercury present to ionic mercury. The liquid is then passed through the TMR IMAC ion-exchange resin where mercury ions are adsorbed. The mercury is then stripped from the spent resin with hydrochloric acid solution (Environmental Canada).

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**Methamidophos (O, S-Dimethyl phosphoramidothioate)**  
**10265-92-6**

**ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND  
TRANSFORMATION**

**Air:** Based on a reported vapor pressure for methamidophos of  $3 \times 10^{-4}$  mm Hg at 30 deg C (Bidleman 1988), methamidophos is expected to exist almost entirely in the vapor phase in the ambient atmosphere (Tomlin 1997). The rate constant for the reaction of methamidophos with photochemically produced hydroxyl radicals is estimated to be  $2.7 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C, which corresponds to a half-life of 4.6 hours for an atmospheric hydroxyl radical concentration of  $5 \times 10^5$  molecules/cu cm (Meylan and Howard 1993).

**Water:** A Koc value of 5 indicates that methamidophos is not expected to adsorb to suspended solids and sediment (Swann et al. 1983). Volatilization from water surfaces is not expected (Lyman et al. 1990) based upon an estimated Henry's Law constant of  $8.7 \times 10^{-10}$  atm-cu m/mole(SRC), developed using a fragment constant estimation method (Meylan and Howard 1991). According to a classification scheme (Franke et al 1994), an estimated BCF of 3.1, from its log Kow of -0.8 (Tomlin 1997) and a regression-derived equation (Meylan et al.1999), suggests the potential for bioconcentration in aquatic organisms is low. Half-lives for the hydrolysis of methamidophos at 22 deg C and pH 4, 7, and 9 have been reported to be 1.8 years, 120 hours, and 70 hours, respectively (Tomlin 1997). The photolysis rate in water is 0.0079 per day (USDA 2003), corresponding to a half-life of 87 days.

**Soil and Sediment:** A Koc value of 5 (Wauchope et al. 1991), indicates that methamidophos is expected to have very high mobility in soil (Swann et al. 1983). Volatilization of methamidophos from moist soil surfaces is not expected to be an important fate process(SRC) given an estimated Henry's Law constant of  $8.7 \times 10^{-10}$  atm-cu m/mole(SRC), using a fragment constant estimation method (Tomlin 1997). Methamidophos is not expected to volatilize from dry soil surfaces (SRC) based upon a vapor pressure of  $3.5 \times 10^{-5}$  mm Hg (Tomlin 1997). Average half-lives in soil under aerobic and anaerobic conditions are 4 and 7.5 days, respectively (USDA 2003). The photolysis rate in soil is reported as 0.210 per day (USDA 2003), corresponding to a half-life of 3.3 days (SRC).

**Biota:** An estimated BCF of 3.1 was calculated for methamidophos, using a log Kow of -0.8 (Tomlin 1997) and a regression-derived equation (Meylan et al.1999). This BCF

suggests the potential for bioconcentration in aquatic organisms is low (Franke et al. 1993).

**Environmental partitioning:** Methamidophos' production may result in its release to the environment through various waste streams; its use as an insecticide and acaricide will result in its direct release to the environment. If released to air, a vapor pressure of  $3.5 \times 10^{-5}$  mm Hg at 25 deg C indicates methamidophos will exist in both the vapor and particulate phases in the ambient atmosphere. Vapor-phase methamidophos will be degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals; the half-life for this reaction in air is estimated to be 12 hours. Particulate-phase methamidophos will be removed from the atmosphere by wet and dry deposition. If released to soil, methamidophos is expected to have very high mobility based upon a Koc of 5. Volatilization from moist soil surfaces is not expected to be an important fate process based upon an estimated Henry's Law constant of  $8.7 \times 10^{-10}$  atm-cu m/mole. Volatilization The log Pow of  $-0.8$  and the animal metabolism and feeding studies suggest that methamidophos should not be described as fat-soluble (Methamidophos 100).

## Regulatory

MCL (Drinking Water): None

Drinking Water Remedial Target: 1.8 ug/L

Universal Treatment Standards (UTS): none

EPA Rf D: 0.00005 mg/kg-day

EPA RfC: none

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
3.1	44	0.18	1.8

## Remediation

Spills and powders should be cleaned up using a dustless method (e.g., by a vacuum cleaner suitable for use with toxic dusts). Alternatively, mix with damp saw-dust and place in a separate container for subsequent disposal. Dry brushing should not be carried out, as this creates dust clouds (ICPS).

Large amounts should be incinerated at high temperature in a unit with effluent gas scrubbing. When no incinerator is available, bury in an approved dump, or in an area where there is no risk of contamination

of surface or groundwater. Before burying, liberally mix with sodium carbonate (washing soda) crystals, to help neutralize the product, and with soil rich in organic matter (ICPS).

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## Methyl Bromide (Bromomethane) 74-83-9

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), **methyl bromide**, which has a vapor pressure of 1,620 mm Hg at 25 deg C (Daubert and Danner 1989), will exist in the gas phase (SRC). Gas-phase **methyl bromide** is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals; the half-life for this reaction in air is about one year, calculated from its rate constant of  $4.0 \times 10^{-14}$  cu cm/molecule-sec at 25 deg C (Atkinson 1989). **Methyl bromide** is not expected to undergo direct photolysis due to the lack of absorption in the environmental UV spectrum ( $>290$  nm) (Robbins 1976). Upward diffusion of **methyl bromide** to the stratosphere is believed to be the dominant loss mechanism of this compound from the troposphere (Robbins 1976). In the stratosphere, into which lower wavelength light can penetrate, photolysis is expected to be the predominant removal mechanism of **methyl bromide** (Robbins 1976).

**Water:** Koc values ranging from 9 to 22 (USDA 2001), indicate that **methyl bromide** is not expected to adsorb to suspended solids and sediment (Swann et al 1983),. Volatilization from water surfaces is expected based upon a Henry's Law constant of  $7.34 \times 10^{-3}$  atm-cu m/mole (Yates and Gan 1998). Using this Henry's Law constant and an estimation method (Lyman et al 1990), volatilization half-lives for a model river and model lake are 1.0 hrs and 3.9 days, respectively. An estimated BCF of 2 (SRC), from its log Kow (Hansch et al 1995) and a regression-derived equation (Meylan et al 1999), suggests the potential for bioconcentration in aquatic organisms is low (Franke et al 1995). Chemical hydrolysis is the primary degradation mechanism for **methyl bromide** in water. The hydrolysis rate constant of **methyl bromide** at 25 deg C and pH 7 is  $4.1 \times 10^{-7}$  1/sec which translates into a half-life of 20 days (Mabey and Mill 1978). Another reported hydrolysis rate constant for **methyl bromide** is  $3 \times 10^{-7}$  1/sec at 25 deg C (Castro and Belser 1981), which translates into a half-life of 26.7 days. The products of **methyl bromide** hydrolysis are methanol and bromide ion (Castro and Belser 1981).

**Soil/Sediment:** Koc values ranging from 9 to 22 (USDA 2001) indicate that **methyl bromide** is expected to have very high mobility in soil (Swann et al 1983),. Volatilization of methyl bromide from moist soil surfaces is an important fate process given a Henry's Law constant of  $7.34 \times 10^{-3}$  atm-cu m/mole (Yates and Gan 1996). The volatilization of methyl bromide from dry soil surfaces is an important removal mechanism from soil (SRC) based upon a vapor pressure of 1,620 mm Hg at 25 deg C (Daubert and Danner 1989). In soil, chemical reactions, likely nucleophilic

substitutions on soil organic matter, were identified as the predominate pathway through which methyl bromide was degraded (Gan and Yates 1996). For example, methyl bromide reacts rapidly with aniline (a model used to simulate reactions with organic matter in soil) compared to direct hydrolysis with water (Gan and Yates 1996); the degradation half-lives of methyl bromide with water and aniline are 20 days and 2.9 days at 24 deg C, respectively (Gan and Yates 1996). The oxidation of  $^{14}\text{C}$ -methyl bromide to  $^{14}\text{CO}_2$  was measured in field experiments with soils collected from two strawberry plots fumigated with mixtures of methyl bromide and chloropicrin (Miller et al 1997). Degradation of methyl bromide by chemical and/or biological processes accounted for 20 to 50% of the loss of methyl bromide during fumigation ( Miller et al 1997 ) with the remainder of methyl bromide loss due to volatilization.

**Environmental Partitioning:** Transformation or degradation of MeBr is an irreversible process that depletes MeBr from the soil-water-air system before it reaches the soil surface and volatilizes into the air. Extremely rapid transformation may deplete MeBr concentrations so quickly that efficacy is compromised. The actual transformation of MeBr in an agricultural soil is the sum of its hydrolysis in water, reactions with soil constituents, and decomposition by soil microorganisms.

## Regulations

MCL (Drinking Water): None

GPS - Groundwater protection standard (DWEL): 0.1 mg/L

Universal Treatment Standards (UTS): 0.1 mg/l<sub>2</sub> in wastewater standards and 15 mg/kg<sup>3</sup>

Groundwater Water Quality Criterion : 10 ug/L

PQL: 2

EPA Rf D: 0.0014 mg/kg-day

EPA RfC: 0.005 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
3.9	13	5.2	8.7

## Remediation

Generators of waste (equal to or greater than 100 kg/mo) containing this contaminant, EPA hazardous waste number U029, must conform with USEPA regulations in storage, transportation, treatment and disposal of waste (40 CFR).

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**Methyl Hydrazine      60-34-4****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), methylhydrazine, which has a measured vapor pressure of 50 mm Hg at 25 deg C (Boublik et al. 1984), will exist solely as a vapor in the ambient atmosphere. Vapor-phase methylhydrazine is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals; the half-life for this reaction in air is estimated to be about 6 hours calculated from its rate constant of  $6.50 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C (Atkinson 1989). Assuming an ozone concentration of  $7 \times 10^{11}$  molecules/cu cm, a maximum half-life of about 3.3 min for the reaction of methylhydrazine with ozone was estimated from a minimum rate constant of  $5 \times 10^{-15}$  cu cm/molecule sec (Tuazon et al. 1982). The half-life for the reaction between ozone and methylhydrazine was estimated to be <1 min during ozone pollution episodes and <12 min in the 'natural' troposphere (Tuazon et al. 1981). Reaction of methylhydrazine with ozone is expected to be the predominant fate of methylhydrazine in the atmosphere.

**Water:** An estimated Koc value of 18 (SRC), determined from a structure estimation method, indicates that methylhydrazine is not expected to adsorb to suspended solids and sediment (Swann et al. 1983), a. Methylhydrazine is a weak base with a pKa of 7.87 (Perrin 1965). This pKa value indicates that this compound will partially exist as a cation in water, and cations may adsorb to suspended solids and sediment more than neutral compounds (SRC). Volatilization of the protonated species will not occur since cations do not volatilize. Volatilization of the free base may occur (Lyman et al. 1990), based on an estimated Henry's Law constant of  $3 \times 10^{-6}$  atm-cu m/mole (SRC), calculated from its vapor pressure, 50 mm Hg (Boublik et al. 1984), and water solubility,  $1 \times 10^6$  mg/L (O'Neil 2001).

**Sediment:** An estimated Koc value of 18, determined from a structure estimation method (Meylan et al. 1992, Hansch et al. 1995), indicates that methylhydrazine is not expected to adsorb to suspended solids and sediment (Swann et al. 1983). Methylhydrazine is a weak base with a pKa of 7.87 (Perrin 1965). This pKa value indicates that this compound will partially exist as a cation in water, and cations may adsorb to suspended solids and sediment more than neutral compounds. Volatilization of the protonated species will not occur since cations do not volatilize. Volatilization of the free base may occur (Lyman et al. 1990), based on an estimated Henry's Law constant of  $3 \times 10^{-6}$  atm-cu m/mole, calculated from its vapor pressure, 50 mm Hg (Boublik et al. 1984), and water solubility,  $1 \times 10^6$  mg/L. Based on this Henry's Law constant, the volatilization

half-life from a model river (1 m deep, flowing 1 m/sec, wind velocity of 3 m/sec) (Perrin 1965) is estimated as 5 days. The volatilization half-life from a model lake (1 m deep, flowing 0.05 m/sec, wind velocity of 0.5 m/sec) (Lyman et al. 1990) is estimated as 62 days. According to a classification scheme (Franke et al. 1994), an estimated BCF of 3, from a log Kow of -1.05 (Hansch et al. 1995) and a regression-derived equation (Meylan et al. 1999), suggests the potential for bioconcentration in aquatic organisms is low. The estimated half-lives of methylhydrazine present at 9.5 mM in pond and sea water are 18.0 and 24.1 days, respectively, and at 19.0 mM are 13.1 days in both pond and sea water (Braun and Zirrolli 1983).

**Soil:** Based on a classification scheme (Swann et al. 1983), an estimated Koc value of 18, determined from a structure estimation method (Meylan et al. 1992), indicates that methylhydrazine is expected to have very high mobility in soil. Methylhydrazine is a weak base with a pKa of 7.87 (Perrin 1965). This pKa value indicates that this compound will partially exist as a cation in moist soils, and cations may adsorb to soil more than neutral compounds. Volatilization of the protonated species from moist soils will not occur since cations do not volatilize. Volatilization of the free base may occur, based on an estimated Henry's Law constant of  $3 \times 10^{-6}$  atm-cu m/mole, calculated from its vapor pressure, 50 mm Hg (Boublik et al. 1984), and water solubility,  $1 \times 10^6$  mg/L (O'Neil 2001). Volatilization from dry soils may occur based on the vapor pressure of this compound (Boublik et al. 1984). Approximately 0%, 5%, 20%, and 8% degradation of methylhydrazine was observed in a 1 hour soil column study using a cleaned sand (100% sand), sandy soil (99.1% sand, 0.4% clay, pH 6.1), organic soil (96.1% sand, 1% clay, 1% carbon, pH 6.4) and clay (69.3% sand, 27.95% clay, pH 3.7) (Braun and Zirrolli 1983).

**Biota:** An estimated BCF value of 3 was calculated for methylhydrazine, using a measured log Kow of -1.05 (Hansch et al. 1995) and a regression-derived equation (Meylan et al. 1995). According to a classification scheme (Franke et al. 1994), this BCF value suggests that bioconcentration in aquatic organisms is low.

**Environmental partitioning:** Methylhydrazine's production and use as a rocket fuel, solvent and chemical intermediate may result in its release to the environment through various waste streams. If released to the atmosphere, methylhydrazine will exist solely in the vapor phase in the ambient atmosphere, based on a measured vapor pressure of 50 mm Hg at 25 deg C. Vapor-phase methylhydrazine is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals and ozone with estimated half-lives of about 6 hours and 1-12 minutes, respectively. If released to soil, an estimated Koc value of 18 suggests that methylhydrazine will have very high mobility in soil. Methylhydrazine is a weak base having a pKa of 7.87, suggesting that it will partially exist in the protonated form in water and moist soils and the protonated form may adsorb to

soils more than the free base. Volatilization from moist soil surfaces will not occur for the protonated species since cations do not volatilize. Volatilization may occur for the neutral species based on an estimated Henry's Law constant of  $3 \times 10^{-6}$  atm-cu m/mole (HSDB).

## Regulatory

MCL (Drinking Water): (FL) FLORIDA 10 ug/l

[USEPA/Office of Water; Federal-State Toxicology and Risk Analysis Committee (FSTRAC). Summary of State and Federal Drinking Water Standards and Guidelines (11/93)]

Universal Treatment Standards (UTS): None

EPA RfD: none mg/kg-day

EPA RfC: none mg/m<sup>3</sup>

Drinking Water Remedial Target: 0.22 µg/L

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
0.44	2.2	0.0061	0.061

## Remediation

The toxicity of methylhydrazine to microbial species was found to be sufficiently high to prevent its degradation by biological waste treatment (Kane and Williamson 1983). Large amounts of methylhydrazine, such as might be released from a spill, are not expected to biodegrade. However, biodegradation of lower methylhydrazine concern may occur. Methylhydrazine at 500 mg/l, present in a wastewater mixture of hydrazine compounds, was incubated with an inoculum prepared from a trickling filter plant; following a 24 hour lag period, this mixture of compounds was biodegraded as measured by oxygen uptake (Wachinski and Farmwald 1980). Contaminated wastewater from the NASA Kennedy Space station was studied in batch cultures and trickle bed reactors to monitor the degradation rates of hydrazine containing rocket fuels (Nwankwoala et al. 1999a, b). The half-life of methylhydrazine in batch cultures (*Rhodococcus* B30 or *Achromobacter* sp.) and trickle bed reactors was about 2.5 and 12 days, respectively (Nwankwoala et al. 1999a). Biodegradation may be a significant removal process at low concentrations in soils or ambient waters, but at higher concentrations hydrazines are toxic to microorganisms (Nwankwoala et al. 1999b). Concentrations of hydrazine and 1,1-dimethylhydrazine that reduced bacterial metabolism by 50% ranged from 14.6 to 145 mg/L and from 19.2 to 9,060 mg/L, respectively (ATSDR 1998).

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## Methyl Isocyanate 624-83-9

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), methyl isocyanate, which has a vapor pressure of 348 mm Hg at 20 deg C (USEPA 1986), is expected to exist solely as a vapor in the ambient atmosphere. Vapor-phase methyl isocyanate is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals; the half-life for this reaction in air is estimated to be 118 days, calculated from its rate constant of  $1.4 \times 10^{-13}$  cu cm/molecule-sec at 25 deg C (SRC), that was derived using a structure estimation method (Meylan and , Howard 1993). Methyl isocyanate does not contain chromophores that absorb at wavelengths >290 nm and therefore is not expected to be susceptible to direct photolysis by sunlight (Lyman et al. 1994).

**Water:** If methyl isocyanate is released to water, it is expected to hydrolyze rapidly based on half-lives of 20 min at 15 deg C and 9 min at 25 deg C, calculated from measured hydrolysis rate constants of  $5.9 \times 10^{-4}$  sec<sup>-1</sup> and  $1.34 \times 10^{-3}$  sec<sup>-1</sup>, respectively (Castro et al. 1985). The products of hydrolysis may include N-carboxymethylamine, methylamine, carbon dioxide, and N, N'-dimethylurea (Castro et al. 1985, USEPA 1986). Bioconcentration, volatilization, adsorption to sediment and suspended solids, and biodegradation are not expected to be important fate processes since hydrolysis proceeds so quickly.

**Soil:** If methyl isocyanate is released to soil, it is expected to hydrolyze rapidly based on aqueous hydrolysis half-lives of 20 min at 15 deg C and 9 min at 25 deg C, calculated from measured rate constants of  $5.9 \times 10^{-4}$  sec<sup>-1</sup> and  $1.34 \times 10^{-3}$  sec<sup>-1</sup>, respectively (Castro et al. 1985). Soil adsorption effects, volatilization, and biodegradation are not expected to be competing processes since hydrolysis proceeds so quickly. Methyl isocyanate is expected to volatilize from dry soil surfaces (SRC) based upon a vapor pressure of 348 mm Hg (USEPA 1986).

**Biota:** Since **methyl isocyanate** hydrolyzes rapidly in water (Castro et al. 1985), biodegradation is not expected to be an important fate process.

### Regulatory

MCL (Drinking Water): none

Universal Treatment Standards (UTS): none

EPA Rf D: none

EPA RfC: none

Drinking Water Remedial Target: none

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m3	tap water ug/L
none	none	none	none

## Remediation

Since methyl isocyanate hydrolyzes rapidly in water (Castro et al. 1985), biodegradation is not expected to be an important fate process (SRC). Methyl isocyanate is expected to volatilize from dry soil surfaces (SRC) based upon an estimated vapor pressure of 348 mm Hg at 20 deg C (USEPA 1986).

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## **Monocrotophos 6923-22-4**

### **ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), monocrotophos, which has a measured vapor pressure of  $2.2 \times 10^{-6}$  mm Hg at 20 deg C (Tomlin 1994), is expected to exist both in the vapor and particulate phases in the ambient atmosphere. Vapor-phase monocrotophos is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals and ozone (SRC); the half-lives for these reactions in air are estimated to be about 9.1 hours and 24 hours, respectively (Meylan and Howard 1993). Particulate-phase monocrotophos may be physically removed from the air by wet and dry deposition (SRC).

**Water:** An estimated Koc value of 19(SRC), determined from an experimental log Kow of -0.20 (Hansch et al. 1995), indicates that monocrotophos is not expected to adsorb to suspended solids and sediment in water (Swann et al. 1983). Monocrotophos is not expected to volatilize from water surfaces (Lyman et al. 1990) based on an estimated Henry's Law constant of  $6.5 \times 10^{-13}$  atm-cu m/mole(SRC), determined from experimental values for vapor pressure,  $2.2 \times 10^{-6}$  mm Hg at 20 deg C (Tomlin 1994), and water solubility,  $1.0 \times 10^6$  mg/l at 20 deg C (Shiu et al. 1990). According to a classification scheme (Shiu et al. 1990), an estimated BCF value of 0.41(Lyman et al. 1990, SRC), from an experimental log Kow(2), suggests that bioconcentration in aquatic organisms is low(SRC). 100% of monocrotophos persisted after eight weeks incubation in river water (Eichelberger and Lichtenberg 1971).

**Soil:** An estimated Koc value of 19(SRC), determined from an experimental log Kow of -0.20 (Hansch et al. 1995) and a recommended regression-derived equation (Lyman et al. 1990), indicates that monocrotophos is expected to have very high mobility in soil (SRC) (Swann et al. 1983). Volatilization of monocrotophos from moist soil surfaces is not expected given an estimated Henry's Law constant of  $6.5 \times 10^{-13}$  atm-cu m/mole(SRC), determined from experimental values for vapor pressure,  $2.2 \times 10^{-6}$  mm Hg at 20 deg C (Tomlin 1994), and water solubility,  $1.0 \times 10^6$  mg/l at 20 deg C (Shiu et al. 1990). Monocrotophos is not expected to volatilize from dry soil surfaces based on a measured vapor pressure of  $2.2 \times 10^{-6}$  mm Hg. Monocrotophos is rapidly degraded in soil with half-lives ranging from 1 to 5 days (Tomlin 1994).

**Biota:** An estimated BCF value of 0.41 was calculated for monocrotophos, using an estimated log Kow of -0.20 (Hansch et al. 1995) and a recommended regression-derived

equation (Lyman et al. 1990). According to a classification scheme (Franke et al. 1994), this BCF value suggests that bioconcentration in aquatic organisms is low.

### Environmental partitioning

Photolysis of **monocrotophos** in soils, in water, and on bean plant leaves was investigated under sunlight and ultraviolet light. On a glass surface, percent recovery of **monocrotophos** in the dark was 98%, while 72.8% was recovered after 8 hr sunlight and 62% after UV light. Percent recovery of **monocrotophos** was 40.2 to 60.1% in soils after UV light exposure and 52.4 to 62.0% after sunlight exposure. Photodegradation in soils was of the order alluvial < black < red loamy < laterite. Percent recovery on bean leaves was 94% after 8 hr darkness and 68% after 8 hr sunlight. The rate of degradation of **monocrotophos** in tap water was approximately twice that in distilled water under sunlight (Dureja 1989).

### Regulatory

MCL (Drinking Water): 1.0 mg/L

Universal Treatment Standards (UTS): none

EPA Rf D: 0.0003 mg/kg-day

EPA RfC: 0.0005 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/l
none	none	none	none

### Remediation

Avoid skin contamination and inhalation of vapor. Cover contaminated areas and absorb spilled liquid with a 1:3 mixture of sodium carbonate crystals and damp sawdust, lime, sand, or earth. Sweep up and place in an impervious container. Ensure that container is tightly closed and labeled before transfer to a safe place for disposal (IPC).

Disposal: Large amounts should be incinerated at high temperature in a unit with effluent gas scrubbing. When no incinerator is available, bury in an approved dump, or in an area where there is no risk of contamination of surface or groundwater. Before burying, liberally mix with sodium carbonate (washing soda) crystals to help neutralize the product, and with soil rich in organic matter. Comply with any local legislation (IPC).

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## Nicotine 54-11-5

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), nicotine, which has a vapor pressure of 0.038 mm Hg at 25 deg C (Boublik et al. 1984), is expected to exist solely as a vapor in the ambient atmosphere. Vapor-phase nicotine is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 4 hours (Meylan and Howard 1993, SRC).

**Soil:** An estimated Koc value of 100(SRC) determined from a log Kow value (Hansch et al. 1995) and a regression-derived equation (Lyman WJ et al. 1990), indicates that nicotine is expected to have high mobility in soil (SRC) (Swann et al. 1983). However, nicotine is a base,  $pK_{b1} = 6.16$ ;  $pK_{b2} = 10.96$  (Tomlin 1994), and protonation under neutral and acidic conditions may result in greater adsorption and less mobility than its estimated Koc or water solubility (Budavari 1996) indicate (Khairy AH et al. 1995, SRC). Volatilization of nicotine from moist soil surfaces is not expected to be important (SRC) given an estimated Henry's Law constant of  $3.0 \times 10^{-9}$  atm-cu m/mole (SRC), using a fragment constant estimation method. Nicotine is not expected to volatilize from dry soil surfaces based on a vapor pressure of 0.038 mm Hg (Boublik T et al. 1994). Based on limited data, biodegradation of nicotine may occur with the formation of oxynicotine, 3-pyridylmethyl ketone, 2,3'-dipyridyl, N-methylmyosmine and an unknown purple crystalline pigment as reaction products (Shulka 1984).

**Sediment:** Nicotine is not expected to adsorb to suspended solids and sediment in water (SRC) (Swann et al. 1984). However, nicotine is a base,  $pK_{b1} = 6.16$ ;  $pK_{b2} = 10.96$  (4), and protonation under neutral and acidic conditions may result in greater adsorption than its estimated Koc value may indicate (Khairy et al. 1990, SRC). Based on limited data, biodegradation of nicotine may occur with the formation of oxynicotine, 3-pyridylmethyl ketone, 2, 3'-dipyridyl, N-methylmyosmine and an unknown purple crystalline pigment as reaction products (Shulka 1984). Nicotine is not expected to volatilize from water surfaces (Lyman et al. 1990, SRC) based on an estimated Henry's Law constant of  $3.0 \times 10^{-9}$  atm-cu m/mole (SRC), developed using a fragment constant estimation method (Meylan and Howard 1991).

**Biota:** According to a classification scheme by Franke et al. (1994), an estimated BCF of 5 (Lyman et al. 1990, SRC), from a log Kow value (Hansch et al. 1995), suggests the potential for bioconcentration in aquatic organisms is low.

## Environmental partitioning

Nicotine is a natural constituent of the tobacco plants *Nicotiana tabacum* and *N. rustica* and it is released in the particulate phase of tobacco smoke. Nicotine may also be released to the environment due to its use as an insecticide and use in formulations for painting surfaces in poultry houses. If released to soil, nicotine may biodegrade to a variety of different products including oxynicotine and 3-pyridylmethyl ketone. Under alkaline conditions this compound should be extremely mobile. Nicotine is not expected to chemically hydrolyze or volatilize from moist soil surfaces. If released to water, nicotine may biodegrade. Slight potential also exists for photolysis. Chemical hydrolysis, bioaccumulation in the aquatic organisms, volatilization and adsorption under alkaline conditions to suspended solids and sediments are not expected to be significant fate processes. If released to air, nicotine is expected to degrade rather quickly in the presence of light and air. In the ambient atmosphere, nicotine may react with photochemically generated hydroxyl radicals (vapor phase half-life 1 day), be removed by wet deposition or photolyze.

## Regulatory

IDLH/NIOSH: 5 mg/m<sup>3</sup>

MCL (Drinking Water): none

Universal Treatment Standards (UTS): none

EPA Rf D: none

EPA RfC: none

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/l
none	none	none	none

## Remediation

Nicotine is a potential candidate for rotary kiln incineration at a temperature range of 820 to 1,600 deg C and residence times of seconds for liquids and gases, and hours for solids. It is also a potential candidate for fluidized bed incineration at a temperature range of 450 to 980 deg C and residence times of seconds for liquids and gases, and longer for solids (USEPA 1981).

The most probable route of human exposure is by inhalation of tobacco smoke. Some people may also be exposed to nicotine in drinking water. Infants breast fed by women who smoke are exposed to nicotine in mother's milk. Worker exposure may occur during formulation, storage and application to pesticides and paints which contain this compound and during processing and extraction of tobacco (Spectrum).

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## Nitric Acid

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** In the atmosphere, nitric acid is present in very minute quantities as a gas or vapor. It reacts with ammonia in the atmosphere to form a nitrate compound that is unlikely to persist for more than 10 to 15 days. Nitric acid does not build up in plant or animal tissues largely because of its highly reactive properties. In a concentrated form, nitric acid reacts very quickly with many other commonly occurring natural or manufactured substances and compounds. In some cases this can produce a vigorous reaction in the form of a chemical fire or explosion (**National Pollutant Inventory**).

**Water:** Elevated nitrate levels will stimulate plankton and aquatic weed growth.

**Soil:** During transport through the soil, **nitric acid** will dissolve some of the soil material; in particular, the carbonate based materials. The acid will be neutralized to some degree with adsorption of the proton also occurring on clay materials. However, significant amounts of acid are expected to remain for transport down toward the ground water table. Upon reaching the ground water table, the acid will continue to move, now in the direction of the ground water flow. A contaminated plume will be produced with dilution and dispersion serving to reduce the acid concentration (Environment Canada).

**Environmental Partitioning:** Nitric acid exists in the atmosphere in the gas phase. The atmospheric removal processes for gaseous nitric acid are by wet and dry deposition. The estimated half-life and lifetime for dry deposition of nitric acid is 1.5 to 2 days and 2 to 3 days, respectively, and efficient rain out during episodic precipitation events. Nitric acid reacts with gaseous ammonia to form particulate or aerosol nitrate, which in turn is removed by wet and dry deposition of the particles. The average half-life and lifetime for particles in the atmosphere is about 3.5 to 10 days and 5 to 15 days (National Pollutant Inventory).

## Remediation

**Land spill:** Dig a pit, pond, lagoon, or holding area to contain liquid or solid material. /SRP: If time permits; pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash or cement powder. Neutralize with agricultural lime (CaO), crushed limestone or sodium bicarbonate.

Water spill: Neutralize with agricultural lime (slaked lime), crushed limestone, or sodium bicarbonate.

Air Release: Apply water spray or mist to knock down vapors. Vapor knockdown water is corrosive or toxic and should be diked for containment (Association of American Railroads).

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## Nitrogen Dioxide 10102-44-0

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Nitrogen dioxide, which has a vapor pressure of 900 mm Hg at 25 deg C (Daubert and Danner 1989), is expected to exist solely as a gas in the ambient atmosphere. Gas-phase nitrogen dioxide is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 35 hours (SRC) from its rate constant of  $1.1 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C (Grosjean 1985). This compound is expected to undergo direct photolysis in the environment (SRC), ultimately leading to the production of ozone and smog conditions in the lower troposphere (Graham 1997).

**Water:** If released into water, nitrogen dioxide is expected to decompose (Graham et al. 1997) to nitric acid.

**Soil:** Nitrogen dioxide is expected to decompose to nitric acid in moist soil surfaces (Graham et al. 1997). The potential for volatilization of nitrogen dioxide from dry soil surfaces may exist based upon a vapor pressure of 900 mm Hg at 25 deg C (Daubert and Danner 1989).

### Remediation

Environmental considerations (air spill): Apply water spray or mist to knock down vapors. Vapor knockdown water is corrosive or toxic and should be diked for containment (Association of American Railroads).

### Regulatory

EPA RfD: has been withdrawn from IRIS as of 09/01/1994

EPA RfC: none

IDLH: 20 ppb

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m3	tap water ug/L
none	none	none	none

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## Nitroglycerine 55-63-0

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), nitroglycerin, which has a vapor pressure of  $2.0 \times 10^{-4}$  mm Hg at 20 deg C (Lyman 1985), is expected to exist in both the vapor and particulate phases in the ambient atmosphere. Vapor-phase nitroglycerin is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 15 days (SRC), calculated from its rate constant of  $1.1 \times 10^{-12}$  cu cm/molecule-sec at 25 deg C (SRC) that was derived using a structure estimation method (Meylan and Howard 1985). Particulate-phase nitroglycerin may be removed from the air by wet and dry deposition (SRC). Nitroglycerin absorbs light above 290 nm weakly and may be susceptible to direct photolysis (Spanggord et al. 1980).

**Water:** An estimated Koc value of 180 (SRC), determined from a log Kow of 1.62 (Hansch et al. 1995) and a regression-derived equation (Lyman et al. 1990), indicates that nitroglycerin is expected to adsorb to suspended solids and sediment (Swann et al. 1983). Volatilization from water surfaces is not expected (Lyman et al. 1990) based upon an estimated Henry's Law constant of  $4.3 \times 10^{-8}$  atm-cu m/mole. According to a classification scheme (Franke et al 1994), an estimated BCF of 4 (SRC), from its log Kow (Hansch et al. 1995) and a regression-derived equation (Meylan et al. 1999), suggests the potential for bioconcentration in aquatic organisms is low (SRC). Hydrolysis is expected to be an important environmental fate process under alkaline conditions based on half-lives of 37 and 96 days at pH 9 and 25 and 18 deg C, respectively (Cappellos et al. 1984).

**Soil/ Sediment:** An estimated Koc value of 180 (SRC), determined from a log Kow of 1.62 (Hansch et al. 1995) and a regression-derived equation (Lyman et al. 1990), indicates that nitroglycerin is expected to have moderate mobility in soil (Swann et al. 1983). Volatilization of nitroglycerin from moist soil surfaces is not expected to be an important fate process (SRC) given an estimated Henry's Law constant of  $4.3 \times 10^{-8}$  atm-cu m/mole (SRC), derived from its vapor pressure,  $2.0 \times 10^{-4}$  mm Hg, and water solubility, 1,380 mg/l (Seidell 1941). Nitroglycerin is not expected to volatilize from dry soil surfaces (SRC) based upon its vapor pressure. Hydrolysis under alkaline conditions may be important based upon half-lives of 37 and 96 days at pH 9 and 25 and 18 deg C, respectively (Cappellos et al. 1984). Biodegradation of nitroglycerin occurs rapidly using batch reactors and digestors (Christodoulatos et al. 1997, Accashian et al. 2000), but its rate of biodegradation in soil is unknown (SRC). Nitroglycerin (10 ppm initial concentration)

was completely biodegraded in 13 days using river water and river water/sediment microcosms obtained from a river near a munitions facility (Spanggord et al. 1980), which may suggest biodegradation can occur in soils (SRC).

**Biota:** According to a classification scheme (Franke et al 1994), an estimated BCF of 4 (SRC), from its log Kow (Hansch et al. 1995) and a regression-derived equation (Meylan et al. 1999), suggests the potential for bioconcentration in aquatic organisms is low (SRC).

## Regulatory

Drinking Water Remedial Target: 48 µg/l

DWEL (Drinking Water): 5 ug/l

Universal Treatment Standards (UTS): none

EPA RfD: 0.0001mg/kg-day (provisional)

EPA RfC: none mg/m<sup>3</sup>

IDLH: 75 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
35	18	0.48	4.8

## Remediation

Nitroglycerin and other aliphatic nitric acid esters undergo aerobic biodegradation readily via successive removal of nitrate groups to isomeric derivatives (Rendic S et al. 1995). Contrary to some earlier reports that it was recalcitrant to biodegradation, nitroglycerin proved to be readily biodegradable in batch and continuous tests (Wendt et al. 1978). Results of the shake flask test that was run at 30 deg C with an activated sludge inoculum showed a 53.6% loss of nitroglycerin in 5 days (Wendt et al. 1978). In a continuous bench-scale activated sludge apparatus, a mean reduction of 92.2% was obtained (Wendt et al. 1978). The apparatus was run at room temperature and the residence time was 84 hr. 1, 3-Dinitroglycerol and 1,2-dinitroglycerol were identified at intermediate stages of the process, but they were also essentially absent from the effluent (Wendt et al. 1978). The third experiment employing a chemostat, a continuous culture, aerobic system with no solids recycling, was designed to simulate a plant where propellant wastes would be treated. After a 8-15 hr detention period, no nitrate esters

were detectable in the effluent (Wendt et al. 1978). It was found that nitroglycerin is not suitable as a source of carbon and nitrogen so nutrients are essential (Wendt et al. 1978). It was speculated that the earlier experiments which showed nitroglycerin to be recalcitrant were conducted using concentrations of nitroglycerin that were toxic to the microorganisms (Wendt et al. 1978).

Complete mineralization of nitroglycerin (350  $\mu\text{mol/l}$  initial concentration) was achieved within 4 days utilizing mixed cultures and an anaerobic digester, using digester sludge from a municipal wastewater treatment plant amended with 2,000 mg/l of glucose (Christodoulatos et al.1997). Without glucose amendments, the complete mineralization of nitroglycerin took about 25 days in the digester (Christodoulatos et al.1997). Nitroglycerin, present at 16 mg/l, was rapidly degraded in a sequencing batch reactor using mixed liquor from an industrial wastewater facility and a domestic wastewater treatment plant, with 1,2-dinitroglycerol and 1,3-dinitroglycerol as the initial products (Accashian JV et al.2000).

Nitroglycerin (10 ppm initial concentration) was completely biodegraded in 13 days using river water and river water/sediment microcosms obtained from a river near a munitions facility in Virginia. Pure cultures isolated from river water and sediment near a munitions factory in Virginia degraded nitroglycerin. The first-order rate constant was 0.60  $\text{hours}^{-1}$ , which corresponds to a half-life of about 1 hour (Spanggord et al.1995).

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## Parathion 56-38-2

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), parathion, which has a vapor pressure of  $6.68 \times 10^{-6}$  mm Hg at 20 deg C (Tomlin 1997), will exist in both the vapor and particulate phases in the ambient atmosphere. Vapor-phase parathion is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 4.2 hours (SRC) from the estimated rate constant (Meylan and Howard 1993). Particulate-phase parathion may be removed from the air by wet and dry deposition (SRC). Parathion absorbs radiation with wavelengths shorter than 320 nm, suggesting a potential for direct photolysis (SRC).

**Water:** Based on a classification scheme (Swann et al. 1983), sediment Koc values ranging from 3,086 to 38,000 (House et al. 1992) indicate that some adsorption of parathion to suspended solids and sediment in the water column is expected (SRC). Parathion is not expected to volatilize from water surfaces (Lyman et al. 1990) based upon a Henry's Law constant of  $2.98 \times 10^{-7}$  atm-cu m/mole (Tomlin 1997). According to a classification scheme (Franke et al. 1994), BCFs ranging from 63 to 462 (USEPA), suggest bioconcentration in aquatic organisms is moderate to high (SRC). Irradiation of parathion for 10 hr in aerated distilled water resulted in 88% degradation (Mansour et al. 1983). 20% of parathion was lost by photolysis in 2 hr in Okefenokee Swamp water (Zepp and Baughman 1978). Reported hydrolysis half-lives at 20 deg C at environmentally relevant pHs range from 3 weeks at pH 9 (Faust and Gomaa 1972) to 43 weeks at pH 5; the half-life for hydrolysis in sterile sea water has been reported to be approximately 1 yr at 4 deg C. It is thought that divalent cations catalyze hydrolysis (Wade 1979). Parathion biodegrades in acclimated natural waters within several weeks (Eichelberger and Lichtenberg 1971). Parathion, 5 ppm, completely degraded within 2 weeks in acclimated water from Holland Marsh, Ontario, being almost quantitatively converted to aminoparathion; only 10% degradation occurred in 16 weeks when the water was sterilized (Sharom 1980). After 30 days incubation in non-sterile (sterile) coastal river water, only 21, 14, and 6% (64, 57, and 49%) of parathion remained at pH values of 6, 7, and 8.16, respectively (Wang and Hoffman 1991).

**Soil:** Persistence of parathion was partially dependent on soil type. In some soils degradation was rapid & probably through combination of hydrolysis & strong microbial activity. In other soils, loss was slow & attributable to hydrolysis (Menzie 1978). Based on a classification scheme (Swann et al. 1983), Koc values ranging from 314 to 15,860

(Mingelgrin and Gerstl 1983) indicate that parathion is expected to have moderate to no mobility in soil (SRC). Volatilization of parathion from moist soil surfaces is not expected to be an important fate process (SRC) given a Henry's Law constant of  $2.98 \times 10^{-7}$  atm-cu m/mole (Tomlin 1997). Parathion is not expected to volatilize from dry soil surfaces based upon a vapor pressure of  $6.68 \times 10^{-6}$  mm Hg at 20 deg C (Tomlin 1997). The half-life for photodecomposition of parathion on 3 soils with various water contents ranged from 31 to 70 hr (Hautala 1978). After 8 weeks of incubation in an organic and a mineral soil, <2 and 6%, respectively, of the 1 ppm parathion applied remained; in sterilized controls 80 and 95% remained (Chapman et al. 1981). Prior exposure of soils to p-nitrophenol resulted in increased mineralization of parathion to carbon dioxide (Somasundaram and Coats 1990). Metabolic pathways involve both oxidative and reductive reactions. The primary oxidative pathway involves an initial hydrolysis to p-nitrophenol and diethylthiophosphoric acid; a second oxidative pathway involves oxidation to paraoxon. Under low oxygen conditions reduction to aminoparathion occurs (Miles et al. 1979).

Biota: Tadpoles, a species resistant to cholinesterase inhibitors such as organophosphate pesticides, had an average bioconcentration factor of 64 (Hall and Kolbe 1980). BCFs ranging from 63 to 462 were observed in bluegill after 0.5 and 3 days exposure to parathion concentrations of 510 and 640 ug/l, respectively (USEPA 1998). BCFs ranging from 68 to 344 were observed in brook trout after 0.33 and 5.83 days exposure to parathion concentrations of 3,180 and 270 ug/l, respectively (USEPA 1998). In a terrestrial-aquatic ecosystem, there was no evidence for bioaccumulation in any of the organisms, including snails, algae, daphnia, mosquito larvae, and fish, at the end of the 38-day experiment, although the ecosystem organisms contained radiolabeled carbon indicating the presence of metabolism products (Sanborn et al. 1977). There is no evidence of bioaccumulated parathion in cattle, sheep, or rabbits (Sanborn et al. 1977). The average BCF in killifish after 24 to 72 hours exposure to the individual pesticide was 98; the average BCF was 88 in killifish after 24 to 72 hours exposure to a mixture of pesticides, including parathion (Tsuda et al. 1995). According to a classification scheme (Franke et al. 1990), these BCFs suggest bioconcentration in aquatic organisms is moderate to high.

## Regulatory

MCL (Drinking Water): none

Universal Treatment Standards (UTS): 0.014 mg/l<sup>2</sup> in wastewater standards

EPA RfD: 0.006 mg/kg-day (provisional)

EPA RfC: none mg/m<sup>3</sup>

IDLH: 10 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
370	5300	22	220

## Remediation

Land spill: Dig a pit, pond, lagoon, holding area to contain liquid or solid material. SRP: If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash, cement powder, or commercial sorbents. Apply "universal" gelling agent to immobilize spill. Parathion and compressed gas mixture (Association of American Railroads).

Water spill: Use natural barriers or oil spill control booms to limit spill travel. Use surface active agent (eg, detergent, soaps, alcohols), if approved by EPA. Inject "universal" gelling agent to solidify encircled spill and increase effectiveness of booms. If dissolved, in region of 10 ppm or greater concentration, apply activated carbon at ten times the spilled amount. Use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates. Parathion and compressed gas mixture (Association of American Railroads).

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**Phenol****108-95-2****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** The gas-phase reaction of phenol with photochemically produced hydroxyl radicals is probably a major removal mechanism in the atmosphere. An estimated half-life for phenol for this reaction is 0.61 days (EPA 1979a). The reaction of phenol with nitrate radicals during the night may constitute a significant removal process. This is based on a rate constant of  $3.8 \times 10^{-12}$  cm<sup>3</sup>/molecule second for this reaction, corresponding to a half-life of 15 minutes at an atmospheric concentration of  $2 \times 10^8$  nitrate radicals per cm<sup>3</sup> (Atkinson et al. 1987). The reaction of phenol with nitrate radicals present in the atmosphere during smog episodes may decrease the half-life of phenol in polluted atmospheres. The above data indicate that phenol has a short half-life in the atmosphere, probably <1 day. Phenol does not absorb light in the region of 290–330 nm (Lide and Milne 1994); therefore, it should not photo-degrade directly in the atmosphere.

**Water:** Although phenol does not absorb light at wavelengths >290, phenols react rapidly to sunlit natural water via an indirect reaction with photochemically produced hydroxyl radicals and peroxy radicals; typical half-lives for hydroxyl and peroxy radical reactions are on the order of 100 and 19.2 hours of sunlight, respectively (Canonica et al. 1995; Mill and Mabey 1985). These reactions require dissolved natural organic materials that function as photosensitizers (Canonica et al. 1995). The estimated half-life for the reaction of phenol with photochemically produced singlet oxygen in sunlit surface waters contaminated by humic substances is 83 days (assuming Switzerland summer sunlight and a singlet oxygen concentration of  $4 \times 10^{-14}$  molar [M]) (Scully and Hoigne 1987). Phenol is readily biodegradable in natural water, provided the concentration is not high enough to cause significant inhibition through microbial toxicity. Complete degradation in <1 day has been reported in water from three lakes; the rates of degradation were affected by the concentration of organic and inorganic nutrients in the water (Rubin and Alexander 1983). Complete removal of phenol in river water has been reported after 2 days at 20 °C and after 4 days at 4 °C (Ludzack and Ettinger 1960). The degradation of phenol is somewhat slower in salt water, and a half-life of 9 days has been reported in an estuarine river (EPA 1979b).

**Sediment and Soil:** Available data indicate that phenol biodegrades in soil under both aerobic and anaerobic soil conditions. The half-life of phenol in soil is generally <5 days (Baker and Mayfield 1980), but acidic soils and some surface soils may have half-lives of up to 23 days (Shiu et al. 1994). Mineralization in an alkaline, parabrown soil under aerobic conditions was 45.5, 48, and 65% after 3, 7, and 70 days, respectively (Haider et

al. 1974). Half-lives for degradation of low concentrations of phenol in two silt loam soils were 2.70 and 3.51 hours (Scott et al. 1983). Plants have been shown to be capable of metabolizing phenol readily (DOE 1987). While degradation is slower under anaerobic conditions, evidence presented in the literature suggests that phenol can be rapidly and virtually completely degraded in soil under both aerobic and anaerobic conditions (Baker and Mayfield 1980).

**Biota:** Phenol is not expected to bioconcentrate significantly in aquatic organisms. Reported log bioconcentration factors (BCF) in fish for phenol include 0.28 for goldfish, (Kobayashi et al. 1979) and 1.3 for golden orfe (Freitag et al. 1984). The highest mean level of phenol detected in bottom fish from Commencement Bay in Tacoma, Washington, was 0.14 ppm (Nicola et al. 1987). The levels of phenol in the water or sediments were not stated.

## Environmental partitioning

Phenol is released into the air and discharged into water from both manufacturing and use. Based on its high water solubility and the fact that it has been detected in rainwater, some phenol may wash out of the atmosphere; however, it is probable that only limited amounts wash out because of the short atmospheric half-life of phenol. During the day, when photochemically produced hydroxyl radical concentrations are highest in the atmosphere, very little atmospheric transport of phenol is likely to occur.

In water, neither volatilization nor sorption to sediments and suspended particulates are expected to be important transport mechanisms. Using the Henry's Law constant of  $3 \times 10^{-7}$  atm m<sup>3</sup>/mol (Gaffney et al. 1987), a volatilization half-life of 88 days was calculated for phenol evaporation from a model river 1 m deep with a current of 1 m/second, and wind velocity of 3 m/second (Lyman et al. 1982). The biological treatment of waste water containing phenol has shown that <1% of phenol is removed by stripping (Kincannon et al. 1983; Petrasek et al. 1983).

The pKa of phenol is 10 (O'Neil 2001), indicating that phenol will primarily exist as the protonated acid at environmental pH values. In alkaline soils and water, phenol will partially exist as an anion, which can affect its fate and transport processes. Although it has been shown that plants readily uptake phenol (DOE 1987), bioaccumulation does not take place due to a high rate of respiratory decomposition of phenol to CO<sub>2</sub>.

## Regulatory

MCL (Drinking Water): 0.3 mg/L cyanide

Universal Treatment Standards (UTS): 0.039 mg/l<sup>2</sup> in wastewater standard

6.2 mg/L in nonwastewater standard

EPA RfD: 0.6 mg/kg-day

EPA RfC: 0.006 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/l
37000	10000	2200	22000

## Remediation

Phenol is quickly broken down in the air, usually within 1–2 days.

Phenol may persist in water for a week or more.

Phenol that remains in soil may be broken down by bacteria or other microorganisms.

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## Phosgene 75-44-5

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), phosgene, which has a vapor pressure of 1420 mm Hg at 25 deg C (Daubert and Danner 1989), is expected to exist solely in the gas phase in the ambient atmosphere. Gas-phase phosgene is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 44 years (SRC), calculated from its rate constant of  $1.0 \times 10^{-15}$  cu cm/molecule-sec at 25 deg C (Atkinson 1994). Phosgene does not absorb UV radiation at wavelengths >290 nm (Montgomery and Rollefson 1993) and therefore would not be subject to direct photolysis by sunlight in the troposphere. Phosgene's hydrolytic half-life in the atmosphere, extrapolating from high temperature data, is 113 yr at sea level, assuming 10 torr of water vapor (Butler et al 1979).

**Water:** Based on a classification scheme (Swann et al 1983), an estimated Koc value of 2.2 (SRC), determined from a structure estimation method (Meylan et al 1992), indicates that phosgene is not expected to adsorb to suspended solids and sediment (SRC).

Volatilization from water surfaces is expected (Lyman et al 1990) based upon a Henry's Law constant of  $1.7 \times 10^{-2}$  atm-cu m/mole (De Bruyn et al 1995). Phosgene has been reported to hydrolyze rapidly in seawater (Talmage et al 2007). Hydrolysis of a 1% solution of phosgene in water is complete within 20 seconds at 0 deg C (Mitretek Systems 2000). Measured aqueous phase hydrolysis rate constants for phosgene have ranged from 0.29 to 100 sec<sup>-1</sup> (Kolb et al 1995). The rapid hydrolysis suggests that bioconcentration will not be an important environmental fate process. Biodegradation data were not available (SRC, 2007).

**Soil/Sediment:** Based on a classification scheme (Swann et al 1983), an estimated Koc value of 2.2, determined from a structure estimation method (Meylan et al 1992), indicates that phosgene is expected to have very high mobility in soil (SRC). Volatilization of phosgene from moist soil surfaces is expected to be an important fate process (SRC) given a Henry's Law constant of  $1.7 \times 10^{-2}$  atm-cu m/mole (De Bruyn et al 1995). Phosgene is expected to hydrolyze in moist soil; hydrolysis of a 1% solution of phosgene in water is complete within 20 seconds at 0 deg C (Mitretek Systems 2000). Phosgene is expected to volatilize from dry soil surfaces (SRC) based upon a vapor pressure of 1420 mm Hg (Daubert and Danner 1989). Biodegradation data were not available (SRC, 2007).

**Biota:** Phosgene does not accumulate in the food chain.

## Environmental Partitioning

When released to air, phosgene will exist solely as a gas. Phosgene gas is degraded in the atmosphere by reacting with substances commonly found in the air, but this is a very slow process. Phosgene in the air may also react with moisture in clouds or rain and be broken down into other compounds.

Phosgene will react with water and be broken down into other products. Some of the phosgene that is not broken down may evaporate into air.

When released to soil, phosgene will not stick to the soil. Small amounts may evaporate into air or pass through the soil surface and contaminate groundwater. Most of the phosgene in soil will be broken down when it comes into contact with moisture.

## Regulations

MCL (Drinking Water): none ug/L

Universal Treatment Standards (UTS): none in wastewater standards and none in non wastewater standards

Groundwater Water Quality Criterion : none ug/L

PQL: none µg/g

IDLH: 2 ppm

EPA RfD: none

EPA RfC: none

## Remediation

Liquid spills can be covered with sodium hydrogen carbonate or an equal mixture of soda ash and slake lime or crystallized urea. Water from an atomizer can then be added cautiously and the mixture transferred to a large volume of water. Gas spills can be mitigated by gaseous ammonia, aqueous ammonia or an ammonia steam curtain or sprays. There should be a holding area beneath any storage or handling installation that can contain a liquid spill. This should have an impermeable flexible membrane liner and must already contain lime, limestone, sodium hydrogen carbonate, urea or any other suitable neutralizing absorbent, sufficient to eliminate the spill (IPCS 2007).

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**Phosphamidon 13171-21-6****Dimethoate 60-51-5 (synonym)****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), phosphamidon, which has a vapor pressure of  $1.65 \times 10^{-5}$  mm Hg at 25 deg C (Tomlin 1997), will exist in both the vapor and particulate phases in the ambient atmosphere. Vapor-phase phosphamidon is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 10.5 hrs (SRC), calculated from its rate constant of  $3.7 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C (SRC) that was derived using a structure estimation method (Meylan and Howard 1993). Particulate-phase phosphamidon may be removed from the air by wet and dry deposition (SRC). The rate constant for the vapor-phase reaction of phosphamidon with ozone has been estimated as  $1.6 \times 10^{-16}$  cu cm/molecule-sec at 25 deg C (SRC), derived using a structure estimation method (Bidleman 1988); this corresponds to an atmospheric half-life of about 7 days at an atmospheric concentration of  $7 \times 10^{11}$  ozone molecules per cu cm (Atkinson and Carter 1984).

**Water:** Based on soil thin layer chromatography study, retardation factors (Rf) of 0.91-0.92 in a silt loam and a fine sandy loam (Khan and Khan 1986) indicate that phosphamidon is not expected to adsorb to suspended solids and sediment (SRC). Volatilization from water surfaces is not expected (Lyman et al. 1990) based upon an estimated Henry's Law constant of  $1.5 \times 10^{-12}$  atm-cu m/mole (SRC), developed using a fragment constant estimation method (Meylan and Howard 1991). According to a classification scheme (Franke et al 1994), a BCF of <1 (Min and Cha 2000) suggests the potential for bioconcentration in aquatic organisms is low.

**Soil:** Based on soil thin layer chromatography study, retardation factors (Rf) of 0.91-0.92 in a silt loam and a fine sandy loam (Khan and Khan 1986), indicate that phosphamidon has very high mobility in soil (SRC). Volatilization of phosphamidon from moist soil surfaces is not expected to be an important fate process (Lyman et al. 1990) given an estimated Henry's Law constant of  $1.5 \times 10^{-12}$  atm-cu m/mole (SRC), using a fragment constant estimation method (Meylan and Howard 1991). Phosphamidon is not expected to volatilize from dry soil surfaces (SRC) based upon a vapor pressure of  $1.65 \times 10^{-5}$  mm Hg (Tomlin 1997). Half-lives in loam, loamy sand, and sand of approximately 6, 3, and <3 days, respectively (Benyon et al. 1973), indicate that biodegradation of phosphamidon

may be an important environmental fate process in soil.

**Sediment:** Based on soil thin layer chromatography study, retardation factors (Rf) of 0.91-0.92 in a silt loam and a fine sandy loam (Khan and Khan 1986) indicate that phosphamidon is not expected to adsorb to suspended solids and sediment (SRC). Volatilization from water surfaces is not expected (Lyman et al. 1990) based upon an estimated Henry's Law constant of  $1.5 \times 10^{-12}$  atm-cu m/mole(SRC), developed using a fragment constant estimation method (Meylan and Howard 1991). According to a classification scheme (Franke et al. 1994), a BCF of <1 (Min and Cha 2000) suggests the potential for bioconcentration in aquatic organisms is low (SRC). Phosphamidon's hydrolysis half-lives at 23 deg C are 74 days, 13.8 days, and 2.2 days at pH 4, 7, and 10, respectively.

## Regulatory

Universal Treatment Standards (UTS): None

Acceptable Daily Intake (ADI) for dimethoate for a human is 0.02 mg/kg/day

CDHS' Archived Advisory Levels for Drinking Water: 0.001 mg/L

EPA Rf D: 0.0002 mg/kg-day

EPA RfC: none mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
120	180	0.73	7.3

## Remediation

**Air spill:** Apply water spray or mist to knock down vapors. /Organophosphorus pesticides, liquid, flammable, toxic; Organophosphorus pesticides, liquid, toxic; Organophosphorus pesticides, solid, toxic.

**Water spill:** Use natural barriers or oil spill control booms to limit spill travel. Remove trapped material with suction hoses. /Organophosphorus pesticides, liquid, flammable, toxic; Organophosphorus pesticides, liquid, toxic; Organophosphorus pesticides, solid, toxic.

**Land spill:** Dig a pit, pond, lagoon, holding area to contain liquid or solid material. /SRP: If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed

polyurethane, or foamed concrete. Absorb bulk liquid with fly ash, cement powder, or commercial sorbents. Organophosphorus pesticides, liquid, flammable, toxic; Organophosphorus pesticides, liquid, toxic; Organophosphorus pesticides, solid, toxic (Association of American Railroads).

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## **Phosphorous Trichloride      7719-12-2**

### **ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** White phosphorus can enter air from its production, use, accidental spills during loading and unloading for shipment and accidental spills during transport. During white phosphorus production, an estimated 0.58 kg total phosphorus/ton of product is emitted into the air (EPA 1989). Part of the air emissions during production of white phosphorus from phosphate rock is expected to be in the form of phosphate.

The air emission of elemental phosphorus during production of munitions at the Pine Bluff Arsenal in Arkansas was estimated at 1 mg/kg of phosphorus used (Berkowitz et al. 1981). The amount of elemental phosphorus released to air during the manufacture of phosphoric acid, phosphates, and other phosphorus chemicals is not known because the air emission factor is estimated for total phosphorus (EPA 1989).

However, the emission of elemental phosphorus from the manufacture of phosphorus chemicals may be low, because the conversion rate from elemental phosphorus to compounds is high. White phosphorus is also released in air during its use as an incendiary device by the military. It will enter the atmosphere during testing, which produces phosphorus smoke in the field. At a smoke density of 0.1 mg/m<sup>3</sup>, the estimated concentration of elemental phosphorus in the smoke was 21 ppb (EPA 1991).

**Water:** The very small amounts of phosphorus trichlorides and oxychlorides that actually enter the environment react rapidly with water to form hydrochloric and phosphorus oxyacids. After deposition, these oxyacids are neutralized and/or buffered to form chloride (by oxidation of phosphite) and phosphate salts. The quantities of these chlorides and phosphate salts are not significant.

**Sediment and Soil:** The exposure of soil to phosphorus aerosols will upset the pH of the soil and create a more acidic layer of soil. This decrease in pH can exceed the buffering capacity of the surface layer of the soil depending upon the amount of applied phosphorus. This process can be mitigated by a larger soil volume area so that phosphorus speciation on a field scale will be minimal. The interaction of metals with phosphorus condensates could lead to their leachability and possible trace metal migration from the soil (Van Vorris et al. 1987).

**Biota:** Phosphorus trichloride and phosphorus oxychloride react with the water component of the tissues they first contact. The resulting hydrogen chloride and

phosphorus oxyacid ions, if absorbed, join the body pools of these ions. Phosphate, chloride, and hydrogen ions are easily excreted by the kidneys by normal physiological mechanisms.

Phosphorus trichloride and phosphorus oxychloride hydrolyse by reacting 3 moles of water per mole and yielding 6 moles of hydrogen ion per mole. As liquids or in high vapor concentrations, they kill most animal and plant tissues they contact. However, the dehydrating and acidifying effects of major releases are transient and local.

## Environmental Partitioning

The persistence of white phosphorus in air is very short and may range from minutes to days. Particulate white phosphorus present as an aerosol may be coated with a protective layer of oxide and may have a longer lifetime in air (Berkowitz et al. 1981). In addition to aerosol age, phosphorus aerosol speciation is also affected by the humidity of the ambient environment (Van Voris et al. 1987). Washout and rainout processes transport both the reaction products of vapor phase phosphorus and unreacted particles of phosphorus to water and land (Berkowitz et al. 1981). Because of its lower water solubility, physical state (gas), and slower reactivity, phosphine formed during the combustion of white phosphorus or released to the atmosphere from other media persists in the atmosphere longer than other reaction products.

## Regulatory

IDLH: 25 ppm

Universal Treatment Standards (UTS): None

EPA RfD: 10 mg/kg-day

EPA RfC: none mg/m<sup>3</sup>

Drinking Water Remedial Target: none µg/L

## Remediation

Water Spill: Neutralize with agricultural lime (CaO), crushed limestone (CaCO<sub>3</sub>), or sodium bicarbonate (NaCO<sub>3</sub>). Use mechanical dredges or lifts to remove immobilized masses of pollutant and precipitates. Adjust pH to neutral (pH=7) Air Spill: Apply water spray or mist to knock down vapors. Vapor knockdown water is corrosive or toxic and should be diked for containment.

Land Spill: Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash or cement powder. Neutralize with agricultural lime (CaO), crushed limestone (CaCO<sub>3</sub>), or sodium bicarbonate (NaHCO<sub>3</sub>) (Association of American Railroads).

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**Phosphoryl trichloride      10025-87-3****(Phosphorus oxychloride)****ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), phosphorus oxychloride, which has a vapor pressure of 40 mm Hg at 27.3 deg C (Lewis 1999), is expected to exist solely as a vapor in the ambient atmosphere. Vapor-phase phosphorus oxychloride is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is reported to be >290 days, calculated from its rate constant of  $<4 \times 10^{-14}$  cu cm/molecule-sec at 25 deg (Martin et al. 2002). Phosphorus oxychloride does not contain chromophores that absorb at wavelengths >290 nm (Lyman et al. 1990) and therefore is not expected to be susceptible to direct photolysis by sunlight (SRC).

**Water:** Phosphorus oxychloride hydrolyzes rapidly in water at room temperature (Lewis 1999).

**Soil:** Hydrolysis is expected to be the most important fate process for phosphorus oxychloride in the environment (Lewis 1999). Phosphorus oxychloride is expected to volatilize from dry soil surfaces (SRC) based upon a vapor pressure of 40 mm Hg (Lewis 1999).

**Biota:** Hydrolyzes rapidly in water at room temperature (Lewis 1999), and therefore is not expected to bioconcentrate in organisms.

**Regulatory**

MCL (Drinking Water): none

Universal Treatment Standards (UTS): None

EPA RfD: none mg/kg-day

EPA RfC: none mg/m<sup>3</sup>

Drinking Water Remedial Target: none µg/L

## Remediation

Keep material out of water sources and sewers; build dikes to contain flow as necessary; use water spray to knock down vapors; do not use water on material itself; and neutralize spilled material with crushed limestone, soda ash, or lime. For a land spill, dig a pit, pond, lagoon, or holding area to contain liquid or solid material; dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete; absorb bulk liquid with fly ash or cement powder; neutralize with agricultural lime (slaked lime), crushed limestone, or sodium bicarbonate. For a water spill, neutralize with agricultural lime (slaked lime), crushed limestone, or sodium bicarbonate; use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates; adjust pH to neutral (pH 7). For air spills apply water spray or mist to knock down vapors; vapor knock down water is corrosive or toxic and should be diked for containment. Stop leak if you can do so without risk. Do not touch spilled material. Keep combustibles (wood, paper, oil, etc.) away from spilled material. Clean up only under supervision of an expert. Keep this chemical out of a confined space, such as a sewer, because of the possibility of an explosion, unless the sewer is designed to prevent the build-up of explosive concentrations. It may be necessary to contain and dispose of this chemical as a hazardous waste. If material or contaminated runoff enters water ways, notify downstream users of potentially contaminated waters (Sittig 2002).

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## Potassium Chloride 7447-40-7

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

Potassium Chloride (KCl) is ubiquitous in the environment, occurring in minerals, soil and sediments, and natural waters (oceans, lakes, rivers). KCl as inorganic salt is not subjected to further degradation processes in the environment.

Water: In water, potassium chloride is highly water soluble, and readily undergoes dissociation. In soil, transport/leaching of potassium and chloride is affected by clay minerals (type and content), pH, and organic matter (Betlach et al.1987, Biesinger and Christensen 1972). K is in general less mobile and less prone to leaching than anions in soil, such as chloride and nitrate ( $\text{NO}_3^-$ ). Chloride binds only weakly to soil particles, and therefore follows water movement. Anions do not leach alone, but always together with a counter-ion (cation). Chloride and nitrate leaching may deplete soils of nutrients such as K. Loss of potassium by erosion or leaching carries no environmental concern. In one published study carried out under national guidelines and with reliable procedure description, the following results on acute aquatic toxicity were found: *Pimephales promelas*: 96 h-LC<sub>50</sub> = 880 mg/l ( $\pm 15$ ); *Daphnia magna*: 48h EC<sub>50</sub> = 177 mg/l; *Ceriodaphnia dubia*: 48 h-EC<sub>50</sub> = 630 mg/l ( $\pm 14$ ) (Lewis 1999). For fish species, these studies show effects in the similar range of doses: two studies on *Lepomis macrochirus* (Lyman 1983, 1983), one study on *Oncorhynchus mykiss* and one study on *Ictalurus punctatus* showed 48 h-LC<sub>50</sub> values in the range 720-2010 mg/l (Marschner1995). The toxicity of KCl has been investigated in one algae species (*Nitzschia linearis*), showing 120 h-EC<sub>50</sub> (growth rate) of 1337 mg/l (Lyman 1983).

Soil: Agricultural soil productivity is dependent on establishing a balanced availability of potassium throughout the growing season, often resulting in the need of fertilizing with potassium. Inadequate maintenance of K in agricultural soils will lead to loss of soil productivity. K deficiencies may also restrict a crop's ability to utilize N, resulting in increased potential for nitrate leaching, and lack of K can enhance a crop's susceptibility to disease and increase the need for pesticide application.

Plants: Toxicity to terrestrial plants: Potassium is one of the three major nutrients and chloride is an essential micronutrient for plants. Literature search has not revealed studies related to toxic effects on terrestrial organisms, therefore this section only describes levels of potassium necessary for good growth of plants and its role in plant physiology (Betlach et al.1987, Biesinger and Christensen 1972, Mitchell et al. 1988).

Demands for potassium and chloride in plants/crops: The potassium requirement for optimal plant growth is in the range 2-5 % of the plant dry weight of vegetative parts. In crops, uptake of K in plants ranges from about 50 to 300 kg K ha<sup>-1</sup> per crop, broadly similar to the uptake of nitrogen. Inadequate supply of potassium makes plants more susceptible to frost damage, fungal attack and drought. In most plant species the Cl requirement for optimal growth is in the range of 0.2-0.4 mg/g dry matter. Excessive Cl is detrimental to chloride-sensitive crops, or if it is necessary to guard against excessive salinity, and in arid areas Cl deficiency may occur in highly leached soils with a low Cl input from rain. The principal effect of Cl deficiency in plants is a reduction in leaf surface area and thereby plants dry weight. With severe deficiency, necrosis might occur.

Environmental Partitioning: Environmental distribution of potassium chloride using a Generic Level I Fugacity Model under three emission scenarios showed that KCl is mainly distributed to water (99.7 %), while other compartments make minor contributions (< 0.03 % to air, < 0.03 % to soil solids, < 0.01 % to sediment solids). This result applied under the three emissions scenarios 100 % release to water, 100 % release to air, or 100 % release to soil [12]. Fugacity estimations may not be reliable for KCl. The modelling programs used are designed for organic chemicals and the chemical is not expected to undergo photolysis or biodegradation. Taking into consideration the physico-chemical properties of potassium chloride, and a calculated octanol/water partition coefficient of – 0.46, no potential for bioaccumulation/ bioconcentration can be identified.

## Regulatory

MCL (Drinking Water): none

Universal Treatment Standards (UTS: none in wastewater standards

NIOSH: = 1500 mg/kg Oral LD<sub>50</sub> Mouse = 2600 mg/kg Oral LD<sub>50</sub> Rat

EPA RfD: none

EPA RfC: none

## Remediation

KCl does not pose a threat.

Spills/Leaks: Clean up spills immediately, observing precautions in the Protective Equipment section. Sweep up or absorb material, then place into a suitable clean, dry, closed container for disposal. Avoid generating dusty conditions. Provide ventilation.

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## Potassium Cyanide 151-50-8

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Most cyanide in the atmosphere exists almost entirely as hydrogen cyanide gas, although small amounts of metal cyanides may be present as particulate matter in the air (EPA 1984). Hydrogen cyanide is very resistant to photolysis at wavelengths of normal sunlight (EPA 1979). The most important reaction of hydrogen cyanide in air is the reaction with photochemically-generated hydroxyl radicals and subsequent rapid oxidation to carbon monoxide (CO) and nitric oxide (NO); photolysis and reaction with ozone are not important transformation processes, and reaction with singlet oxygen ( $O^1D$ ) is not a significant transformation process except at stratospheric altitudes where singlet oxygen is present in significant concentrations (Cicerone and Zellner 1983). The rate of hydroxyl radical reaction with hydrogen cyanide in the atmosphere depends on the altitude, and the rate of the reaction is at least an order of magnitude faster at lower tropospheric altitudes (0–8 km) than at upper tropospheric altitudes (10–12 km) (Cicerone and Zellner 1983). Based on a reaction rate constant of  $3 \times 10^{-14}$  cm<sup>3</sup>/(molecule-sec) at 25 °C (Fritz et al. 1982) and assuming an average hydroxyl radical concentration of  $5 \times 10^5$  molecules/cm<sup>3</sup>, the residence time for the reaction of hydrogen cyanide vapor with hydroxyl radicals in the atmosphere is  $\approx 2$  years. This value compares well with the atmospheric residence time derived by Cicerone and Zellner (1983) of approximately 2.5 years, with a range of 1.3–5.0 years, depending on the hydroxyl radical concentrations assumed. Using the equation  $t_{1/2} = 0.693\tau$  for converting residence time ( $\tau$ ) to half-life ( $t_{1/2}$ ) (Lyman 1982) and an estimated atmospheric residence time for hydrogen cyanide of 2–3 years, and assuming first-order kinetics for the reaction of hydrogen cyanide with hydroxyl radicals, an atmospheric half-life of 1.4–2.9 years can be calculated for hydrogen cyanide.

**Water:** Cyanide occurs most commonly as hydrogen cyanide in water, although it can also occur as the cyanide ion, alkali and alkaline earth metal cyanides (potassium cyanide, sodium cyanide, calcium cyanide), relatively stable metalocyanide complexes (ferricyanide complex  $[Fe(CN)_6]^{-3}$ ), moderately stable metalocyanide complexes (complex nickel and copper cyanide), or easily decomposable metalocyanide complexes (zinc cyanide  $[Zn(CN)_2]$ , cadmium cyanide  $[Cd(CN)_2]$ ). The environmental fate of these cyanide compounds varies widely (EPA 1979).

**Biota:** Most of the available information on the mechanisms of biodegradation of cyanides in water comes from studies on the evaluation and use of this process as a means of detoxifying cyanide-containing wastes (Akcil and Mudder 2003; EPA 1994; Raybuck 1992). It is known that there is a natural attenuation of the cyanide ion and thiocyanide concentrations in waste waters, for example reduction in gold mill tails that is due the acclimation of indigenous microflora in the tailings (Akcil and Mudder 2003; Oudjehani et al. 2002; Zagury et al. 2004). A number of microorganisms have been identified that are capable of uptake, conversion, sorption, and/or precipitation of the cyanide ion, cyanate, and thiocyanate, including species of the

genera, *Actinomyces*, *Alcaligenes*, *Arthrobacter*, *Bacillus*, *Micrococcus*, *Neisseria*, *Paracoccus*, *Pseudomonas*, and *Thiobacillus* (Akcil and Mudder 2003). Some of these species, for example *Pseudomonas*, are capable of using the cyanide ion and thiocyanate as the sole source of carbon and nitrogen and therefore, are particularly effective at cyanide degradation. In fact, *Pseudomonas* is the basis of commercial applications for degrading the cyanide ion to ammonia and carbonate in waste waters generated in mining operations that use the cyanide ion to leach gold and other precious metals for low-grade ores (Akcil and Mudder 2003).

**Sediment and Soil:** Analogous to the fate of cyanides in water, it is predicted that the fate of cyanides in soil would be dependent on cyanide concentrations, pH, temperature, metal content, concentration of microbes, availability of nutrients, and acclimation of microbes. Cyanide may occur as hydrogen cyanide, alkali metal salts, or as immobile metalocyanide complexes. In soil, cyanide present at low concentrations would biodegrade under aerobic conditions with the initial formation of ammonia, which would be converted to nitrite and nitrate in the presence of nitrifying bacteria. Under anaerobic conditions, the cyanides ion will denitrify to gaseous nitrogen (Richards and Shieh 1989). Upper limits of 200 and 2 ppm (mg/kg  $\text{CN}^-$ ), respectively, have been reported for uninhibited aerobic and anaerobic biodegradation of cyanide in soil (Fueller 1985); however, these limits have not been confirmed in other studies (Thomas and Lester 1993). Cyanide ions in soil are not involved in oxidation-reduction reactions but may undergo complexation reactions with metal ions in soil (EPA 1978).

**Environmental partitioning:** Because hydrogen cyanide is a gas and has a relatively slow degradation rate in air, the atmosphere will be the ultimate sink for this compound. Almost all of the hydrogen cyanide released to the atmosphere remains in the lower altitudes (troposphere); only 2% of tropospheric hydrogen cyanide is transferred to the stratosphere (Cicerone and Zellner 1983). Cyanide has the potential to be transported over long distances from its emission source. Despite higher water solubility at saturated pressure, the removal of hydrogen cyanide by rainwater appears to be a negligible partitioning pathway (Cicerone and Zellner 1983). Because hydrogen cyanide is a gas, its removal from air by dry deposition is also likely to be negligible. However, metal cyanide particles, particularly water-soluble cyanide particles, are expected to be removed from the air by both wet and dry deposition.

Volatilization and sorption are the two physical processes that contribute to the loss of cyanide from water. At pH <9.2, most of the free cyanide in solution should exist as hydrogen cyanide, a volatile cyanide form (EPA 1978). On the basis of Henry's law constant and the volatility characteristics associated with various ranges of Henry's law constant (Thomas 1982), volatilization is a significant and probably dominant fate process for hydrogen cyanide in surface water (EPA 1992). The most common alkali metal cyanides (e.g., sodium and potassium cyanide) may also be lost from surface water primarily through volatilization; whereas, the sparingly soluble metal cyanides such as copper (I) cyanide are removed from water predominantly by sedimentation and biodegradation (EPA 1992). Variations in the volatilization rate are expected because this process is affected by several parameters including temperature, pH, wind speed, and cyanide concentration (EPA 1979). The volatilization rate was pH-dependent, with the rate faster at a lower pH. Data indicated that cyanide volatilization is a more important fate process than cyanide loss due to chemical and biodegradation reactions (Ludzack et al. 1951; Raef et al. 1977). Because volatilization is not an important fate process for cyanide in groundwater, cyanide would be expected to persist for considerably longer periods of time in underground aquifers than in surface water.

## Regulatory

MCL (Drinking Water): 0.2 mg/L cyanide

Universal Treatment Standards (UTS): 1.2 mg/l<sup>2</sup> in wastewater standard

590 mg/L in nonwastewater standard

EPA RfD: 0.05 mg/kg-day

EPA RfC: none mg/m<sup>3</sup>

Soil PQL: 3 mg/kg

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/l
11	35	3.1	6.2

## Remediation

Biodegradation of cyanides has been investigated, with varying results, for several industrial processes, and additional research in this area would be valuable. While investigations of the potential for microbial species found in mineral processing waste waters demonstrate effective removal of cyanide, metal complexed cyanide, and thiocyanate (Boucabeille 1994; EPA 1994), complex cyanides did not appear amenable to biodegradation at gasworks sites (Thomas and Lester 1993). Application of formaldehyde to electroplating waste under basic conditions can convert the cyanide anion to substituted acetates in addition to recovering copper and silver as free metals with formaldehyde reduction (Tucker and Carson 1985). Calcium or sodium polysulfide treatment converts some cyanide wastes into less toxic thiocyanate (Higgins and Desher 1988). These examples suggest that typical treatments involve the decomposition of cyanides to less toxic compounds by physical or chemical processes. More than 97% of cyanide is typically removed from waste waters by alkaline chlorination, electrolysis, or ozonation process. Cyanide from some wastes can be removed by ion-exchange resins. After using an appropriate treatment method such as those described above, cyanide wastes may be disposed of in a secured sanitary landfill (Grosse 1986; Higgins and Desher 1988; Tucker and Carson 1985). Disposal by injection of high-pH cyanide wastes into sandstone was investigated by Scrivner et al. (1986). Currently, the injection of waste water containing hydrogen cyanide and cyanide compounds through underground injection is a major method for disposal of these wastes.

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## Propylene Oxide 75-56-9

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), 1,2-propylene oxide, which has a vapor pressure of 538 mm Hg at 25 deg C (Boublik et al. 1984) is expected to exist solely as a vapor in the ambient atmosphere. Vapor-phase 1,2-propylene oxide is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals(SRC); the half-life for this reaction in air is estimated to be 30 days (SRC), calculated from its rate constant of  $5.2 \times 10^{-12}$  cu cm/molecule-sec at 25 deg C.

**Water:** Based on a classification scheme (Swann et al. 1983), an estimated Koc value of 25, determined from an measured log Kow of 0.03 and a regression-derived equation (Lyman et al. 1994), indicates that 1,2-propylene oxide is not expected to adsorb to suspended solids and sediment (SRC). Volatilization from water surfaces is expected (Lyman et al. 1994) based upon an estimated Henry's Law constant of  $6.96 \times 10^{-5}$  atm-cu m/mole, based upon its vapor pressure, 538 mm Hg (Boublik et al. 1984), and water solubility,  $5.9 \times 10^5$  mg/l (Bogyo et al. 1980). Using this Henry's Law constant and an estimation method (Lyman et al. 1994), volatilization half-lives for a model river and model lake are 12 hours and 6 days, respectively (SRC). According to a classification scheme (Franke et al. 1994), an estimated BCF of 3 (SRC), from its log Kow (Hansch et al. 1995) and a regression- derived equation (Meylan et al. 1999), suggests the potential for bioconcentration in aquatic organisms is low (SRC). In freshwater, propylene oxide will hydrolyze with estimated half-lives of 11.6 days (pH's 7-9) and 6.6 days (pH 5) at 25 deg C (Bogyo et al. 1980). 1, 2-Propylene oxide, present at 100 mg/l, reached 95% of its theoretical BOD in 3 weeks using an activated sludge inoculum at 30 mg/l and the Japanese MITI test (CITI 1992), suggesting biodegradation will be an important fate process.

**Soil:** An estimated Koc value of 25 (SRC), determined from a measured log Kow of 0.03 (Hansch et al. 1995) and a regression-derived equation (Lyman et al. 1994), indicates that 1, 2-propylene oxide is expected to have very high mobility in soil (Swann et al. 1983). Volatilization of 1,2-propylene oxide from moist soil surfaces is expected to be an important fate process (SRC) given an estimated Henry's Law constant of  $6.96 \times 10^{-5}$  atm-cu m/mole (SRC), calculated from its vapor pressure, 538 mm Hg (Boublik et al. 1984), and water solubility,  $5.9 \times 10^5$  mg/l (Bogyo et al. 1980). The potential for volatilization of 1,2-propylene oxide from dry soil surfaces may exist (SRC) based upon a vapor pressure of 538 mm Hg (Boublik et al. 1984). The aqueous hydrolysis of propylene oxide occurs at

an environmentally important rate; therefore, hydrolysis in moist soil is likely to be important (Bogyo et al. 1980). 1,2-Propylene oxide, present at 100 mg/l, reached 95% of its theoretical BOD in 3 weeks using an activated sludge inoculum at 30 mg/l and the Japanese MITI test (CITI 1992), suggesting biodegradation will be an important fate process.

**Biota:** An estimated BCF of 3 was calculated for 1,2-propylene oxide(SRC), using a log Kow of 0.03 (Hansch et al. 1995) and a regression-derived equation (Meylan et al. 1994). According to a classification scheme (Franke et al. 1994), this BCF suggests the potential for bioconcentration in aquatic organisms is low.

## Regulatory

MCL (Drinking Water): none

Universal Treatment Standards (UTS): None

OPPTS RfD: 0.001 mg/kg-day

EPA RfC: 0.03 mg/m<sup>3</sup>

Drinking Water Remedial Target: 35 µg/L (see EPA 2005 c)

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
1.9	9.1	0.52	0.22

## Remediation

**Land spill:** Dig a pit, pond, lagoon, or holding area to contain liquid or solid material. If time permits, pits, ponds, lagoons, soak holes, or holding areas should be sealed with an impermeable flexible membrane liner. Dike surface flow using soil, sand bags, foamed polyurethane, or foamed concrete. Absorb bulk liquid with fly ash, cement powder, sawdust, or commercial sorbents. Apply "universal" gelling agent to immobilize spill.

**Water spill:** Use natural barriers or oil spill control booms to limit spill. Use surface active agent (eg detergent, soaps, alcohols) to compress and thicken spilled material. Inject "universal" gelling agent to solidify encircled spill and increase effectiveness of booms. If dissolved, apply activated carbon at ten times the spilled amount in region of 10 ppm or greater concentration. Use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates.

**Air release:** Apply water spray or mist to knock down vapors (Association of American Railroads 1994).

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## Sodium Arsenite 7784-46-5

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Arsenic is released into the atmosphere primarily as arsenic trioxide or, less frequently, in one of several volatile organic compounds, mainly arsines (EPA 1982b). Trivalent arsenic and methyl arsines in the atmosphere undergo oxidation to the pentavalent state (EPA 1984), and arsenic in the atmosphere is usually a mixture of the trivalent and pentavalent forms (EPA 1984; Scudlark and Church 1988). Photolysis is not considered an important fate process for arsenic compounds (EPA 1979).

**Water:** Arsenic in water can undergo a complex series of transformations, including oxidation-reduction reactions, ligand exchange, precipitation, and biotransformation (EPA 1979, 1984; Sanders et al. 1994; Welch et al. 1988). Rate constants for these various reactions are not readily available, but the factors most strongly influencing fate processes in water include Eh, pH, metal sulfide and sulfide ion concentrations, iron concentrations, temperature, salinity, distribution and composition of the biota, season, and the nature and concentration of natural organic matter (Farago 1997; Redman et al. 2002; Wakao et al. 1988). Organic arsenical pesticides, such as MSMA, DSMA, and DMA do not degrade by hydrolysis or by aquatic photolysis (EPA 2006). No formation of arsine gas from marine environments has been reported (Tamaki and Frankenberger 1992). Inorganic species of arsenic are predominant in the aquatic environment. In the pH range of natural waters, the predominant aqueous inorganic As(V) species are the arsenate ions,  $\text{H}_2\text{AsO}_4^-$  and  $\text{HAsO}_4^{2-}$ ; the predominant inorganic As(III) species is  $\text{As}(\text{OH})_3$  (Aurillo et al. 1994; EPA 1982b). As(V) generally dominates in oxidizing environments such as surface water and As(III) dominates under reducing conditions such as may occur in groundwater containing high levels of arsenic. However, the reduction of arsenate to arsenite is slow, so arsenate can be found in reducing environments. Conversely, the oxidation of arsenite in oxidizing environments is moderately slow (half-life, 0.4–7 days in coastal systems) and therefore, arsenite can be found in oxidizing environments (Mariner et al. 1996; Sanders et al. 1994). The main organic species in fresh water are MMA and DMA; however, these species are usually present at lower concentrations than inorganic arsenic species (Eisler 1994). Aquatic microorganisms may reduce the arsenate to arsenite, as well as methylate arsenate to its mono- or dimethylated forms (Aurillo et al. 1994; Benson 1989; Braman and Foreback 1973; Edmonds and Francesconi 1987; Sanders et al. 1994). Methylated species are also produced by the biogenic reduction of more complex organoarsenic compounds like arsenocholine or arsenobetaine.

**Sediment and Soil:** In soil, arsenic is found as a complex mixture of mineral phases, such as co-precipitated and sorbed species, as well as dissolved species (Roberts et al. 2007). The degree of arsenic solubility in soil will depend on the amount of arsenic distributed between these different mineral phases. The dissolution of arsenic is also affected by particle size. The distribution between these phases may reflect the arsenic source (e.g., pesticide application, wood treatment, tanning, or mining operations), and may change with weathering and associations with iron and manganese oxides and phosphate minerals in the soil (Roberts et al. 2007; Ruby et al. 1999). Davis et al. (1996) reported that in soil in Anaconda, Montana, a smelting site from 1860 to 1980, contained arsenic that is only in a sparingly soluble form, consisting of primarily arsenic oxides and phosphates.

The arsenic cycle in soils is complex, with many biotic and abiotic processes controlling its overall fate and environmental impact. Arsenic in soil exists in various oxidation states and chemical species, depending upon soil pH and oxidation-reduction potential. Under most environmental conditions, inorganic As (V) will exist as a mixture of arsenate anions,  $\text{H}_2\text{AsO}_4$  and  $\text{HAsO}_4^{2-}$ , and inorganic As (III) will exist as  $\text{H}_3\text{AsO}_3$ . The arsenate and arsenite oxyanions have various degrees of protonation depending upon pH (EPA 1982b; McGeehan 1996). As (V) predominates in aerobic soils, and As (III) predominates in slightly reduced soils (e.g., temporarily flooded) or sediments (EPA 1982b; Sanders et al. 1994). As (III) commonly partitions to the aqueous phase in anoxic environments, and would be more mobile. As (V) usually remains bound to minerals, such as ferrihydrite and alumina, limiting its mobility and bioavailability (Rhine et al. 2006).

Arsenite is moderately unstable in the presence of oxygen; however, it can be found under aerobic conditions as well (Sanders et al. 1994). While arsenate is strongly sorbed by soils under aerobic conditions, it is rapidly desorbed as the system becomes anaerobic. Once it is desorbed, arsenate can be reduced to arsenite, which exhibits greater mobility in soils (McGeehan 1996). Transformations between the various oxidation states and species of arsenic occur as a result of biotic or abiotic processes (Bhumbla and Keefer 1994). While degradation of an organic compound is typically considered complete mineralization, in the case of organic arsenic compounds, the element arsenic itself cannot be degraded. However, the organic portion of the molecule can be metabolized (Woolson 1976). Organic arsenical pesticides, such as MSMA, DSMA, and DMA, do not degrade by hydrolysis or by soil photolysis (EPA 2006).

**Environmental partitioning:** Arsenic in soil may be transported by wind or in runoff or may leach into the subsurface soil. However, because many arsenic compounds tend to partition to soil or sediment under oxidizing conditions, leaching usually does not transport arsenic to any great depth (EPA 1982a; Moore et al. 1988; Pantisar-Kallio and Manninen 1997; Welch et al. 1988). Arsenic is largely immobile in agricultural soils; therefore, it tends to concentrate and remain in upper soil layers indefinitely. Downward migration has been shown to be greater in a sandy soil than in a clay loam (Sanok et al. 1995). Arsenic from lead arsenate that was used for pest control did not migrate downward below 20 cm in one fruit orchard; in another orchard, 15 years after sludge amendments and deep plowing, essentially all arsenic residues remained in the upper 40 cm of soil (Merwin et al. 1994). Leaching of arsenic in polluted wetland soil was low; leaching was correlated with the amount of dissolved organic matter in the soil (Kalbitz and Wennrich 1998). The effect of soil characteristics, namely pH, organic matter content, clay content, iron oxide content, aluminum oxide content, and cation exchange capacity (CEC), on the adsorption of various metals, including the metalloid arsenic, to 20 Dutch surface soils was assessed by regression analysis (Janssen et al. 1997). The most influential parameter affecting arsenic adsorption was the iron content of the soil.

Transport and partitioning of arsenic in water depends upon the chemical form (oxidation state and counter ion) of the arsenic and on interactions with other materials present. Soluble forms move with the water, and may be carried long distances through rivers (EPA 1979). However, arsenic may be adsorbed from water onto sediments or soils,

especially clays, iron oxides, aluminum hydroxides, manganese compounds, and organic material (EPA 1979, 1982a; Welch et al. 1988). Under oxidizing and mildly reducing conditions, groundwater arsenic concentrations are usually controlled by adsorption rather than by mineral precipitation. The extent of arsenic adsorption under equilibrium conditions is characterized by the distribution coefficient,  $K_d$ , which measures the equilibrium partitioning ratio of adsorbed to dissolved contaminant. The value of  $K_d$  depends strongly upon the pH of the water, the arsenic oxidation state, and the temperature. In acidic and neutral waters, As(V) is extensively adsorbed, while As (III) is relatively weakly adsorbed. Trivalent inorganic arsenic exists predominantly as arsenous acid ( $H_3AsO_3$ ) at environmental pH and is not strongly adsorbed to suspended solids and sediments in the water column.

Pentavalent inorganic arsenic exists predominantly as  $H_2AsO_4$  and  $HAsO_4^{2-}$  in most environmental waters, which has considerably greater adsorption characteristics than arsenous acid. While in acidic and neutral waters, As(V) is more strongly adsorbed relative to As(III), in high-pH waters (pH >9) aquifer  $K_d$  values are considerably lower for both oxidation states (Mariner et al. 1996). Sediment-bound arsenic may be released back into the water by chemical or biological interconversions of arsenic species. Arsenic enters rivers from where mining operations occurred and is transported downstream, moving from water and sediment into biofilm (attached algae, bacterial, and associated fine detrital material), and then into invertebrates and fish. The source of arsenic in the water column may be resuspended sediment. While arsenic bioaccumulates in animals, it does not appear to biomagnify between trophic levels (Eisler 1994; Farag et al. 1998; Williams et al. 2006).

Terrestrial plants may accumulate arsenic by root uptake from the soil or by absorption of airborne arsenic deposited on the leaves, and certain species may accumulate substantial levels (EPA 1982b). Yet, even when grown on highly polluted soil or soil naturally high in arsenic, the arsenic level taken up by the plants is comparatively low (Gebel et al. 1998b; Pitten et al. 1999). Kale, lettuce, carrots, and potatoes were grown in experimental plots surrounding a wood preservation factory in Denmark where waste wood was incinerated to investigate the amount and pathways for arsenic uptake by plants (Larsen et al. 1992). On incineration, the arsenate in the wood preservative was partially converted to arsenite; the arsenic emitted from the stack was primarily particle bound. Elevated levels of inorganic arsenic were found in the test plants and in the soil around the factory

## Regulatory

MCL (Drinking Water): 0.01 mg/L

Universal Treatment Standards (UTS): 1.4mg/l<sup>2</sup> in wastewater standards (Arsenic)

5.0 mg/L TCLP nonwastewater standard (Arsenic)

EPA Rf D: 0.0003 mg/kg-day

EPA RfC: None

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
0.39	2.7	0.00045	0.045

## Remediation

The practice of liming to remediate contaminated soils and mine tailings has the potential to mobilize arsenic. Experiments performed by Jones et al. (1997) indicate that the increased mobility appears to be consistent with the pH dependence of sorption reactions of arsenic on iron oxide minerals rather than dissolution-precipitation reactions involving arsenic. They recommend that remediation of acidic mine tailings or other arsenic-contaminated soils be carefully evaluated with respect to potential arsenic mobilization, especially at contaminated sites hydraulically connected to surface or groundwater.

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## Sodium Cyanide 143-33-9 (Cyanide)

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Most cyanide in the atmosphere exists almost entirely as hydrogen cyanide gas, although small amounts of metal cyanides may be present as particulate matter in the air (EPA 1984). Hydrogen cyanide is very resistant to photolysis at wavelengths of normal sunlight (EPA 1979). The most important reaction of hydrogen cyanide in air is the reaction with photochemically-generated hydroxyl radicals and subsequent rapid oxidation to carbon monoxide (CO) and nitric oxide (NO); photolysis and reaction with ozone are not important transformation processes, and reaction with singlet oxygen ( $O^1D$ ) is not a significant transformation process except at stratospheric altitudes where singlet oxygen is present in significant concentrations (Cicerone and Zellner 1983). The rate of hydroxyl radical reaction with hydrogen cyanide in the atmosphere depends on the altitude, and the rate of the reaction is at least an order of magnitude faster at lower tropospheric altitudes (0–8 km) than at upper tropospheric altitudes (10–12 km) (Cicerone and Zellner 1983). Based on a reaction rate constant of  $3 \times 10^{-14}$  cm<sup>3</sup>/(molecule-sec) at 25 °C (Fritz et al. 1982) and assuming an average hydroxyl radical concentration of  $5 \times 10^5$  molecules/cm<sup>3</sup>, the residence time for the reaction of hydrogen cyanide vapor with hydroxyl radicals in the atmosphere is  $\approx 2$  years. This value compares well with the atmospheric residence time derived by Cicerone and Zellner (1983) of approximately 2.5 years, with a range of 1.3–5.0 years, depending on the hydroxyl radical concentrations assumed. Using the equation  $t_{1/2} = 0.693\tau$  for converting residence time ( $\tau$ ) to half-life ( $t_{1/2}$ ) (Lyman 1982) and an estimated atmospheric residence time for hydrogen cyanide of 2–3 years, and assuming first-order kinetics for the reaction of hydrogen cyanide with hydroxyl radicals, an atmospheric half-life of 1.4–2.9 years can be calculated for hydrogen cyanide.

**Water:** Cyanide occurs most commonly as hydrogen cyanide in water, although it can also occur as the cyanide ion, alkali and alkaline earth metal cyanides (potassium cyanide, sodium cyanide, calcium cyanide), relatively stable metalcyanide complexes (ferricyanide complex  $[Fe(CN)_6]^{-3}$ ), moderately stable metalcyanide complexes (complex nickel and copper cyanide), or easily decomposable metalcyanide complexes (zinc cyanide  $[Zn(CN)_2]$ , cadmium cyanide  $[Cd(CN)_2]$ ). The environmental fate of these cyanide compounds varies widely (EPA 1979).

In water, hydrogen cyanide and cyanide ion exist in equilibrium with their relative concentrations primarily dependent on pH and temperature. At pH < 8, >93% of the free cyanide in water will exist as undissociated hydrogen cyanide (EPA 1978). Hydrogen cyanide can be hydrolyzed to formamide, which is subsequently hydrolyzed to ammonium and formate ions (EPA 1979). However, the relatively slow rates of hydrolysis reported for hydrogen cyanide in acidic solution (Kriebel and McNally 1929; Kriebel and Peiker 1933) and of cyanides under alkaline conditions (Wiegand and Tremelling 1972) indicate that hydrolysis is not competitive with volatilization and biodegradation for removal of free cyanide from ambient waters (EPA 1979).

**Sediment and Soil:** Analogous to the fate of cyanides in water, it is predicted that the fate of cyanides in soil would be dependent on cyanide concentrations, pH, temperature, metal content, concentration of microbes, availability of nutrients, and acclimation of microbes. Cyanide may occur as hydrogen cyanide, alkali metal salts, or as immobile metalcyanide complexes. In soil, cyanide present at low concentrations would biodegrade under aerobic conditions with the initial formation of ammonia, which would be converted to nitrite and nitrate in the presence of nitrifying bacteria. Under anaerobic conditions, the cyanides ion will denitrify to gaseous nitrogen (Richards and Shieh 1989). Upper limits of 200 and 2 ppm (mg/kg CN<sup>-</sup>), respectively, have been reported for uninhibited aerobic and anaerobic biodegradation of cyanide in soil (Fueller 1985); however, these limits have not been confirmed in other studies (Thomas and Lester 1993). Cyanide ions in

soil are not involved in oxidation-reduction reactions but may undergo complexation reactions with metal ions in soil (EPA 1978).

**Environmental partitioning:** Because hydrogen cyanide is a gas and has a relatively slow degradation rate in air, the atmosphere will be the ultimate sink for this compound. Almost all of the hydrogen cyanide released to the atmosphere remains in the lower altitudes (troposphere); only 2% of tropospheric hydrogen cyanide is transferred to the stratosphere (Cicerone and Zellner 1983). Cyanide has the potential to be transported over long distances from its emission source. Despite higher water solubility at saturated pressure, the removal of hydrogen cyanide by rainwater appears to be a negligible partitioning pathway (Cicerone and Zellner 1983). Because hydrogen cyanide is a gas, its removal from air by dry deposition is also likely to be negligible. However, metal cyanide particles, particularly water-soluble cyanide particles, are expected to be removed from the air by both wet and dry deposition.

Volatilization and sorption are the two physical processes that contribute to the loss of cyanide from water. At pH <9.2, most of the free cyanide in solution should exist as hydrogen cyanide, a volatile cyanide form (EPA 1978). On the basis of Henry's law constant and the volatility characteristics associated with various ranges of Henry's law constant (Thomas 1982), volatilization is a significant and probably dominant fate process for hydrogen cyanide in surface water (EPA 1992). The most common alkali metal cyanides (e.g., sodium and potassium cyanide) may also be lost from surface water primarily through volatilization; whereas, the sparingly soluble metal cyanides such as copper (I) cyanide are removed from water predominantly by sedimentation and biodegradation (EPA 1992). Variations in the volatilization rate are expected because this process is affected by several parameters including temperature, pH, wind speed, and cyanide concentration (EPA 1979). EPA (1979) summarized results of a laboratory study that indicated that the volatilization half-life of hydrogen cyanide from solutions at concentrations of 25–200 µg/L ranged from 22 to 110 hours. First-order kinetics were observed. In outdoor experiments with moderate winds, the rate of hydrogen cyanide loss increased by a factor of 2–2.5. In a study to evaluate the effect of cyanide on biochemical oxidation, there was a 50% loss of 6 ppm (mg/L) cyanide in river water kept in open biochemical oxygen demand bottles (without aeration) at pH 7.4 within ≈10 days (Ludzack et al. 1951). When the bottles were aerated (rate of aeration not given), 50% loss occurred in only ≈10 hours. The kinetics of the rate of loss due to volatilization were not rigorously investigated. The volatilization rate was pH-dependent, with the rate faster at a lower pH. Data indicated that cyanide volatilization is a more important fate process than cyanide loss due to chemical and biodegradation reactions (Ludzack et al. 1951; Raef et al. 1977a). Because volatilization is not an important fate process for cyanide in groundwater, cyanide would be expected to persist for considerably longer periods of time in underground aquifers than in surface water.

## Regulatory

Canada Drinking Water Standards and Criteria (MAC): 200 ug/L

Canadian Water Quality Guidelines for the Protection of Aquatic Life: 5 ug/L

MCL (Drinking Water): 0.2 mg/L (CN)

Universal Treatment Standards (UTS) for CN totals: 1.2 mg/l<sup>2</sup> in wastewater standards

590 mg/L in nonwastewater standards

EPA RfD: 0.04 mg/kg-day

EPA RfC: 3.0 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
11	35	3.1	6.2

## Remediation

The biodegradation of cyanides has been investigated, with varying results, for several industrial processes, and additional research in this area would be valuable. While investigations of the potential for microbial species found in mineral processing waste waters demonstrate effective removal of cyanide, metal complexed cyanide, and thiocyanate (Boucabeille 1994; EPA 1994), complex cyanides did not appear amenable to biodegradation at gasworks sites (Thomas and Lester 1993). Application of formaldehyde to electroplating waste under basic conditions can convert the cyanide anion to substituted acetates in addition to recovering copper and silver as free metals with formaldehyde reduction (Tucker and Carson 1985). Calcium or sodium polysulfide treatment converts some cyanide wastes into less toxic thiocyanate (Higgins and Desher 1988). These examples suggest that typical treatments involve the decomposition of cyanides to less toxic compounds by physical or chemical processes. More than 97% of cyanide is typically removed from waste waters by alkaline chlorination, electrolysis, or ozonation process. Cyanide from some wastes can be removed by ion-exchange resins. After using an appropriate treatment method such as those described above, cyanide wastes may be disposed of in a secured sanitary landfill (Grosse 1986; Higgins and Desher 1988; Tucker and Carson 1985). Disposal by injection of high-pH cyanide wastes into sandstone was investigated by Scrivner et al. (1986). Currently, the injection of waste water containing hydrogen cyanide and cyanide compounds through underground injection is a major method for disposal of these wastes.

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## **Sulfur Dioxide 7446-09-5**

### **ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION**

**Air:** Wet deposition is, in general, much more easily measured than is dry deposition. Routine measurement of wet deposition is determined from sulfate concentration in precipitation samples and precipitation amount. Typically, the removal rate for particulate sulfate is of the order of 40% per hr, and for sulfur dioxide, an order of magnitude less. The overall efficiency of wet removal depends on many factors: precipitation type, intensity, duration, frequency, the relative amounts of sulfur dioxide and sulfate present, and the size distribution of particulate sulfate

Direct surface uptake of sulfur dioxide is the most important dry removal process for atmospheric sulfur; good sinks /include/ oceans (pH= 8), other non acidic moist surfaces, and some crops and forest species at certain growth stages; where as dry, snow covered surfaces and soils, for example, are less efficient (Monitoring and Assessment Research Centre 1978).

**Water:** Although snow covered surfaces are inefficient receptors of gaseous and particulate sulfur compound, the spring melt of the accumulated winter snowpack can result in rapid, short term inputs of high sulfate, low pH water to freshwater systems with resulting disastrous effects on fish (Monitoring & Assessment Research Centre 1978).

**Environmental Partitioning:** When released into the environment, sulfur dioxide moves into the air. In the air, it can be converted to sulfuric acid, sulfur trioxide, and sulfates. Sulfur dioxide dissolves in water. Once dissolved in water, sulfur dioxide can form sulfurous acid. Sulfur dioxide can be absorbed into the soil, but we don't know if or how it moves in soil (ATSDR 1998).

### **Regulations**

IDLH: 100 ppm

Acute MRL: 10ppm

MCL (Drinking Water): None

Universal Treatment Standards (UTS): none

Groundwater Water Quality Criterion : none ug/L

EPA Rf D: none mg/kg-day

EPA RfC: none mg/m<sup>3</sup>

## Remediation

Ventilate area of leak to disperse gas. If sulfur dioxide is in gaseous form, stop the flow of gas. Remove to safe place in open air and allow emptying. If in the liquid form, allow to vaporize (Mackison et al. 1981).

Approach release from upwind. Stop or control the leak, if this can be done without undue risk. Use water spray to cool and disperse vapors and protect personnel. Control runoff and isolate discharged material for proper disposal. Releases may require isolation or evacuation (Fire Protection Guide 1997).

## References

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Monitoring and Assessment Research Centre; Report #7 p.13-14, 21 (1978).

## Sulfur Trioxide 7446-11-9

### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** Sulfuric acid is formed in the atmosphere from sulfur dioxide. Both gaseous and aqueous phase reactions can occur, with gas phase reactions accounting for only about 5% of the sulfur dioxide oxidized during the summer (Pienaar and Helas 1996). In the process of sulfuric acid formation, sulfur trioxide is formed. After sulfur trioxide is formed, it rapidly reacts with water vapor to form sulfuric acid, so that processes that form sulfur trioxide in moist atmospheres are equivalent to the formation of sulfuric acid (Pienaar and Helas 1996). In the gaseous phase, substances that react with sulfur dioxide resulting in the production of sulfuric acid include O, HO•, HO<sub>2</sub>•, and CH<sub>2</sub>O<sub>3</sub>•. Substances that result in the oxidation of sulfur dioxide in the aqueous phase include ozone and hydrogen peroxide. Metal ions including ClO<sub>4</sub><sup>2-</sup>, VO<sub>2</sub><sup>+</sup>, Fe<sup>2+</sup>, Fe<sup>3+</sup>, Mn<sup>2+</sup>, and Ni<sup>2+</sup> could also directly oxidize SO<sub>3</sub><sup>2-</sup>, or catalyze SO<sub>3</sub><sup>2-</sup> oxidation by molecular oxygen. In a study using electron spin resonance, spin trapping, and high-performance liquid chromatography, Shi (1994) found that reactions of SO<sub>3</sub><sup>2-</sup> with NO<sub>2</sub><sup>-</sup> also generate SO<sub>3</sub>• radicals. The major reactions that form sulfuric acid from sulfur dioxide require light. Therefore, levels of sulfuric acid in the atmosphere show both seasonal and diurnal variations.

The production of sulfuric acid from sulfur dioxide can become limited if the concentrations of oxidants are less than the concentrations of sulfur dioxide. Oxidant limitations of the formation of sulfuric acid were shown in a study of sulfur dioxide and hydrogen peroxide concentrations over Columbus, Ohio; the study authors suggested that similar conditions exist over a large area of the eastern United States (Kleinman and Daum 1991). A deficiency of hydrogen peroxide relative to sulfur dioxide was also shown in a study completed at Whiteface Mountain, NY (Dutkiewicz et al. 1995). These results suggest that there would be a nonlinear relationship between sulfur dioxide emissions and downwind acid precipitation in the northeastern United States where aqueous-phase oxidation by hydrogen peroxide is the principal mechanism for forming sulfuric acid. The aqueous oxidation of sulfur dioxide to sulfuric acid in precipitation has been estimated to account for 50-80% of the sulfuric acid found in precipitation (Fung et al. 1991).

**Water:** Sulfuric acid is very corrosive and is highly toxic to aquatic and terrestrial life at low concentrations. In water, sulfuric acid dissociates. The sulfate anion may associate with other cations including calcium, magnesium, and aluminum. Sulfur that is in water may be oxidized to sulfuric acid by sulfur bacteria (Thiobacilli) that use sulfur to obtain energy for growth (Takeuchi and Suzuki 1994). Sulfate in water can also be reduced.

Because sulfur dioxide and sulfate are transformed through similar pathways in water, the effect of sulfur on aquatic systems is not dependent on the chemical or physical form of deposition (wet or dry) (EPA 1985). In the surface layer of the ocean the sulfate anion may be formed from dissolved sulfur dioxide, which is transformed to sulfurous acid ( $\text{H}_2\text{SO}_3$ ) and subsequently oxidized. Because of the relative lack of salt in freshwater, the oxidation of sulfur dioxide to sulfate is less likely to occur in freshwater. In the depths of the ocean, sulfate is reduced to sulfur dioxide, sulfur, and hydrogen sulfide by bacteria (Kellogg et al. 1972).

**Sediment and Soil:** When released into the soil, this material may leach into groundwater. When released into the air, this material may be removed from the atmosphere to a moderate extent by wet deposition. When released into the air, this material may be removed from the atmosphere to a moderate extent by dry deposition. The ions (sulfate, hydrogen) can adsorb to soil particles or be converted to gases (EPA 1985). Anaerobic bacteria in sediments and soil can reduce sulfate to sulfur and hydrogen sulfide (Kellogg et al. 1972).

**Transport and Partitioning:** Factors that affect the dispersion of sulfur compounds and other air pollutants include the temperature and efflux velocity of the emissions, stack height, topography and the proximity of other buildings, and meteorology. Meteorological conditions that result in temperature inversions can result in the trapping of pollutants, which increases concentrations up to several hundred times the concentrations usually found (WHO 1979). Sulfates, including sulfuric acid, are removed from the air by both dry and wet deposition processes. Wet deposition processes including rainout (a process that occurs within the clouds) and washout (removal by precipitation below the clouds) contribute to the removal of sulfate from the atmosphere (Kellogg et al. 1972). In the stratosphere, sulfuric acid aerosols have lifetimes of about 14 and 2.4 days at altitudes of 15 and 20 km, respectively (Kellogg et al. 1972). At cloud level, the residence time is about 6 days, with shorter residence times in surface air.

## **Regulatory-Sulfuric Acid**

MCL (Drinking Water): None

IDLH: 15 mg/m<sup>3</sup>

Universal Treatment Standards (UTS): None

MRL: None

EPA Rf D: None

EPA RfC: None

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m	tap water ug/L
None	None	None	None

## Remediation

Spent sulfuric acid can often be reprocessed to obtain a product of virgin quality. Enormous amounts of spent sulfuric acid are reprocessed since most of the sulfuric acid used for industrial processes acts only as a reagent and does not form part of the final product; one exception is the fertilizer industry (IARC 1992). It has been suggested that waste sulfuric acid can also be recycled using sulfate-reducing bacteria to produce hydrogen sulfide (Stucki et al. 1993).

It is not recommended that sulfuric acid or sulfur trioxide be placed in a landfill. Environmental regulatory agencies should be consulted for acceptable disposal practices (HSDB 1998). Sulfuric acid has been disposed of by being placed in sealed containers and by being absorbed in vermiculite, dry sand, or earth. Sulfuric acid may also be diluted and then neutralized. One method of neutralization is to add the acid slowly to a solution of soda ash and slaked lime, and to then flush with a large volume of water. Once sulfuric acid is diluted, and neutralized it can be discharged to a sewer. When diluting, the acid should always be added to a large volume of water because the heat released when a small bolus of water is added can cause the water to turn to steam, and the resulting effervescence can splatter the acid.

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### Trimethylamine 75-50-3

#### ENVIRONMENTAL TRANSPORT, DISTRIBUTION, AND TRANSFORMATION

**Air:** According to a model of gas/particle partitioning of semivolatile organic compounds in the atmosphere (Bidleman 1988), trimethylamine, which has a vapor pressure of  $1.6 \times 10^3$  mm Hg at 25 deg C (Daubert and Danner 1989), is expected to exist solely as a gas in the ambient atmosphere. Gas-phase trimethylamine is degraded in the atmosphere by reaction with photochemically-produced hydroxyl radicals (SRC); the half-life for this reaction in air is estimated to be 9 hrs (SRC), calculated from its rate constant of  $6.1 \times 10^{-11}$  cu cm/molecule-sec at 25 deg C (Atkinson 1989). Trimethylamine does not contain chromophores that absorb at wavelengths  $>290$  nm and therefore is not expected to undergo direct photolysis by sunlight (Lyman et al 1990).

**Water:** Based on a classification scheme (Swann et al 1983), an estimated Koc value of 29, determined from a measured log Kow of 0.16 (Hansch et al 1995) and a regression-derived equation (Lyman et al 1990), indicates that trimethylamine is not expected to adsorb to suspended solids and sediment (SRC). However, trimethylamine has a pKa of 9.8 (Perinn 1972) and should exist primarily as a cation under environmental conditions (pH 5-9) (SRC). As a result, trimethylamine may have greater adsorption to suspended solids and sediment than its estimated Koc value indicates (SRC). Volatilization of trimethylamine from water surfaces will not be an important fate process, since cations do not volatilize (SRC). According to a classification scheme, an estimated BCF of 3 (SRC), from its log Kow (Hansch et al 1995) and a regression-derived equation (Meylan et al 1999), suggests the potential for bioconcentration in aquatic organisms is low (SRC). Trimethylamine incubated in a marine sediment (Lowes Cove, ME) slurry underwent about 35% removal in 12 hrs, as measured by production of CO<sub>2</sub> and CH<sub>4</sub> (King et al 1983). Degradation products formed under aerobic conditions include dimethylamine, formaldehyde, formate, and CO<sub>2</sub> (Tate and Alexander 1976), while products formed under anaerobic conditions include dimethylamine, NH<sub>4</sub><sup>+</sup>, and CH<sub>4</sub> (Meiberg and Harber 1978). Trimethylamine achieved 66-92% of its theoretical BOD in 2 weeks using an activated sludge inoculum in the Japanese MITI test (Chemicals Evaluation Research Institute).

**Soil/Sediment:** Based on a classification scheme (Swann et al 1983), an estimated Koc value of 29 (SRC) determined from a measured log Kow of 0.16 (2) and a regression-derived equation (Lyman et al 1990), indicates that trimethylamine is expected to have very high mobility in soil (SRC). However, trimethylamine has a pKa of 9.8 (Perinn 1972) and should exist primarily as a cation under environmental conditions (pH 5-9)

(SRC). As a result, trimethylamine may have greater adsorption and less mobility than its estimated Koc value indicates since cations generally adsorb more strongly to soils containing organic carbon and clay than neutral species (Doucette 2000). Sorption coefficients for trimethylamine adsorption on montmorillonite, kaolinite and Flax Pond sediment (7% clay, 2.8% OM; Long Island, NY) were 15, 2 and 7 ml/g, respectively (Wang and Lee 1993). The trimethylamine cation adsorbed strongest to the negatively-charged montmorillonite via electrostatic interactions (Wang and Lee 1993). Volatilization of trimethylamine from moist soil surfaces will not be an important fate process because cationic trimethylamine does not volatilize. The potential for volatilization of trimethylamine from dry soil surfaces may exist (SRC) based upon a vapor pressure of 1,610 mm Hg at 25 deg C (Daubert and Danner 1989). Microbial production of dimethylamine from trimethylamine in soil was found to be greater under acidic conditions than at near neutral pH and greater under aerobic conditions than anaerobic conditions (Tate and Alexander 1976). Degradation products formed under aerobic conditions include dimethylamine, formaldehyde, formate, and CO<sub>2</sub> (Meiberg and Harber 1978), while products formed under anaerobic conditions include dimethylamine, NH<sub>4</sub><sup>+</sup>, and CH<sub>4</sub> (Hippe et al 1977).

**Biota:** An estimated BCF of 3 was calculated for trimethylamine (SRC), using a log Kow of 0.16 (Hansch et al 1995) and a regression-derived equation (Meylan et al 1999). According to a classification scheme (Franke et al 1994), this BCF suggests the potential for bioconcentration in aquatic organisms is low.

## Regulations

MCL (Drinking Water): none

Universal Treatment Standards (UTS): None

EPA RfD: none mg/kg-day

EPA RfC: 0.007 mg/m<sup>3</sup>

PRG: residential soil mg/kg	industrial soil mg/kg	ambient air ug/m <sup>3</sup>	tap water ug/L
23	88	7.3	12

## Remediation

Water spill: Add sodium bisulfate (NaHSO<sub>4</sub>). If dissolved, in region of 10 ppm or greater concentration, apply activated carbon at ten times the spilled amount. Use mechanical dredges or lifts to remove immobilized masses of pollutants and precipitates.

(Association of American Railroads).

Air release: Apply water spray or mist to knock down vapors. Vapor knockdown water is corrosive or toxic and should be diked for containment (Association of American Railroads) .

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## **Appendix D: TIC-Master Fact Sheet**



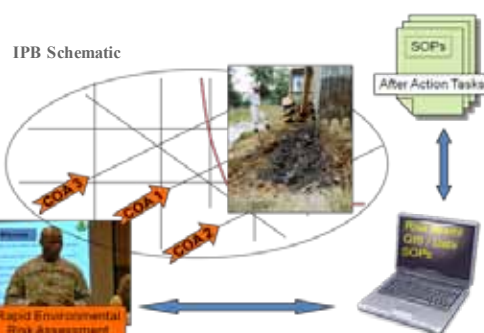
November 2010

## TIC Master and its Future with the IPB Process

Toxic Industrial Chemical (TIC) exposures may result during military urban operations involving industrial or manufacturing facilities. Even if soldiers are adequately protected, TIC releases can affect neutral or friendly populations in the environs of an urban operation. A better understanding and mapping of the types of TICs that may be present in a battlefield from industrial sources is needed as an important part of an assessment tool that could be incorporated into the Intelligence Preparation of the Battlefield (IPB) process.

### TIC Database/GIS Platform Integration:

A database has been built that currently contains over 2,000 TICs, primarily from the International Task Force 40 and EPA Toxic Release Inventory (TRI). The database includes physical properties, hazard assessment information, and environmental effects and remediation information. A data model has been built to relate these chemicals and industrial types as listed as 6-digit codes from the North American Industrial Classification System (NAICS). Data to relate the chemicals to the industries was queried from the TRI database. Analysis of the data allows for a crude estimation of the probability to encounter specific TICs at a particular industrial facility. A custom GIS standards based platform has been designed to specifically map industrial facilities within a designated region. The integration of the database and the GIS platform allows relevant information about the chemical(s) to be queried from the TIC database for a specific site and the mapping of areas of highest hazard probability. This technology offers a physical layout of industrial chemical hazards that can be the basis for the continuation of a more comprehensive effort to assess potential battlefields for TIC hazards in regions with less available data (OCONUS (Outside the Continental United States)).



### Test 5: Comparison of Two Facilities

- Evaluated potential operation against 2 facilities, determined which presented least risk.
- Successfully determined industries from partial list of chemicals
- Determined community risk (schools, population etc.)



### Developmental Testing:

A series of test scenarios were created by a semi-independent team to test the predictive aspects of the integrated geospatial system along with the TIC database. The tests were conducted using actual industrial facilities in the Vickburg Mississippi area using EPA Tier 2 data. There were five scenarios given to the test team that represented a broad spectrum of industries and hazardous situations. The industrial process, TIC release mechanism (depending on military action) and atmospheric conditions changed with each. Each scenario provided the geospatial location of the industrial facilities and various pieces of information such as industry type, partial list of chemicals, possible hazards, and/or other site characteristics. Sites varied in proximity to the community and geographic features. A limitation during the testing was that some of the chemical, industry, and hazard prediction was done manually because the analytic programming has not been completed.

## Appendix E: Email from the Army Intelligence Center

-----Original Message-----

From: Enright, Kevin W CIV USA USACE [mailto:[kevin.enright@us.army.mil](mailto:kevin.enright@us.army.mil)]

Sent: Friday, April 10, 2009 2:24 PM

To: Medina, Victor F ERDC-EL-MS

Subject: RE: Thank you (UNCLASSIFIED)

Classification: UNCLASSIFIED

Caveats: NONE

Victor: Again, sorry about the small audience.

My thoughts were to keep it unclassified so disseminate of your product can be used by many not just a few.

In addition, we discussed doing a comparison between open source and Classified data holding (Automated Intelligence File; AIF) or the (Modernized Integrated Database; MIDB) to determine if there was a vast difference or no difference. This would lead you to evaluate whether or not to enter the classified realm.

I do believe that what you are planning will be helpful for maneuver commanders by showing them a new aspect of the battlespace and its effect on friendly and enemy movement.

We talked a little about interfacing with weather to obtain the latest on atmospheric conditions to provide more information on the spread of TIC/TIMs.

I have sent your brief to others and if they have any comments I will forward to you. Kevin

Modernized Integrated Database (MIDB) is a Department of Defense Intelligence Information System (DODIIS) Intelligence Mission Application, serves as the primary repository for data production and dissemination of military intelligence involving worldwide orders of battle, facilities, command and control networks, targeting, battle damage assessments, and other related information required for strategic assessments and national policy decision-making.

## **Appendix F: White Paper for Modification of TIC-Master for Use by installations**



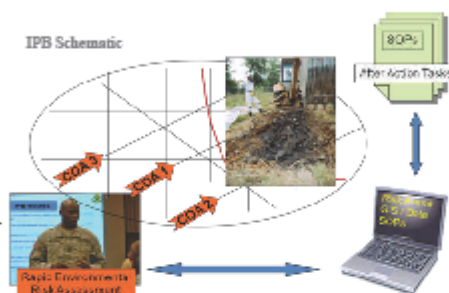
November 2010

## Installation Chemical Planning, Mapping and Response Tool

Previously, ERDC developed a GIS linked database tool that allows military planners to predict toxic industrial chemicals from industrial facilities during the Intelligence Preparation of the Battlefield (IPB). Modifying this tool for application at military facilities could have tremendous value in planning and responding to chemical incidents.

### TIC Database/GIS Platform Integration:

A database has been built that currently contains over 2,000 TICs, primarily from the International Task Force 40 and EPA Toxic Release Inventory (TRI). The database includes physical properties, hazard assessment information, and environmental effects and remediation information. A data model has been built to relate these chemicals and industrial types as listed as 6-digit codes from the North American Industrial Classification System (NAICS). Data to relate the chemicals to the industries was queried from the TRI database. Analysis of the data allows for a crude estimation of the probability to encounter specific TICs at a particular industrial facility. A custom GIS standards based platform has been designed to specifically map industrial facilities within a designated region. The integration of the database and the GIS platform allows relevant information about the chemical (s) to be queried from the TIC database for a specific site and the mapping of areas of highest hazard probability. This technology offers a physical layout of industrial chemical hazards that can be the basis for a more localized mapping of chemical hazards associated with buildings and facilities at military installations.



### Test 5: Comparison of Two Facilities

- Evaluated potential operation against 2 facilities, determined which presented least risk.
- Successfully determined industries from partial list of chemicals
- Determined community risk (schools, population etc.)



### Developmental Testing:

A series of test scenarios were created by a semi-independent team to test the predictive aspects of the integrated geospatial system along with the TIC database. The tests were conducted using actual industrial facilities in the Vickburg Mississippi area using EPA Tier 2 data. There were five scenarios given to the test team that represented a broad spectrum of industries and hazardous situations. The industrial process, TIC release

mechanism (depending on military action) and atmospheric conditions changed with each. Each scenario provided the geospatial location of the industrial facilities and various pieces of information such as industry type, partial list of chemicals, possible hazards, and/or other site characteristics. Sites varied in proximity to the community and geographic features. A limitation during the testing was that some of the chemical, industry, and hazard prediction was done manually because the analytic programming has not been completed.

**CAUTION CHEMICAL HAZARD**

**DPW Application:**

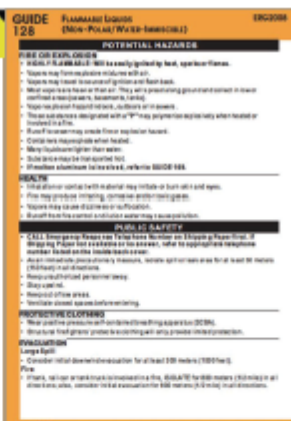
This tool can be modified to be used as a chemical planning and response tool for use at military facilities. The goal will be develop a tool that will allow installations to map their chemical containing facilities/buildings on a GIS system along with chemical inventories and quantities. Capabilities will include providing a library of properties and response information for a wide range of chemicals; release modeling for spills, atmospheric releases, fire, and explosion; and the ability to develop different scenarios for aid in planning and training. All the information necessary to manage chemical incidents would be in one fast and easy to use format. Access to the system could be done on a laptop computer, and perhaps eventually with a portable handheld device.


### Steps to Developing Installation Chemical Planning, Mapping and Response Tool:

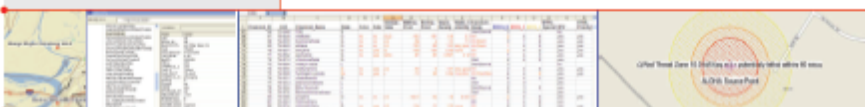
1. Implement previously suggested improvements to chemical database.
2. Obtain detailed information from targeted installation: buildings, inventory, procedures, etc.



- HAZARDOUS CHEMICALS**



- ★ Improved response to chemical incidents by having critical chemical information in an easy to use format.
  - ★ The ability to use the tool to develop different scenarios will aid in planning and training.
  - ★ Simple computing requirements
  - ★ Eliminate the hardcopy references
4. Focus GIS system on installation details including terrain and special locations of interest.
  5. Incorporate improved release modeling program in place of Aloha and include liquid/surface release modeling capability.
  6. Develop user interface program for intuitive easy-to-use access to GIS system, database, and methodology for spill management.
  7. Test tool for functionality and applicability.
- 



U.S. Army Corps  
of Engineers®  
Engineering Research and  
Development Center

Victor F. Medina, Ph.D., P.E.  
Environmental Laboratory  
3903 Halls Ferry Rd., Vicksburg, MS 39180-6799  
Victor.F.Medina@usace.army.mil  
601-634-4283

K. James Hay, Ph.D.  
Construction Engineering Research Laboratory  
2902 Newmark Dr., Champaign, IL 61822-3076  
Karl.J.Hay@usace.army.mil  
312.373.3495

## **Appendix G: Endorsement from MANSCEN**



REPLY TO  
ATTENTION OF:

ATZT-CDB

DEPARTMENT OF THE ARMY  
U.S. ARMY MANEUVER SUPPORT CENTER AND FORT LEONARD WOOD  
320 MANSCEN LOOP STE 316  
FORT LEONARD WOOD, MISSOURI 65473-8929

JUN 26 2009

MEMORANDUM FOR DIRECTOR, U.S. ARMY ENGINEER RESEARCH &  
DEVELOPMENT CENTER, 3909 HALLS FERRY ROAD, VICKSBURG, MS  
39180-6199

SUBJECT: Endorsement of Proposed Environmental Intelligence Preparation of the Battlefield  
Work Package

1. The U.S. Army Maneuver Support Center endorses the proposed work package titled Environmental Intelligence Preparation of the Battlefield.
2. This effort supports the following Warfighter Outcomes:
  - a. Detection and Identification of Toxic Industrial Chemicals and Material (TIC/TIM)
  - b. Detect, identify, and neutralize CBRN/TIM agent dispersal modes (sprayers, CBRN-filled IEDs, bomblets, etc) at standoff
  - c. Computer-aided Analysis and Reasoning
  - d. Geospatial Data Management
3. This project is developing a database of significant Toxic Industrial Chemicals, including their important physical and chemical properties, and including information on their environmental transport, fate, and remediation approaches. This database will allow users to access these chemicals if the industrial processes are known and, conversely, estimate industries based on identified chemicals. This is being linked to a GIS system that will allow mapping of the industrial sites and allow plume modeling to estimate effects. A successful product generated from the project could aid in identifying environmental hazards associated with operational planning and in the evaluation of environmental problems associated with overseas bases.

ATZT-CDB

SUBJECT: Endorsement of Proposed Environmental Intelligence Preparation of the Battlefield Work Package

4. The MANSCEN POC is Dennis Hutchinson, [dennis.hutchinson@us.army.mil](mailto:dennis.hutchinson@us.army.mil), commercial (573) 563-5253, DSN 676.



PETER J. TABACCHI  
COL, GS

Director, Capability Development  
and Integration Directorate

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14. ABSTRACT The Army needs to manage threats from toxic industrial chemicals (TICs) during the Intelligence Preparation of the Battlefield (IPB) Process. This work unit involved developing a database of TICs that could pose environmental/operational risk, creating an industry/chemical crosswalk, linking the chemical database to a GIS component, applying and evaluating the system in five hypothetical scenarios. The resultant program, which was christened "TIC-Master," has the capability to enhance urban theater of operations analysis to identify TIC/toxic industrial material (TIM) affiliated activities within the battlefield operational footprint and evaluate consequence management alternatives for decision support during the IPB process. TIC-Master was tested in a series of demonstrations. Facilities in the Jackson/Vicksburg MS area were used to test the program. Actual chemical data were obtained for several facilities by securing Tier 2 EPA (Environmental Protection Agency) reports for the facilities of interest. A test team was assembled to develop scenarios based on the industries represented by the facilities. TIC-Master was then used to address the scenarios. Although some deficiencies were identified, the program was successful at encompassing the chemical risk presented by the TICs found at the facilities. Plume modeling was used to assess release scenarios. Subsequent briefings of the program were conducted with potential users at the Army Maneuver Support Center of Excellence (MSCoE) and the Army Intelligency Center (AIC). Constructive suggestions were received and are presented in this document. Overall, the potential users agree that the TIC-Master could be a valuable tool for the Army and other organizations. Ultimately, populating tactical decision-making tools with critical information leads to a greater knowledge of the battlefield and increases survivability.					
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